

MDCAT Chemistry Chapter 9 Period Online Test

Sr	Questions	Answers Choice
1	Which one of the followings has polar covalent bonds hut is overall nom-polar molecule:	A. HF B. CO ₂ C. CH ₄ D. N ₂
2	Energy required to remove electron from an atom	A. Ionization potential B. Electronegativity C. Electropositivity D. Electron affinity
3	Elements having high ionization potential values are	A. metals B. non- metal C. liquids D. solid
4	Bonding in MgO is an example of	A. Ionic bond B. Polar bond C. Covalent bond D. Coordination covalent bond
5	Geometry of simple molecule with sp ² hybridization	A. Triangular planar B. Trigonal C. Square planner D. Pyramidal
6	At compromise distance the forces dominating between atoms are	A. repulsive forces B. attractive forces C. Dipole induced dipole force D. H-bonding
7	what is the exact value of angle in BF ₃	A. 90 B. 104.51 C. 119.5 D. 120°
8	Octet rule is not allowed in the formation of	A. NF ₃ B. B.CF ₄ C. CCl ₄ D. PCl ₅
9	Bond will be covalent when electronegativity difference of bonded atom is	A. Equal to 1.7 B. between 0.5 to 1.7 C. Greater to 1.7 D. zero
10	In a group, the atomic radii from top to bottom	A. increase B. decrease C. don't change D. show variable trend
11	The shielding effect of inner electron is responsible for	A. Having no effect on ionization energy B. Decreasing ionization energy C. Increasing ionization energy D. Increasing electronegativity
12	What is true for a molecule with standard geometry	A. It lacks a lp B. It can't be a donor C. It can be an acceptor D. All
13	Which of the following molecule has zero dipole moment?	A. PCl ₃ B. BF ₃ C. NH ₃ D. H ₂ O
14	For formation of ionic bond, electronegativity difference should be	A. Equal to zero B. Equal to 0.5 C. More than 1.7 D. Less than 1.7
15	A molecule which contains two lone pairs and two bond pairs of electrons in valence shell of central atom, geometrical shape of molecules will be	A. Tetrahedral B. Trigonal pyramidal C. Angular D. ...

D. Linear

16 Among the following molecules, which one has coordinate covalent (dative) bond?

- A. CCl_4
- B. CO_2
- C. CO
- D. CH_4

17 Total number of valence electrons in phosphonium ion (PH_4^-) is

- A. 8
- B. 9
- C. 12
- D. 10

18 Carbon-Carbon double bond length in C_3H_6

- A. 154 pm
- B. 134 pm
- C. 120 pm
- D. 105 pm

19 The ionization energy of hydrogen atom is

- A. Zero
- B. 131.3 kJ/mole
- C. 13.13 kJ/mole
- D. 1313 kJ/mole

20 A covalent bond may be

- A. 100% covalent
- B. Partial ionic
- C. 100% ionic
- D. Both a and b