

## MDCAT Chemistry Chapter 9 Period Online Test

Sr	Questions	Answers Choice
1	Elements having high ionization potential values are	A. metals B. non- metal C. liquids D. solid
2	At compromise distance the forces dominating between atoms are	A. repulsive forces B. attractive forces C. Dipole induced dipole force D. H-bonding
3	Low IE is a symbol of	A. high electronegativity B. small size C. High electron affinity D. Metallic character
4	Elements of group IA and IIA are	A. electronegative B. neutral C. electropositive D. non-metals
5	The ionization energy	A. generally increases from left to right in a period B. increases from top to bottom in a group C. does not change in a period D. does not change in a group
6	A molecule which contains two lone pairs and two bond pairs of electrons in valence shell of central atom, geometrical shape of molecules will be	A. Tetrahedral B. Trigonal pyramidal C. Angular D. Linear
7	Which one of the following has zero dipole moment	A. NH <sub>3</sub> B. CHCl <sub>3</sub> C. H <sub>2</sub> O D. BF <sub>3</sub>
8	Octet rule is not allowed in the formation of	A. NF <sub>3</sub> B. B.CF <sub>4</sub> C. CCl <sub>4</sub> D. PCl <sub>5</sub>
9	Total number of valence electrons in phosphonium ion (PH <sub>4</sub> <sup>+</sup> ) is	A. 8 B. 9 C. 12 D. 10
10	The no. of lp's on oxygen in CO are	A. 1 B. 3 C. 4 D. 2
11	Geometry of simple molecule with sp <sup>2</sup> hybridization	A. Triangular planar B. Trigonal C. Square planner D. Pyramidal
12	For formation of ionic bond, electronegativity difference should be	A. Equal to zero B. Equal to 0.5 C. More than 1.7 D. Less than 1.7
13	The ionization energy of hydrogen atom is	A. Zero B. 131.3kJ/mole C. 13.13kJ/mole D. 1313kJ/mole
14	In a period the atomic radii	A. increase B. decrease C. remain same D. first increase, then decreased
15	A covalent bond may be	A. 100% covalent B. Partial ionic C. 100% ionic

D. Both a and b

16 Ionic bond is produced after complete transfer of

- A. nucleus
- B. neutrons
- C. electrons
- D. protons

17 pi-bond can be formed by sideways overlap of

- A. s-orbital
- B. d-orbital
- C. p-orbital
- D. sp orbital

18 Greater shielding effect corresponds to ionization potential value

- A. greater
- B. lesser
- C. remain same
- D. no effect

19 What is true for a molecule with standard geometry

- A. It lacks a lp
- B. It can't be a donor
- C. It can be an acceptor
- D. All

20 A molecule that has polar bonds but is overall non - polar

- A. IF
- B. CCl4
- C. PCl3
- D. All