

MDCAT Chemistry Chapter 9 Period Online Test

| Sr | Questions | Answers Choice |
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| 1 | Elements having high ionization potential values are | A. metals B. non-metal C. liquids D. solid |
| 2 | At compromise distance the forces dominating between atoms are | A. repulsive forces B. attractive forces C. Dipole induced dipole force D. H-bonding |
| 3 | Low IE is a symbol of | A. high electronegativity B. small size C. High electron affinity D. Metallic character |
| 4 | Elements of group IA and IIA are | A. electronegative B. neutral C. electropositive D. non-metals |
| 5 | The ionization energy | A. generally increases from left to right in a period B. increases from top to bottom in a group C. does not change in a period D. does not change in a group |
| 6 | A molecule which contains two lone pairs and two bond pairs of electrons in valence shell of central atom, geometrical shape of molecules will be | A. Tetrahedral B. Trigonal pyramidal C. Angular D. Linear |
| 7 | Which one of the following has zero dipole moment | A. NH ₃ B. CHCl ₃ C. H ₂ O D. BF ₃ |
| 8 | Octet rule is not allowed in the formation of | A. NF ₃ B. BCl ₃ C. CCl ₄ D. PCl ₅ |
| 9 | Total number of valence electrons in phosphonium ion (PH ₄ ⁺) is | A. 8 B. 9 C. 12 D. 10 |
| 10 | The no. of lp's on oxygen in CO are | A. 1 B. 3 C. 4 D. 2 |
| 11 | Geometry of simple molecule with sp ² hybridization | A. Triangular planar B. Trigonal C. Square planar D. Pyramidal |
| 12 | For formation of ionic bond, electronegativity difference should be | A. Equal to zero B. Equal to 0.5 C. More than 1.7 D. Less than 1.7 |
| 13 | The ionization energy of hydrogen atom is | A. Zero B. 131.3kJ/mole C. 13.13kJ/mole D. 1313kJ/mole |
| 14 | In a period the atomic radii | A. increase B. decrease C. remain same D. first increase, then decreased |
| 15 | A covalent bond may be | A. 100% covalent B. Partial ionic C. 100% ionic |

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| | | D. Both a and b |
| 16 | Ionic bond is produced after complete transfer of | A. nucleus B. neutrons C. electrons D. protons |
| 17 | pi-bond can be formed by sideways overlap of | A. s-orbital B. d-orbital C. p-orbital D. sp orbital |
| 18 | Greater shielding effect corresponds to ionization potential value | A. greater B. lesser C. remain same D. no effect |
| 19 | What is true for a molecule with standard geometry | A. It lacks a lp B. It can't be a donor C. It can be an acceptor D. All |
| 20 | A molecule that has polar bonds but is overall non - polar | A. IF B. CCl ₄ C. PCl ₃ D. All |