

## MDCAT Chemistry Chapter 7 Reaction Kinetics Online Test

Sr	Questions	Answers Choice
1	The enthalpies of all elements in their standard states are	A. Unity B. always +ve C. always -ve D. zero
2	The lattice energy of NaCl is	A. 787 j/ mole B. 790 kj/mol C. 780 kJ/ mol D787 kl / mole
3	The change in enthalpy of a system when one mole of the substance is completely burnt in excess of air or oxygen is called	A. Heat of reaction B. Heat of formation C. Heat of atomization D. Heat of combustion
4	Total heat content of a system is called	A. Internal energy B. Entropy C. Enthalpy D. All of these
5	The change in enthalpy when one mole of a substance is dissolved in a specified quantity of solvent at a given temperature is called	A. Heat of reaction B. Heat of solvation C. Heat of combustion D. Heat of solvent
6	$\Delta H^{\circ}$ represent the enthalpy change at	A. 0°C and 1 atm pressure B. 25°Cand 1atm C. 0K and I atm pressure D. 25°C and 2 atm pressure
7	The enthalpy of formation of a compound is	A. Positive B. Either positive or negative C. Negative D. None
8	What is not correct about ΔHF	A. It is always negative B. Its value gives an idea about the relative stability of reactants and the products. C. Its value can be greater or less than zero D. Value depends upon nature of bonds
9	During an exothermic or endothermic reaction which one of the following formula is used to calculate the amount of heat evolved or absorbed	A. $\Delta H = \Delta E + PV$ B. $\Delta E = q + w$ C. $\Delta p = \Delta H$ D. $q = m \times s \times \Delta T$
10	Enthalpy of a system can be calculated by which of following relationship	A. q=ΔE B. q=m×S×ΔT C. q=pv D. q=m×v×ΔT
11	The net heat change in a chemical reaction is the same whether it is brought about in two or more different ways in one or several steps.it is known as	A. Henry's law B. Hess's law C. joule's law D. Law of conservation of energy
12	NaOH+HCI- NaCI+ H2O. Enthalpy change in the above reaction is called	A. Enthalpy of reaction B. Enthalpy of Neutralisation C. Enthalpy of formation D. Enthalpy of combustion
13	The enthalpy change for the reaction C2H2 + 5/2 O2> 2CO2 + H2O is known as enthalpy of	A. Fomation of CO2 B. Fusion of C2H4 C. Combustion of C2H4
14	How much heat is absorbed by 100 g of water when its temperature decreases from 25°C to 5°C? (heat capacity is 4.2 J/gK)	D. Vaporization of C2H2  A. 84,000J B. 2000/4.2J C2000/4.2j D8400J

15	Whenever a reaction is endothermic, then it means that	A. Heat is transferred system to the surrounding B. Heat is transferred from surrounding to the system C. Heat content of the products is less than that of reactants D. Heat content of the reactants is greater than the products
16	In order to determine $\Delta H$ (latt) of ionic compound which is correct relationship	A. $\triangle H$ latt. = $\triangle Hf$ - $\triangle Hx$ B. $\triangle H$ latt. = $\triangle Ha$ + $\triangle Hv$ C. $\triangle H$ latt. = $\triangle Hf$ + $\triangle Hx$ D. $\triangle H$ latt. = $\triangle Hf$ - $\triangle H$ sol.
17	Which of the following enthalpy change always have a negative value	A. ΔHf B. ΔH sol C. ΔHc D. ΔHat
18	Choose from the followings the correct statement about Born Haber cycle	A. Born Haber cycle is different from Hess's law B. The energy changes in a cyclic process is not zero C. The lattice energy of crystalline substances can be calculated easily D. None
19	If a reaction involves only solids and liquids, which of the following is true?	A. $\Delta H = \Delta E$ B. $\Delta H = \Delta E$ C. $\Delta H \otimes gt; \Delta E$ D. $\Delta H = \Delta E + nRT$
20	A system absorbs 100 kJ heat and performs 50 kJ work on the surroundings. The increase in internal energy of the system is	A. 50kJ B. 100 kJ C. 150kJ D. 5000 kJ