

MDCAT Chemistry Chapter 6 Chemical Equilibrium Online Test

Sr	Questions	Answers Choice
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1	For a chemical reaction to occur	A. The vessel shall be open B. Reacting molecules should have less energy than Ea at time of collision C. Reacting molecules must be properly oriented and energy more than or equal to Ea D. The reacting molecules must not collide with each other
2	The order of reaction provides valuable information about of reaction	A. Condition B. Concentration C. Mechanism D. Parameters
3	Rusting of iron is the example of	A. Fast B. Slow C. moderate D. depends upon conditions
4	If reactants are conductor of electricity, then method is used to measure the change in concentration of reaction	A. Optical rotationB. RefractrometricC. DilatometricD. Electrical conductivity
5	The rate of reaction between A and B increases by a factor of 100, when the concentration of A is increased 10 folds, the order of reaction with respect to A is	A. 10 B. 1 C. 4 D. 2
6	When the concentration of reactants is taken as unity the rate of reaction is equal to	A. average rate B. concentratian of reactant C. instantaneous rate D. specific rate constant
7	The radioactive disintegration of 238U92 is	A. First order B. Second order C. Third order D. Zero order
8	If the rate of the reaction is equal to the rate constant, the order of the reaction is	A. 3 B. 1 C. 0 D. 2
9	In the reaction A+B→ Products, if B is taken in excess, then it is an example of	A. Second order reaction B. zero order reaction C. Pseudo first order reaction D. first order reaction
10	The collision which results in chemical reaction	A. Effective collision B. Ineffective collision C. Useless collision D. All of the above
11	Rate of which reaction increases with temperature?	A. Exothermic and endothermic reactions B. Endothermic reactions C. Exothermic reactions D. None of these
12	The increase in reaction rate as a result of increase in temperature from 10K to 90K is	A. 512 B. 256 C. 400 D. 112
13	The concentration of product is increasing from 30 mole/dm3 to 40mol/dm3 in 0.5 sec then rate of reaction will bemoledm-3sec-1	A. 0 B. 20 C. 15 D. 25
14	When the concentration of product is increased the instantaneous rate of reaction with reference to reactants will be	A. Positive B. Negative C. the same

		D. falling curve
15	The number of atoms or molecules whose concentrations determines the rate of a chemical reaction is called the	A. Molecularity of the reaction B. specific activity of the reaction C. Order of the reaction D. rate constant of the reaction
16	All the Hydrolytic reactions are	A. First order B. Second order C. Third order D. pseudo-first order
17	Which property of liquid is measured by polarimeter	A. Conductance B. Optical activity C. Refractiye Indéx D. Change in volume
18	The reaction which is zero order	A. Decomposition of N2O5 B. Formation of Glucose in plant C. Formation of Fel2 D. Chorination of methane in sunlight
19	Which of the following statement about the order of reaction is true?	A. The order of reaction can only be determined by experiment B. a second order reaction is also bimolecular C. The order of reaction is always non-zero D. The order of reaction increases with increasing temperature
20	For reaction of methane and chlorine light is not available then	A. Reaction will take place rapidly B. No Reaction take place C. Reaction occurs at double the ra D. May all cases occur