

MDCAT Chemistry Chapter 6 Electrochemistry Online Test

Sr	Questions	Answers Choice
1	In the reaction $A+B \rightarrow \text{Products}$, if B is taken in excess, then it is an example of	A. Second order reaction B. zero order reaction C. Pseudo first order reaction D. first order reaction
2	When does average rate become equal to instantaneous rate of reaction	A. At the start of reaction B. time interval is zero C. at the end of reactor D. time interval approaches zero
3	A reaction $A \rightarrow B$ is independent of concentration of reactant A. The order of reaction will be	A. First order B. Second order C. Third order D. Zero order
4	If the energy of the activated complex lies close to energy of reactants, it means that reaction is	A. Slow B. Exothermic C. Endothermic D. Exothermic and fast
5	The collision which results in chemical reaction	A. Effective collision B. Ineffective collision C. Useless collision D. All of the above
6	Consider gas is measure in bars then the units of rate of reaction is	A. Mole $\text{dm}^{-3} \text{ sec}$ B. Bars sec C. Mole $\text{dm}^{-3} \text{ sec}^{-1}$ D. Bars sec^{-1}
7	The reaction which is zero order	A. Decomposition of N_2O_5 B. Formation of Glucose in plant C. Formation of FeI_2 D. Chlorination of methane in sunlight
8	When the concentration of product is increased the instantaneous rate of reaction with reference to reactants will be	A. Positive B. Negative C. the same D. falling curve
9	All the Hydrolytic reactions are	A. First order B. Second order C. Third order D. pseudo-first order
10	Substance which is formed as well as consumed during a chemical reaction and have temporary existence.	A. Reactant B. product C. Catalyst D. Intermediate
11	Which property of liquid is measured by polarimeter	A. Conductance B. Optical activity C. Refractive Index D. Change in volume
12	When the concentration of reactants is taken as unity the rate of reaction is equal to	A. average rate B. concentration of reactant C. instantaneous rate D. specific rate constant
13	The study of which one of the followings guides to the mechanism of the reaction	A. Order of reaction B. Rate of reaction C. Half-life period of reaction D. Rate determining step
14	The radioactive disintegration of ^{238}U is	A. First order B. Second order C. Third order D. Zero order
15	Higher the surface area available for reaction	A. slower the reaction B. faster the reaction C. constant the reaction D. lower the E_a

16	The slope of the graph is steepest at the beginning of reaction showing	A. Rapid decrease in concentration of reactants B. Rapid increase in concentration of reactants C. Fast rate of reaction D. All of the above
17	The rate of reaction between A and B increases by a factor of 100, when the concentration of A is increased 10 folds, the order of reaction with respect to A is	A. 10 B. 1 C. 4 D. 2
18	The rate of reaction for a reaction is $30 \text{ mol dm}^{-3}\text{sec}^{-1}$ if the product of concentration of 10.reactant is unity the specific rate constant is	A. 25 B. 2.5 C. 30 D. 15
19	If reactants are conductor of electricity, then method is used to measure the change in concentration of reaction	A. Optical rotation B. Refractometric C. Dilatometric D. Electrical conductivity
20	Which of the following statement about the order of reaction is true?	A. The order of reaction can only be determined by experiment B. a second order reaction is also bimolecular C. The order of reaction is always non-zero D. The order of reaction increases with increasing temperature