

MDCAT Chemistry Chapter 6 Electrochemistry Online Test

Sr	Questions	Answers Choice
1	If reactants are conductor of electricity, then method is used to measure the change in concentration of reaction	A. Optical rotation B. Refractometric C. Dilatometric D. Electrical conductivity
2	If the energy of the activated complex lies close to energy of reactants, it means that reaction is	A. Slow B. Exothermic C. Endothermic D. Exothermic and fast
3	Amount of product formed increases with time, this statement is true for reactions-----with kinetics	A. 1s order B. 3rd order C. zero order D. Any order
4	In dilatometric method is directly proportional to extent of reaction	A. Change in concentration B. Change in pressure C. Change in volume D. Change in temperature
5	Which property of liquid is measured by polarimeter	A. Conductance B. Optical activity C. Refractive Index D. Change in volume
6	The number of reacting molecules whose concentration change during reaction is called	A. Activated molecule B. Rate of reaction C. Order of reaction D. half-life
7	Consider gas is measure in bars then the units of rate of reaction is	A. Mole dm ⁻³ sec B. Bars sec C. Mole dm ⁻³ sec ⁻¹ D. Bars sec⁻¹
8	A reaction A- B is independent of concentration of reactant A. The order of reaction will be	A. First order B. Second order C. Third order D. Zero order
9	Reaction kinetics is important to discover the---under which reaction will proceed most economically:	A. rate constant B. Conditions C. volume D. equilibrium point
10	When the concentration of product is increased the instantaneous rate of reaction with reference to reactants will be	A. Positive B. Negative C. the same D. falling curve
11	The order of reaction provides valuable information about of reaction	A. Condition B. Concentration C. Mechanism D. Parameters
12	The slope of the graph is steepest at the beginning of reaction showing	A. Rapid decrease in concentration of reactants B. Rapid increase in concentration of reactants C. Fast rate of reaction D. All of the above
13	Unit of the rate constant depends upon the	A. Molecularity of reaction B. Order of reaction C. Concentration terms D. Number of reactants
14	All the Hydrolytic reactions are	A. First order B. Second order C. Third order D. pseudo-first order
15	For reaction of methane and chlorine light is not available then	A. Reaction will take place rapidly B. No Reaction take place C. Reaction occurs at double the rate

		D. May all cases occur
16	The number of atoms or molecules whose concentrations determines the rate of a chemical reaction is called the	A. Molecularity of the reaction B. specific activity of the reaction C. Order of the reaction D. rate constant of the reaction
17	The reaction takes place among the molecules when they have:	A. Activation energy B. Properly oriented C. Concentrated D. Activation energy and proper orientation
18	The reaction kinetics concerned with the	A. Rate of reaction B. Direction of reaction C. Factor effecting rate of reaction D. both a & b
19	Rate of which reaction increases with temperature?	A. Exothermic and endothermic reactions B. Endothermic reactions C. Exothermic reactions D. None of these
20	The radioactive disintegration of ^{238}U is	A. First order B. Second order C. Third order D. Zero order