

MDCAT Chemistry Chapter 6 Chemical Equilibrium Online Test

Sr	Questions	Answers Choice
1	When does average rate become equal to instantaneous rate of reaction	A. At the start of reaction B. time interval is zero C. at the end of reactior D. time interval approaches zero
2	Spectrometry method is applicable if a reactant or a product absorbs radiation	A. Ultraviolet B. Visible C. Infrared D. Any of these
3	Higher the surface area available for reaction	A. slower the reaction B. faster the reaction C. constant the reaction D. lower the Ea
4	For a chemical reaction in which one of the reactant also act as solvent, the order will be	A. First order B. Third order C. Second order D. pseudo-first order
5	The collision which results in chemical reaction	A. Effective collision B. Ineffective collision C. Useless collision D. All of the above
6	Half-lives required to convert 100% reactant to produet for a first order reaction are	A. 10 B. 1000 C. 100 D. Infinity
7	The reaction which is zero order	A. Decomposition of N2O5 B. Formation of Glucose in plant C. Formation of Fel2 D. Chorination of methane in sunlight
8	The rate of reaction between A and B increases by a factor of 100, when the concentration of A is increased 10 folds, the order of reaction with respect to A is	A. 10 B. 1 C. 4 D. 2
9	Amount of product formed increases with time, this statement is true for reactionswith kinetics	A. 1s order B. 3rd order C. zero order D. Any order
10	The slope of the graph is steepest at the beginning of reaction showing	A. Rapid decrease in concentration of reactants B. Rapid increase in concentration of reactants C. Fast rate of reaction D. All of the above
11	For a chemical reaction which can never be a fractional no	A. order B. molecularity C. half-life D. rate constant
12	If the reaction "P+Q \rightarrow R+S' is described as being of zero order with respect to P, it means that	A. P is catalyst in this reaction B. P molecules do not possess sufficient energy to react C. The concentration of P does not change during the reaction D. The rate of reaction is independent of the concentration of P
13	Unit of the rate constant depends upon the	A. Molecularity of reaction B. Order of reaction C. Concentration terms D. Number of reactants
		A. The vessel shall be open B. Reacting molecules should have

B. Reacting molecules should have less energy than Ea at time of

14	For a chemical reaction to occur	C. Reacting molecules must be properly oriented and energy more than or equal to Ea D. The reacting molecules must not collide with each other
15	If reactants are conductor of electricity, then method is used to measure the change in concentration of reaction	A. Optical rotationB. RefractrometricC. DilatometricD. Electrical conductivity
16	The order of reaction provides valuable information about of reaction	A. Condition B. Concentration C. Mechanism D. Parameters
17	The reaction takes place among the molecules when they have:	A. Activation energy B. Properly oriented C. Concentrated D. Activation energy and proper orientation
18	If the energy of the activated complex lies close to energy of reactants, it means that reaction is	A. Slow B. Exothermic C. Endothermic D. Exothermic and fast
19	All the Hydrolytic reactions are	A. First order B. Second order C. Third order D. pseudo-first order
20	Which property of liquid is measured by polarimeter	A. Conductance B. Optical activity C. Refractiye Indéx D. Change in volume