

MDCAT Chemistry Chapter 6 Chemical Equilibrium Online Test

Cr.	Quanting	Anguara Cheire
Sr	Questions	Answers Choice
1	Amount of product formed increases with time, this statement is true for reactionswith kinetics	A. 1s order B. 3rd order C. zero order D. Any order
2	Rate of which reaction increases with temperature?	A. Exothermic and endothermic reactions B. Endothermic reactions C. Exothermic reactions D. None of these
3	For reaction of methane and chlorine light is not available then	A. Reaction will take place rapidly B. No Reaction take place C. Reaction occurs at double the rate D. May all cases occur
4	The increase in reaction rate as a result of increase in temperature from 10K to 90K is	A. 512 B. 256 C. 400 D. 112
5	For a chemical reaction which can never be a fractional no	A. order B. molecularity C. half-life D. rate constant
6	For a chemical reaction to occur	A. The vessel shall be open B. Reacting molecules should have less energy than Ea at time of collision C. Reacting molecules must be properly oriented and energy more than or equal to Ea D. The reacting molecules must not collide with each other
7	Higher the surface area available for reaction	A. slower the reaction B. faster the reaction C. constant the reaction D. lower the Ea
8	The reaction takes place among the molecules when they have:	A. Activation energy B. Properly oriented C. Concentrated D. Activation energy and proper orientation
9	If the rate of the reaction is equal to the rate constant, the order of the reaction is	A. 3 B. 1 C. 0 D. 2
10	Reaction kinetics is important to discover theunder which reaction will proceed most economically:	A. rate constant B. Conditions C. volume D. equilibrium point
11	Which of the following statement about the order of reaction is true?	A. The order of reaction can only be determined by experiment B. a second order reaction is also bimolecular C. The order of reaction is always non-zero D. The order of reaction increases with increasing temperature
12	The reaction which is zero order	A. Decomposition of N2O5 B. Formation of Glucose in plant C. Formation of Fel2 D. Chorination of methane in sunlight
13	The concentration of product is increasing from 30 mole/dm3 to 40mol/dm3 in 0.5 sec then rate of reaction will bemoledm-3sec-1	A. 0 B. 20 C. 15 D. 25

14	The study of which one of the followings guides to the mechanism of the reaction	A. Order of reaction
		B. Rate of reaction
	3.3.	C. Half-life period of reaction
		D. Rate determining step
15	In the reaction A+B \rightarrow Products, if B is taken in excess, then it is an example of	A. Second order reaction
		B. zero order reaction
		C. Pseudo first order reaction
		D. first order reaction
	Substance which is formed as well as consumed during a chemical reaction and have temporary existence.	A. Reactant
		B. product
16		C. Catalyst
	isinpolary ontolonion	D. Intermediate
	The rate of reaction for a reaction is 30 mol dm-3sec-1 if the product of concentration of	A. 25
17		B. 2.5
	10.reactant is unity the specific rate constant is	C. 30
		D. 15
	In which of the following techniques rate of reaction is directly related with number ofions	A. Spectrometry
		B. Dilatometric method
8		C. Conductometric method
		D. Refractometric method
	Half-lives required to convert 100% reactant to product for a first order reaction are	A. 10
19		B. 1000
ı J		C. 100
		D. Infinity
	The radioactive disintegration of 238U92 is	A. First order
20		B. Second order
		C. Third order
		D. Zero order