

MDCAT Chemistry Chapter 4 Chemical Bonding Online Test

Sr	Questions	Answers Choice
1	Hydrogen bonding is not present in which of following compound?	A. Ammonia B. Ethanol C. Ether D. Water
2	The boiling of water may be 120°C, when the external pressure is	A. greater than 760 torr B. less than 760 torr C. equal to 760 torr D. variable
3	Polarizability is responsible for intermolecular forces and it	A. increases down the group B. decreases down the group C. almost remains the same D. increased along a period
4	Dipole-dipole interaction are present in the	A. atoms of the He gas B. molecules of CCl ₄ C. molecules of solid iodine D. molecules of :NH ₃
5	Which of the following liquid has highest boiling point	A. HCl B. HBr C. H ₂ O D. Br ₂
6	The nature of the attractive force in acetone and chloroform are	A. dipole-induced dipole forces B. dipole-dipole forces C. ion-dipole forces D. instantaneous forces
7	Ice floats on water because	A. the hydrogen bonding in ice is stronger than that of in water B. empty spaces are left in ice C. ice has two-dimensional structure D. the bond length of the oxygen and hydrogen bond is different in water and ice
8	Amount of heat absorbed when one mole of a solid melts into liquid form at its melting point is called:	A. heat of vaporization B. latent heat of fusion C. molar heat of fusion D. molar heat of sublimation
9	Dipole-induced dipole forces are also called	A. dipole-dipole forces B. ion-dipole forces C. Debye forces D. London-dispersion forces
10	The boiling points of the halogens	A. increases down the group B. decreases down the group C. remains constant D. can not be predicted
11	Ice occupies more space than liquid water	A. 9% B. 10% C. 11% D. 12%
12	Liquids evaporate at every temperature. When the temperature becomes constant for a liquid, then:	A. rate of evaporation is greater than the rate of condensation B. the rate of condensation is greater than the rate of evaporation C. The rate of condensation and evaporation become equal D. it depends upon the nature of the liquid
13	The weakest intermolecular forces present in a liquid may be	A. Dipole-induced dipole forces B. dipole-dipole forces C. instantaneous forces D. electrostatic forces between ions in a ionic solid
14	In order to maintain the boiling point of water at 110 C°. the external pressure should be	A. 550 torr B. between 500 and 760 tor

		<p>C. between 760 and 1500 torr</p> <p>D. any pressure can be maintained</p>
15	The B.P of glycerine at 760 torr pressure is	<p>A. 200°C</p> <p>B. 290°C°</p> <p>C. 250°C°</p> <p>D. 262°C°</p>
16	The B.P. of compound is mostly raised by	<p>A. dipole-induced dipole interactions</p> <p>B. london dispersion forces</p> <p>C. intramolecular H-bonding</p> <p>D. intermolecular H-bonding</p>
17	The vapour pressure of a liquid depends upon	<p>A. amount of the liquid</p> <p>B. surface area</p> <p>C. temperature</p> <p>D. size of container</p>
18	The long chains of amino acids are coiled around one another into a spiral by	<p>A. ionic bond</p> <p>B. Van der Waal's forces</p> <p>C. hydrogen bonding</p> <p>D. overlapping of orbitals</p>
19	Hydrogen bonding is extensively present in proteins which form the spiral. The hydrogen bond being produced is between	<p>A. nitrogen and hydrogen atom</p> <p>B. oxygen and hydrogen atom</p> <p>C. carbon and hydrogen atom</p> <p>D. oxygen and carbon atom</p>
20	Point out the substance which has maximum vapour pressure at a given temperature?	<p>A. Acetone</p> <p>B. Water</p> <p>C. Ethanol</p> <p>D. Acetic acid</p>