

## MDCAT Chemistry Chapter 4 Chemical Bonding Online Test

Sr	Questions	Answers Choice
1	The weakest intermolecular forces present in a liquid may be	A. Dipole-induced dipole forces B. dipole-dipole forces C. <b>instantaneous forces</b> D. electrostatic forces between ions in a ionic solid
2	Which of the following liquid has highest boiling point	A. HCl B. HBr C. <b>H<sub>2</sub>O</b> D. Br <sub>2</sub>
3	Liquid hydrocarbon is	A. methane B. propane C. ethane D. <b>hexane</b>
4	Halogens form halogen acids. HF is the weakest among all of them This is due to the reason that	A. fluorine is a very small-sized atom B. fluorine is highly electronegative atom C. <b>there is strong hydrogen bonding in HF</b> D. the polarity of HF bond is less
5	Which of following factor affect vapour pressure of a liquid?	A. temperature B. inter molecules forces C. size of the molecules D. <b>all of these</b>
6	The vapour pressure of a liquid depends upon	A. amount of the liquid B. surface area C. <b>temperature</b> D. size of container
7	Liquids evaporate at every temperature. When the temperature becomes constant for a liquid, then:	A. rate of evaporation is greater than the rate of condensation B. the rate of condensation is greater than the rate of evaporation C. <b>The rate of condensation and evaporation become equal</b> D. it depends upon the nature of the liquid
8	Saturated hydrocarbons having carbon atoms more than 20 in a molecule are solids due to	A. higher densities B. higher molar masses C. the chain, are more zig-zag D. <b>all are correct</b>
9	The boiling point of higher alkanes are greater than those of lower alkanes due to reason that	A. higher alkanes have greater number of atoms B. <b>the polarizabilities of higher alkanes are greater</b> C. higher alkanes have greater hydrogen bonding D. higher alkanes have zig-zag structures
10	In order to maintain the boiling point of water at 110 C°, the external pressure should be	A. 550 torr B. between 500 and 760 tor C. <b>between 760 and 1500 torr</b> D. any pressure can be maintained
11	H <sub>2</sub> O and HF are the hydrides of the second period. Fluorine is more electronegative than oxygen. Anyhow, the boiling point of water is greater than that of HF. This is due to:	A. water is more polar than HF B. water has a bent structure C. HF has a zig zag structure after making hydrogen bonding D. <b>the number of hydrogen bonds produced by water are greater than that of HF</b>
12	The nature of crystals formed due to London forces of interaction are	A. <b>molecular</b> B. metallic C. ionic D. covalent
		A. dipole-induced dipole forces

13	The forces which are present between the ions and the water molecules are known as	B. dipole-dipole forces C. ion-dipole forces D. London dispersion forces
14	The B.P. of compound is mostly raised by	A. dipole-induced dipole interactions B. london dispersion forces C. intramolecular H-bonding D. intermolecular H-bonding
15	Vapour pressure of a substance does not depend upon:	A. physical state of matter B. temperature C. intermolecular forces D. surface area
16	Dipole-dipole interaction are present in the	A. atoms of the He gas B. molecules of CCl <sub>4</sub> C. molecules of solid iodine D. molecules of :NH <sub>3</sub>
17	Ice occupies more space than liquid water	A. 9% B. 10% C. 11% D. 12%
18	At freezing point of water, the density decreases due to	A. change of bond angles B. change of bond lengths C. cubic structure of ice D. empty spaces present in the structure of ice
19	The long chains of amino acids are coiled around one another into a spiral by	A. ionic bond B. Van der Waal's forces C. hydrogen bonding D. overlapping of orbitals
20	The B.P of glycerine at 760 torr pressure is	A. 200°C B. 290°C C. 250°C D. 262°C