

MDCAT Chemistry Chapter 4 Chemical Bonding Online Test

Sr	Questions	Answers Choice
1	Vapour pressure of a substance does not depend upon:	A. physical state of matter B. temperature C. intermolecular forces D. surface area
2	The weakest intermolecular forces present in a liquid may be	A. Dipole-induced dipole forces B. dipole-dipole forces C. instantaneous forces D. electrostatic forces between ions in a ionic solid
3	The forces which are present between the ions and the water molecules are known as	A. dipole-induced dipole forces B. dipole-dipole forces C. ion-dipole forces D. London dispersion forces
4	Point out the substance which has maximum vapour pressure at a given temperature?	A. Acetone B. Water C. Ethanol D. Acetic acid
5	Oxygen and sulphur are present in VI-A group of the periodic table The hydride of oxygen i.e., H ₂ O is liquid at room temperature but the hydride of sulphur (H ₂ S) is a gas. This is due to	A. greater bond angle of water than H ₂ S B. greater bond lengths in HS than H ₂ O C. hydrogen bonding in water D. acidic character of HS
6	The boiling point of glycerin at 1 atmospheric pressure is:	A. 290°C B. 390°C C. 190°C D. 210°C
7	The vapour pressure of a liquid depends upon	A. amount of the liquid B. surface area C. temperature D. size of container
8	Strong dipole-dipole forces among the liquid molecules are responsible for	A. very high heat of vaporization B. very low heat of vaporization C. cannot be predicted D. negligible forces are these
9	Liquids evaporate at every temperature. When the temperature becomes constant for a liquid, then:	A. rate of evaporation is greater than the rate of condensation B. the rate of condensation is greater than the rate of evaporation C. The rate of condensation and evaporation become equal D. it depends upon the nature of the liquid
10	The nature of the attractive force in acetone and chloroform are	A. dipole-induced dipole forces B. dipole-dipole forces C. ion-dipole forces D. instantaneous forces
11	Ice occupies more space than liquid water	A. 9% B. 10% C. 11% D. 12%
12	The B.P. of compound is mostly raised by	A. dipole-induced dipole interactions B. london dispersion forces C. intramolecular H-bonding D. intermolecular H-bonding
13	In order to maintain the boiling point of water at 110 C°, the external pressure should be	A. 550 torr B. between 500 and 760 tor C. between 760 and 1500 torr D. any pressure can be maintained
14	At freezing point of water, the density decreases due to	A. change of bond angles B. change of bond lengths C. cubic structure of ice D. none of these

		D. empty spaces present in the structure of ice
15	The nature of crystals formed due to London forces of interaction are	A. molecular B. metallic C. ionic D. covalent
16	Dipole-dipole interaction are present in the	A. atoms of the He gas B. molecules of CCl ₄ C. molecules of solid iodine D. molecules of :NH ₃
17	Ice floats on water because	A. the hydrogen bonding in ice is stronger than that of in water B. empty spaces are left in ice C. ice has two-dimensional structure D. the bond length of the oxygen and hydrogen bond is different in water and ice
18	The boiling points of the halogens	A. increases down the group B. decreases down the group C. remains constant D. can not be predicted
19	Which of following factor affect vapour pressure of a liquid?	A. temperature B. inter molecules forces C. size of the molecules D. all of these
20	Saturated hydrocarbons having carbon atoms more than 20 in a molecule are solids due to	A. higher densities B. higher molar masses C. the chain, are more zig-zag D. all are correct