MDCAT Chemistry Chapter 22 Online Test

Sr	Questions	Answers Choice
1	Halogen are being used as fire extinguisher, mild antiseptic, CFCs and many other organic chemicals. Which of the following halogen is used to kill the bacteria in drinking water.	A. Bromine B. Fluorine C. Chlorine D. Iodine
2	The essential property of a fertilizer is that it should be.	A. Partially soluble B. Highly soluble C. In soluble D. Immiscible
3	Nitrogen is present in air as a major constituent it is an inactive gas in comparison with oxygen which is the next major constituent of air Nonreactive nature of nitrogen is due to the reason.	A. There is one lone pair of electron on each nitrogen atom in its molecule B. Nitrogen have three unpaired electron i its 2p orbital which is comparatively stable elecrtonic configuration C. There is a triple covalent bond in nitrogen molecule which in very strong and molecule is polar D. There is a triple covalent bond in nitrogen molecule which is very strong and molecule is non polar
4	Which of the following would react with ozone in the atmosphere?	A. F B. Cl C. O2 D. O
5	Which of these pollutants is produced by burning of coal and causes acid rain	A. SO2 B. CO2 C. CO D. NO
6	Which is the structure of polyvinyl chloride?	A. [H2C=CH-Cl] B. -[HCCl-CH-Cl]- C. -[H2C-CH-Cl]- D. -[CCl2-CCl2]-
7	Which one will be act as a strong acid.	A. Dichloroethanoic acid B. Emanoic acid C. Chloroethanoic acid D. Trichlorothanoic acid
8	Which product is formed by teh reaction of carboxylic acid with alcohol?	A. Aldehyde B. Ether C. Alkane D. Ester
9	Which mechanism of reaction is shown by carbonyl compounds?	A. Nucleophilic addition B. Electrophilic substituion C. Free radical substitution D. Electrophilic addition
10	Why is it necessary to distill aldehyde formed from oxidation of primary alcohol through acidified potassium dichromate (VI) solution or acidified sodium dichromate (VI) solution.	A. Aldehyde formed may be oxidised further to carboxylic and concerned B. Aldehyde formed may react with primary alcohol the original reactant C. Aldehyde formed may be oxidised further to a ketone D. Aldehyde formed is unstable and decomposed back to original precuser, ie. primary alcohol
11	Which of the following reagents is used to distinguish between aldehydes and ketones.?	A. 2,4 NDPH B. Bromine C. Alkaline Iodine D. Tollen's reagent

12	Alcohol in which carbon atom bonded to OH group is further attached with three alkyl group is	A. Aromatic alcohol B. Tertiary alcohol C. Primary alcohol D. Secondary alcohol
13	Which one of the following compounds is known as tertiary alcohol?	A. 1-Propanol B. 2-methyl-1 propanol C. 2-propanol D. 2-methyl-2-propanol
14	Which one of the following compounds act as catalyst when alcohols react with carboxylic acids.	A. Pt B. Conc. H2SO5 C. Conc HNO3 D. Ni
15	Select one which is alcohol	A. CH3-CH2-OH B. CH3COOH C. CH3-O-CH3 D. CH3-CH2-Br
16	Which compound will be produced by the oxidation of ethanol by acidified K2Cr2O7?	A. Ethanone B. Ethene C. Ethanoic acid D. Ethanol
17	Nylon -6,6 also called	A. Polyvinyl alcohol B. Polystyren C. Polyamide D. Polyester
18	Alcohol in which carbon atom bonded to OH group is further attached with three alkyl group is .	A. Aromatic alcohol B. Tertiary alcohol C. Primary Alcohol D. Secondary Alcohol
19	Which of the following acts as a nucleophile in the reaction of alkyl halide with alcoholic aqueous ammonia?	A. NH3 B. H+ C. Br- D. NO2-
20	What is teh order of increasing reactivity of alkyl halides?	A. Iodoalkane <bromoalkane<chloroalkane<Fluoroalkane B. Fluoroalkane<Chloroalkane<Bromoalkane<Iodoalkane C. Iodoalkane<Bromoalkane<Chloroalkane<Fluoroalkane D. Fluroalkane<Chloroalakane< Bromo alkane<Iodo alkane
21	Which of following compound is solid and room temperature?	A. Ethanal B. Phenol C. Butane D. Methanol
22	Halogen is a halo derivative of	A. Ethanol B. Methane C. Methanol D. Ethane
23	Reaction mechanism of alkanes with halogens is known as	A. Propagation B. Additon C. Elimination D. Free radical substitution
24	Butane molecule can have max no of isomers.	A. 4 B. 5 C. 3 D. 2
25	Which of the following acts as a electrocphile in the electrphilic substitution of benzene.with bromine.?	A. Fe⁻³ B. Br- C. FeCl⁴⁻ D. Fe⁺²
26	Which compound is obtained by the elimination of bromopropane?	A. Propene B. Ethene C. Propane D. Butane
	The species which are produced by	

27	are produced by electrolytic bond breaking and can act as electron pair donors are known as.	A. Cations B. Anions C. Nucleophiles D. Free radical
28	Ligands having two lone pair of electrons for donation to the central transition metal ion are known as.	A. Bidentate ligands B. Hexadentate ligands C. Polydentate ligands D. Monodentate ligands
29	Which is the correct electronic configuration of chromium.	A. 1s²,2s²,3s²,2p⁶,3p⁶,4s²,3d⁶ B. 1s²,2s²,2p⁶,3s⁶,3p⁶,3d⁶ C. 1s²,2s²,3s²,2p⁶,3p⁶,4s²,3d⁴ D. 1s²,2s²,2p⁶,3s²,3p⁶,4s²,3d⁵
30	Down the group acid base behavior of metallic oxides of group 2 elements changes to .	A. More basic B. No change C. Less basic D. More acidic
31	Which one the following is the structure of Teflon?	A. (-CH2-CH2-)n B. (-CF2-CH2-)n C. (-CF2-CF2-)n D. (-CF2-CCl2-)n
32	The catalyst used for the manufacture of H2SO4 by contact process is with bromine?	A. SO3 B. Pt/pd C. V2O5 D. Fe2O3
33	The dilute solution of _____ is called vinegar.	A. Formic acid B. Acetic acid C. Oxalic acid D. Benzoic acid
34	In period 1 and period 3 maximum melting point shown by elements.	A. Nitrogen and phosphorous B. Carbon and silicon C. Lithium and sodium D. Neon and argan
35	Gas is enclosed in a container of 20 cm3 with the moving piston. According to kinetic theory of gases, what will be the effect on freely moving molecules of the gas if temperature is increased from 20 ^{o}C in 100 ^{o}C?	A. Volume will be increased B. Decrease rate of a reaction C. Decrease yield of a reaction D. Increase yield of product
36	According to Lowry-bronsted acid and bass concept, H2O is	A. A salt B. An acid C. A base D. An amphoteric species
37	The potential difference of an electrochemical cell is measured by	A. Calorimeter B. Voltmeter C. Galvanometer D. Ammeter
38	The standard electrode potential of hydrogen is arbitrarily taken at 298 K is	A. 1.00 volt B. 0.00 volt C. 10.0 volt D. 0.10 volt
39	Which one of the following enthalpy change ins always exothermic?	A. Enthalpy of atomization B. Enthalpy of combustion C. Enthalpy of solution D. Enthalpy of formation
40	Relation of water with quick lime result in the rise in the temperature of the system using the concept of	A. Endothermic reaction B. Third order reaction C. Exothermic reaction D. Non spontaneous reaction

	energy change, indicate the nature of the reaction?	D. Non spontaneous reaction
41	The shape of $[Co(NH_3)_6]^{3+}$ complex is.	A. Linear B. Octahedral C. T Tetrahedral D. Square planer
42	Electron affinity of the atom is the energy released when	A. electron is removed from gaseous atom B. Covalent bond of molecule is broken C. Electron is added to gaseous atom D. Covalent bond is formed between the atom
43	Which of the following molecule has largest number of shared pair of electrons?	A. CO₂ B. N₂ C. NH₃ D. C₂H₄
44	Which option shows all the molecule with bond angle 109.5^{o}.	A. CH_{4,}CCl₄,NH₃ B. CH₄,NH₄ PH₃ C. SiCl₄ H₂O BeCl₂ D. SiCl₄ Nh₄CH₄
45	Liquid in the container have temperature 70 C. What will be the temperature in Kelvin Scale?	A. 203 K B. 350 K C. 343 K D. 300 K
46	While finding the relative atomic mass, which of the following standard is used to compare the atomic mass of chlorine .	A. Carbon-12 B. Neon -20 C. Carbon -13 D. Nucleon number
47	3.0 mole of calcium will contained_____ g of calcium.	A. 100 gm B. 105 gm C. 80 gm D. 120 gm
48	Percentage of nitrogen by volume in air is	A. 20% B. 78% C. 98% D. 50%