

## MDCAT Chemistry Chapter 11 S and P Block Elements Online Test

Sr	Questions	Answers Choice
1	which one pair has the same oxidation state of-Fe?	A. FeSO <sub>4</sub> and FeCl <sub>4</sub> B. FeCl <sub>4</sub> and FeCl <sub>3</sub> C. FeSO <sub>4</sub> and FeCl <sub>2</sub> D. Fe <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub> and FeSO <sub>4</sub>
2	Zinc does not show variable oxidation state, because	A. Its d-subshell is incomplete B. Its d-subshell is complete C. It is relatively soft metal D. It has two electrons in outermost shell
3	Which of the elements has seven electrons in d-subshell?	A. Zn B. Co C. Cu D. Fe
4	In the electronic configuration of Cr one electron from 4s sub-shell is transferred to 3d sub-shell because	A. The 3d orbital is of lower energy than 4s B. The half-filled d-subshell is more stable than 4 electrons having d-subshell C. The 4s orbital is of equal energy to 3d orbital D. 6 unpaired electron make Cr more paramagnetic
5	At which oxidation state Cu achieves electronic configuration of Zn <sup>+2</sup>	A. 0 B. +2 C. +1 D. +3
6	D-block elements are also called	A. Non-typical transition element B. Outer transition elements C. Abnormal transition elements D. Inner transition
7	All 3d series elements show an oxidation state of oxidation state	A. +1 B. +2 C. +3 D. Zero
8	Which of the following transition metal forms colourless compounds in +4 oxidation state?	A. Ti B. Cr C. Cu D. Zn
9	What is the sequence of electron take up and removal from 4s orbital a transition metal in 3d series?	A. Enters first, leaves after 3d electrons removal B. Enters after 3d electrons, leaves after 3d electrons C. Enters after 3d electrons, leaves first D. Enters first and leaves first
10	Transition compounds which occur as tripositive ions have no	A. 4s-electron B. 3p-electron C. 3s-electron D. 2s-electron
11	Which of the following pair has the same no. of electrons in d- subshell	A. Sc <sup>+3</sup> ,Ti <sup>+3</sup> B. Mn <sup>+2</sup> ,Fe <sup>+3</sup> C. Ti <sup>+3</sup> ,V <sup>+3</sup> D. Cr <sup>+3</sup> ,Co <sup>+2</sup>
12	Group VIB of transition elements contains	A. Zn. Cd. Hg B. Cr. Mo, w C. Fe. Ru, Os D. Mn. Te. Re
13	TiCl <sub>4</sub> is used as catalyst for manufacture of	A. Sulphuric acid B. Plastics C. Ethanol D. Tetraethyl lead
		A. +6

14	The maximum oxidation state of Mn is	B. +7 C. +5 D. +4
15	d-d transition cannot be shown by	A. Cu+1 B. Sc+3 C. Zn+2 D. All
16	which of the following is a typical transition metal?	A. Sc B. Y C. Ra D. Co
17	The total number of 3d-series transition elements is	A. 10 B. 40 C. 14 D. 58
18	In $[\text{Ti}(\text{H}_2\text{O})]^{3+}$ which colour is transmitted	A. Yellow B. Blue and red C. Blue and yellow D. red and yellow
19	Which of the following compound is expected to be colored	A. $\text{Na}_2\text{SO}_4$ B. $\text{ZnCl}_2$ C. $\text{MgF}_2$ D. $\text{CuF}_2$
20	The oxidation state of transition elements is usually	A. Variable B. Single C. Constant D. Infinite