

Chemistry Fsc Part 2 Chapter 1 Online Test

| Sr | Questions | Answers Choice |
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| 1 | What criteria did Mendeleev use in arranging his periodic table. | A. Atomic mass B. Atomic number C. Mass number D. Density |
| 2 | Mark the correct statement. | A. Na ⁺ is smaller than Na atom B. Na ⁺ is large than Na atom C. Cl ⁻ is smaller than Cl atom D. Cl ⁻ and Cl are equal in size |
| 3 | Which element has lowest melting point | A. Beryllium B. Magnesium C. Calcium D. Barium |
| 4 | In a period, from left to right in the periodic table, the size of atom generally. | A. Increases B. decreases C. Remains constant D. First increase upto the middle of period and then decreases |
| 5 | In which group, melting point and boiling point increase downward in a group | A. IA B. II A C. VII A D. Both a and b |
| 6 | Which element when react with chlorine form polymeric halide. | A. Na B. Be C. Ba D. P |
| 7 | Ionic Hydrides react with water to form | A. Proton B. Hydride ions C. Hydroxide ions D. Hydronium ions |
| 8 | In which compound, oxidation state of sulphur is +6 | A. H ₂ S B. H ₂ SO ₄ C. H ₂ SO ₃ D. SO ₃ |
| 9 | Elements of Groups IIA are called | A. Alkali metals B. Alkaline earth metals C. Coinage metals D. Halogens |
| 10 | In which group of periodic table is the element which has atomic number 14. | A. II B. IV C. III D. VI |
| 11 | The first ionization energy of Na, Mg, Al and Si are in the order of. | A. Na < Mg < Al < Si B. Na > Mg > Al > Si C. Na > Mg < Al < Si D. Na < Mg > Al < Si |
| 12 | In modern periodic table VI period contains elements | A. 8 B. 18 C. 10 D. 32 |
| 13 | Across a period from left to right in the periodic table, the melting and boiling point. | A. Decrease B. Increase C. Remain constant D. First increase upto the middle of period and then decrease |
| 14 | What are the total numbers of periods in the modern periodic table | A. 3 B. 5 C. 7 D. 8 |
| 15 | As you proceed across a period in the periodic table the first ionization energy | A. Decrease B. Increase C. Remains constant |

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| | | D. First increase up the middle of period and then decreases |
| 16 | Electron affinity is measure of energy | A. Required to remove the electron B. Released by adding an electron C. Required to excite an electron D. Released by removing an electron |
| 17 | Mark the incorrect statement | A. Metallic character increase down the group B. Metallic character increase from left to right along a period C. Metallic character decrease from left to right along a period D. Metallic character remains the same down the group |
| 18 | In which one of the following sets do all three particles have same number of total electrone | A. F ⁻ , Cl ⁻ , Br ⁻ B. Li ⁺ , Na ⁺ , K ⁺ C. N ³⁻ , O ²⁻ , F ⁻ , D. Na ⁺ , Mg ²⁺ , K ⁺ |
| 19 | Which one of the following oxides is amphoteric in nature. | A. MgO B. Na ₂ O C. SO ₂ D. ZnO |
| 20 | Which statement is incorrect. | A. All the metals are good conductors of electricity B. All the metals are good conductor of heat C. All the metal form positive ions D. All the metal form acidic oxides |