Acid-Base Chemistry

Sr	Questions	Answers Choice
1	A base taht ionizes completely in aqueous solution is called.	A. Strong base B. Weak base C. Conjugate acid D. Neutral
2	Which indicator is typically used for titration involving strong acids and strong bases.	A. Methyl red B. Phenolphthalein C. Bromothymol blud D. Litmus solution
3	Which of the following is the conjugate base of water.	A. OH- B. OH+ C. H2O D. H2O+
4	If Ka of acetic acid is 1.8 X 10^{-5}, IT IS	A. Strong acid B. Base C. Weak acid D. Neutral
5	NaOH is not considered an Bronsted Lowry acid because.	A. It's neutral B. It's not soluble C. Doesn't donate H+ D. Doesn't produce H+
6	The pH of a buffer doesn't chagne much when	A. Diluted with water B. Acid or base is added C. Heated D. Temp changes
7	pH of 1 x 10 -4 M HCl is	A. 4 B. 1 C. 3.5 D. 7
8	The sgtrength of base is measured by its.	A. Ka B. Kb C. Pkw D. Kw
9	Which of these is NOT a property of acids.	A. Conduct electricity B. Turn litmus red C. Sour taste D. Feel slippery
10	The pKa value for HCOOH is.	A. 4.74 B. 3.78 C. 4.78 D. 4.24
11	Which is amphoteric	A. H2O B. HCl C. NaOH D. NH3
12	CH3COONA is a salt of.	A. Strong acid+ Strong base B. Weak acid+Strong base C. None D. Weak base+Weak acid
13	The conjugate base fo H2CO3 is	A. HCO3 B. CO2 C. CO3 D. H2CO3
14	The solubility prodct of Pb SO4 is	A. 1.6 x 10⁻⁸ B. 1.4 x 10⁻⁴ C. 1.6 x 10⁻⁶ D. 1.6 x 10⁻²
15	If the concentration of Cl- ion in a solution is increaed, the soubility of silver chloride will	A. Decreasse B. Increase C. Remain unchanged D. Become zero

16	Which one is a monoprotic acid	A. H3PO4 B. H2SO4 C. HCl D. H2CO3
17	The Bronsted Lowry definition identifies acid as.	A. Electron pair acceptors B. Proton donors C. Proton acceptors D. Elecron pair donors
18	Which has the lowest pH?	A. 0.1 M CH3COOH B. 1 M HCl C. 1 M NH3 D. 0.1 M HCl
19	In a titration, the equivalence point is when	A. No acid left B. pH = 7 C. Moles of acid= moles of base D. Buffer forms
20	The solubility product of AgCl is 2.0 x 10^{-10} mol2 dm-6. The maximum concentration of Ag+ ions in the solution is.	A. 2.0 x 10⁻¹⁰ mol dm⁻³ B. 1.41 x 10⁻⁵ mol dm⁻³ C. 1.0 x 10⁻¹⁰ mol dm⁻³ D. 4.0 x 10⁻²⁰ mol dm⁻³