Chemistry Fsc Part 1 Chapter 9 Online Test

Sr	Questions	Answers Choice
1	Which solution is an example of solid in gas	A. Fog B. Steel C. smoke D. Air
2	Butter is solution of	A. Liquid in liquid B. Solid and liquid C. Liquid and solid D. Liquid and gas
3	The temperature which partially immiscible pair of liquid leads to the formation of a single phase in called.	A. Transition temperature B. Absolute temperature C. Consulate temperature D. Room temperature
4	Which pair of mixture is called ideal solution.	A. Chlorobenzene and bromobenzene B. Water alcohol C. Water ether D. HCl and water
5	A negative deviation from Raoult's law in solution means, the solution has	A. High boiling point and high vapour pressure B. High boiling point and low vapour pressure C. Low boiling point and low vapour pressure D. Low boiling point and high vapour pressure
6	Azeotropic mixture	A. Obey Raoult's law B. Do not obey Raoult's law C. Boils at low temperature only D. Boils at high temperature only
7	Relative lowering of vapour pressure is equal to.	A. Mole fraction of solute B. Mole fraction of solvent C. Mole fraction of solute and solvent D. Molality of solution
8	In case of non volatile solute, lowering of vapour pressure is proportional to.	A. Mass fraction of solute B. Mole fraction of solvent C. Mole fraction of solute D. None of the above
9	Which of the following mixtures of liquids show negative deviation	A. Methyl alcohol water B. Hydrochloric acid water C. Carbon di sulphide chloroform D. Chlorobenzene bromobenzene
10	Unit of mole fraction is	A. mol dm-3 B. mol kg-1 C. g dm-3 D. No unit
11	What amount of NaOH is required to prepare 500 g of 0.5 molal solution.	A. 10 g B. 20 g C. 30 g D. 40 g
12	Which concentration unit is independent of temperature.	A. Molarity B. Molality C. ppm D. both a and b
13	In a solution 7.8 g of benzene and 46 g of toluene is present The mole fraction of benzene is.	A. 1/2 B. 1/3 C. 1/5 D. 1/6
14	The molarity of 2% w/v NaOH solution is	A. 2 B. 0.25 C. 0.05 D. 0.5

15	The sum of mole percent of all the components of solution is always equal to.	A. Unity B. 100 C. Less than one D. Less than 100
16	10 g of NaOH have been dissolved per kg of solvent The molality of solution.	A. 0.25 m B. 0.5 m C. 1.0 m D. 2.0 m
17	10 g of NaOH has been dissolved per dm3 of solution. The molarity of solution is.	A. 0.5 M B. 0.25 M C. 1 M D. 2 M