Acid-Base Chemistry

Sr	Questions	Answers Choice
1	A salt from stron acid+weak base gives.	A. Neutral solution B. Acidic solution C. Precipitate D. Basic solution
2	CH3COONA is a salt of.	A. Strong acid+ Strong base B. Weak acid+Strong base C. None D. Weak base+Weak acid
3	The definition of amphoteric is	A. Only acid B. Only base C. Botha acid and base D. Neither
4	Which of the following is a weak base.	A. KOH B. NaOH C. NH3 D. Ca(OH)2
5	If Ka of acetic acid is $1.8 X 10^{-5}$, IT IS	A. Strong acid B. Base C. Weak acid D. Neutral
6	The logarithmic nature of pH means.	A. Each pH unit = 2 x H+ B. 10 X H+ Per unit C. H+ drop linearly D. No change in H+
7	The Ka of a weak acid is 10-5 its pKa is	A. 2 B. 3 C. 5 D. 10
8	The sgtrength of base is measured by its.	A. Ka B. Kb C. Pkw D. Kw
9	Strong bases have	A. Low Kb B. High Kb C. Low pH D. High Ka
10	The pH scale ranges typically from	A. 1 -10 B. 0 - 14 C. -1 to 1 D. 7 - 14
11	A buffer solution is.	A. Stron acid+ strong base B. Weak Acid + Its salt C. Weak base + Its salt D. Weak base+ salt of strong acid
12	The suppression of ionization of a weak electrolyte by adding common ion is	A. Le-Chatelier's effect B. Common ion effect C. Auf Bau effect D. Hund's rule
13	Litmus turns which color in a basic solution.	A. Red B. Orange C. Blue D. Colorless
14	The Bronsted Lowry definition identifies acid as.	A. Electron pair acceptors B. Proton donors C. Proton acceptors D. Elecron pair donors
15	Which has the lowest pH?	A. 0.1 M CH3COOH B. 1 M HCl C. 1 M NH3 D. 0.1 M HCl

16	The pH of a buffer doesn't chagne much when	A. Diluted with water B. Acid or base is added C. Heated D. Temp changes
17	Which of these is NOT a property of acids.	A. Conduct electricity B. Turn litmus red C. Sour taste D. Feel slippery
18	pH of blood is around.	A. 6.0 B. 7.4 C. 5.4 D. 8.0
19	Which of the following is a tribasic acid.	A. H3PO4 B. HNO3 C. HCl D. H2SO4
20	The conjugate base of HClO4 is	A. ClO4 B. ClO4 ²⁻ C. hClO3 D. H2ClO4+