Chemical Equilibrium

Sr	Questions	Answers Choice
1	Na2CO3 in wate gives.	A. Acidic solution B. Basic Solution C. Neutral solution D. Buffer
2	Le Chatelier's Principle applies to.	A. Irreversibel reactions B. Static equilibrium C. Dynamic equilibrium D. Precipitation reactions
3	Q < K implies.	A. Reaction proceeds forward B. quilibrium is established C. System stops D. Reaction fhifts in reverse
4	Removing a reactant shifts equilibrium to	A. Left B. Right C. No shift D. Depends on temperture
5	Increase in concentratio of reactants	A. Increase K B. Decrease K C. Shifs equilibrium forward D. Stops reverse reaction
6	Law of mass action was proposed by	A. Le chatelier B. Arrheneius C. Guldberg and Waage D. Dalton
7	For a specific reaction the valu eof the equilibrium constant, Kc?	A. Always remains the same at differrent reaction conditions B. Increases if the concentration of one ofthe product is increased C. Changes with changes in the temperature D. Increasees if the concentration of one of the ratants is increased
8	Dynamic equilibrium occurs.	A. In open systems B. Only in gases C. In closed systems D. Only at low temperature
9	Contact process is used for	A. Sulfuric acid production B. Ammonia Synthesis C. Nitric Acid production D. Hydrogenation
10	Kc = 0.040 at 450 oC for the teh given reaction, evaluate Kp for the reaction. Pcl5 ..... PCl2 + Cl2	A. 0.40 B. 2.4 C. 0.64 D. 0.052
11	Removing product from equilibrium	A. Shifts equilibrium left B. Stops reaction C. Shifts rquilibrium right D. Has no effect
12	Equilibrium constant depends on	A. Pressure B. Temperature C. Volume D. Concentration
13	Law of mass action is applicable to.	A. Reversible reaction B. Combustion reactions C. Irreversible reactions D. Endothermic reactions only
14	Kc is expressed in terms of.	A. Pressure B. Mole fraction C. Concentration D. Volume

15	Catalyst used in contact process	A. Fe B. V_2O_5 C. Ni D. Al_2O_3
16	A saturted solution represents a dynamic equilibrium Macroscopically, the concentration of dissolved solute is constant, Microscopically this occurs because.	A. No more solute particles are dissolving B. The rate of dissolution of solute is zero C. Solute particles are dissolving and precipitating at the same rate D. All solute particles have dissolved
17	A reversibel reaction is one which	A. Proceeds to completion B. Occurs only in one direction C. Proceeds in both directions D. Has no products
18	Which one of the following is not an example of reversible reaction.	A. Formation of ammonia B. Foramtion ow water C. Decomposition of PCl_5 D. Decomposition of NO_2
19	At equilirbium the observabel properties.	A. Keep changing B. Fluctuae randomly C. Remain constant D. Oscillate
20	Salts of weak acid + strong base are.	A. Basic B. Acidic C. Amphoteric D. Neutral