Chemical Equilibrium

Sr	Questions	Answers Choice
1	Whcih of the following reactions reaches equilibrium	A. Reversible B. Irreversible C. Combustion D. Neutralization
2	Law of mass action was proposed by	A. Le chatelier B. Arrheneius C. Guldberg and Waage D. Dalton
3	At equilirbium	A. Products dominate B. Reactants dominate C. Forward and reverse rates are zero D. Forward rate = reverse rate
4	Salts of weak acid + strong base are.	A. Basic B. Acidic C. Amphoteric D. Neutral
5	Which is NOT a feature of dynamic equilibrium.	A. Closed system B. Constant Temperature C. Uequal reaction rates D. No net change
6	Dynamic equilibrium occurs.	A. In open systems B. Only in gases C. In closed systems D. Only at low temperature
7	NH4Cl in water makes solution	A. Neutral B. Acidic C. Basic D. amphoteric
8	Untis of Kc depend on	A. Catalyst B. Reaction stoichimometty C. Activation energy D. Delta H
9	Removing a reactant shifts equilibrium to	A. Left B. Right C. No shift D. Depends on temperture
10	Catalyst affects	A. Value of K B. Equilibrium position C. Activation energy D. Final concentrations
11	Removing product from equilibrium	A. Shifts equilibrium left B. Stops reaction C. Shifts rquilibrium right D. Has no effect
12	Law of mass action applies to	A. Closed systems only B. Open systems C. Irreversibel systems D. Combustion reactions
13	A reversibel reaction is one which	A. Proceeds to completion B. Occurs only in one direction C. Proceeds in both directions D. Has no products
14	Optimum temperatur ein Haber process is	A. 50 oC B. 450 oC C. 200 oC D. 1000 oC
15	Law of mass action is applicable to.	A. Reversible reaction B. Combustion reactions C. Irreversible reactions

		D. Endothermic reactions only
16	According to law of mass action rate of reaction is proportional to.	A. Temperature B. Pressure C. Product of active masses D. Atomic mass
17	Which statement is true at dynamic equilibrium.	A. No reactio is occurring B. Concentrations are changing C. Rates of forward and reverse reactions are equal D. Rate forward reaction < reverse
18	Active mass means	A. Moles B. Volume C. Mass D. Molar concentration
19	Catalyst used in Haber process is	A. Fe B. Pt C. Cu D. Zn
20	Which does not alter equilibrium constant.	A. Catalyst B. Temperture C. Pressure D. Both a and c