Chemical Equilibrium

Sr	Questions	Answers Choice
1	Le Chatelier's Principle applies to.	A. Irreversibel reactions B. Static equilibrium C. Dynamic equilibrium D. Precipitation reactions
2	Na2CO3 in wate gives.	A. Acidic solution B. Basic Solution C. Neutral solution D. Buffer
3	Which statement is true at dynamic equilibrium.	A. No reactio is occurring B. Concentrations are changing C. Rates of forward and reverse reactions are equal D. Rate forward reaction < reverse
4	Lowering temperatur ein an exothermic reaction.	A. Favors reverse B. Favors forward C. No effect D. Stop the reaction
5	For a specific reaction the valu eof the equilibrium constant, Kc?	A. Always remains the same at differrent reaction conditions B. Increases if the concentration of one ofthe product is increased C. Changes with changes in the temperature D. Increasees if the concentration of one of the ratants is increased
6	Untis of Kc depend on	A. Catalyst B. Reaction stoichimometty C. Activation energy D. Delta H
7	At equilirbium the observabel properties.	A. Keep changing B. Fluctuae randomly C. Remain constant D. Oscillate
8	At equilirbium	A. Products dominate B. Reactants dominate C. Forward and reverse rates are zero D. Forward rate = reverse rate
9	Catalyst used in contact process	A. Fe B. V2O5 C. Ni D. Al2O3
10	Optimum temperatur ein Haber process is	A. 50 oC B. 450 oC C. 200 oC D. 1000 oC
11	According to law of mass action rate of reaction is proportional to.	A. Temperature B. Pressure C. Product of active masses D. Atomic mass
12	Active mass means	A. Moles B. Volume C. Mass D. Molar concentration
13	Reaction in Haber process is	A. Endothermic B. Exothermic C. Irrversible D. Neutral
14	A large K value indicates.	A. Prodcuts are favored B. Reactants are favored C. No reaction

		D. Slow reaction
15	Kc is expressed in terms of.	A. Pressure B. Mole fraction C. Concentration D. Volume
16	Consider the gas phase equilibrium sysem represented by teh equation 2 H2O..... 2H2+O2 Givent that the forward reactio is endothermic, which of the following chagnes will decrease the equilibrium amount of H2O.	A. Adding more oxygen B. Adding a solid phase catalyst C. Decreasing the volume of the container D. Increaseing the temperature at constant pressure
17	Which is NOT a feature of dynamic equilibrium.	A. Closed system B. Constant Temperature C. Uequal reaction rates D. No net change
18	A reversibel reaction is one which	A. Proceeds to completion B. Occurs only in one direction C. Proceeds in both directions D. Has no products
19	Law of mass action was proposed by	A. Le chatelier B. Arrheneius C. Guldberg and Waage D. Dalton
20	Addition of inert gas at constant volume	A. Affects equilibrium B. Shifts reation left C. Shifts reaction right D. No effect