Chemical Equilibrium

Sr	Questions	Answers Choice
1	Reaction quotient Q is calculated using	A. Initial conceration B. Equlibrium concentrations C. Temperature D. Catalyst
2	Which of the following is an irreversible reaction.	A. Haber process B. Precipitation of Ag Cl C. Synthesis of ammonia D. Esterification
3	Whcih of the following reactions reaches equilibrium	A. Reversible B. Irreversible C. Combustion D. Neutralization
4	Addition of inert gas at constant volume	A. Affects equilibrium B. Shifts reation left C. Shifts reaction right D. No effect
5	Which is NOT a feature of dynamic equilibrium.	A. Closed system B. Constant Temperature C. Uequal reaction rates D. No net change
6	Kp is used for.	A. Solid state reactions B. Gaseous reactions C. Aqueous reactions D. Liquid reactions
7	Catalyst affects	A. Activation energy B. Equilibrium position C. Kp value D. Enthalpy
8	Which does not alter equilibrium constant.	A. Catalyst B. Temperture C. Pressure D. Both a and c
9	Law of mass action was proposed by	A. Le chatelier B. Arrheneius C. Guldberg and Waage D. Dalton
10	A large K value indicates.	A. Prodcuts are favored B. Reactants are favored C. No reaction D. Slow reaction
11	Which statement is true at dynamic equilibrium.	A. No reactio is occurring B. Concentrations are changing C. Rates of forward and reverse reactions are equal D. Rate forward reaction < reverse
12	If Kc> 1, the reaction.	A. Favors products B. Favors reactants C. Is at equilirbium D. Does not occur
13	Contact process is used for	A. Sulfuric acid production B. Ammonia Synthesis C. Nitric Acid production D. Hydrogenation
14	Consider the gas phase equilibrium sysem represented by teh equation 2 H2O..... 2H2+O2 Givent that the forward reactio is endothermic, which of the following chagnes will decrease the equilibrium amount of H2O.	A. Adding more oxygen B. Adding a solid phase catalyst C. Decreasing the volume of the container D. Increaseing the temperature at constant pressure
		A. Closed systems only

15	Law of mass action applies to	B. Open systems C. Irreversibel systems D. Combustion reactions
16	Salts of weak acid + strong base are.	A. Basic B. Acidic C. Amphoteric D. Neutral
17	Le Chatelier's Principle applies to.	A. Irreversibel reactions B. Static equilibrium C. Dynamic equilibrium D. Precipitation reactions
18	At equilirbium the observabel properties.	A. Keep changing B. Fluctuae randomly C. Remain constant D. Oscillate
19	NH4Cl in water makes solution	A. Neutral B. Acidic C. Basic D. amphoteric
20	Kc = 0.040 at 450 oC for the teh given reaction, evaluate Kp for the reaction. Pcl5 ..... PCl2 + Cl2	A. 0.40 B. 2.4 C. 0.64 D. 0.052