Reaction Kinetics

Sr	Questions	Answers Choice
1	Intermediates are	A. Reactants B. Products C. Short lived species formed and consumed D. Stable compounds
2	The units of the rate constant (k) for a reaction depend on the	A. Activation energy of the reaction B. Temperature of the reaction C. Overall order of the reaction D. Stoichiometry of the balanced chemical equation
3	Collision frequency depends on	A. Pressure B. Concentration C. Temperature D. All of these
4	In an energy profile, peak represents.	A. Reactants B. Activated complex C. Products D. Catalyst
5	Enzyme is a	A. Solid catalyst B. Homogeneous catalyst C. Biological catalyst D. Promoter
6	Which factor affects rate of reaction.	A. Concentration B. Temperature C. Surface Area D. All of these
7	Half life formula for 1st order reaction is	A. 0.693 /k B. k x t C. 1 /k D. 2 k
8	The rate of reaction .	A. Increase as the reaction proceeds B. Decreases as the reaction proceeds C. Remains the same as the reaction proceeds D. May decrease or increse as the reaction proceeds
9	Which order has constant half life	A. First B. Zero C. Second D. Third
10	In pressure of catalyst Delta H	A. Increases B. Decreases C. Double D. Unchanged
11	For gaseous reactions, rate is ofter monitored by	A. Mass change B. pH change C. Pressure change D. Temperature change
12	Considr two reactions with different activation energies at the same temperature, The reaction with the lower activation energy will have	A. A smaller rate constatn B. A larger rate constant C. The same rate constant D. A rate constant that depends on the enthalpy change
13	Rate = k [A] ; defines	A. Stoichiometry B. Reaction order C. Reaction mechanism D. Titration
14	If we double the conentration of a reactant, the rate increases by four times, the reactionis.	A. First Order B. Second Order C. Third Order

		D. Zero order
15	Which one reduces activation energy.	A. Catlyst B. Inhibitor C. Product D. Reactant
16	Which unit is common for all orders	A. mol dm-3 B. No common unit C. J mol -1 D. s-1
17	Unit of rae constant for zero order reaction is	A. mol dm-3 s-1 B. s-1 C. mol -1 dm³ s-1 D. mol⁻² dm⁶ s-1
18	Units of rate constant depend on	A. Order of reaction B. Temperature C. Surface area D. Pressure
19	Rate constant can be determined by	A. Titration B. spectroscopy C. conductometry D. all of these
20	Effect of temperature on rate is	A. Negligible B. Linear C. Exponential D. Logarithmic