

## Reaction Kinetics

Sr	Questions	Answers Choice
1	In color change reactions, which method is best	A. Colorimetry B. Titration C. Conductometry D. Manometry
2	Rate determining step controls	A. Overall rate B. Temperature C. Volume D. Final yield
3	Enzyme is a	A. Solid catalyst B. Homogeneous catalyst C. Biological catalyst D. Promoter
4	Effective collisions are those that.	A. Have enough energy and proper orientation B. Are elastic C. Are in vacuum D. Produce no change
5	Rate of fastest	A. At start B. In middle C. At end D. After completion
6	Which of the following methods measures reaction rate.	A. Conductivity B. Titration C. Volume of gas D. All of these
7	In a chemical reaction, a catalyst.	A. Alters Delta H B. Lowers activation energy C. Is consumed D. Forms product
8	Energy profile shows	A. Concentration B. Temperature C. Energy vs reaction progress D. Time
9	The order of a chemical reaction, that is dependent of concentration is.	A. Second order reaction B. First order reaction C. Zero order reaction D. Pseudo first order reaction
10	Unit of rate constant for first order reaction is	A. mol dm <sup>-3</sup> B. mol <sup>-1</sup> dm <sup>3</sup> s <sup>-1</sup> C. s <sup>-1</sup> D. J mol <sup>-1</sup>
11	Unit of rate for gaseous reaction are.	A. mol dm <sup>-3</sup> s <sup>-1</sup> B. atm C. mol D. kg s <sup>-1</sup>
12	On an energy profile diagram the presence of a catalyst is represented by	A. A high peak representing the activation energy B. A lower peak representing the activation energy C. A change in the energy level of the reactants or products D. A shift in the equilibrium position
13	If we double the concentration of a reactant, the rate increases by four times, the reaction is.	A. First Order B. Second Order C. Third Order D. Zero order
14	Unit of half life is	A. mol B. s C. mol dm <sup>-3</sup> D. J

15	k has unit s <sup>-1</sup> for	A. First order B. Second order C. Third order D. zero order
16	How does the presence of a catalyst affect the rate of a chemical reaction	A. It always decreases the rate of the reaction B. It always increases the rate of the reaction C. It increases the rate of the forward and decreases the rate of the reverse reaction D. It increases the rate of both the forward and reverse reactions
17	In an energy profile, peak represents.	A. Reactants B. Activated complex C. Products D. Catalyst
18	A reaction mechanism describes	A. Experimental conditions B. Overall stoichiometry C. Stepwise molecular events D. Heat of reaction
19	A flat curve in rate vs time graph indicate.	A. Reaction completed B. High rate C. Constant rate D. Temperature increase
20	Half life is time in which	A. Rate becomes zero B. Concentration reduces to 1/4 C. Concentration reduces to half D. Product is half formed