Chemistry Fsc Part 1 Chapter 7 Online Test

Sr	Questions	Answers Choice
1	The rate of reaction .	A. Increase as the reaction proceeds B. Decreases as the reaction proceeds C. Remains the same as the reaction proceeds D. May decrease or increse as the reaction proceeds
2	On a Boltzmann distribution curve, the activation energy is represented by	A. The height of the peak B. The area under the entire curve C. A vertical line drawn at a specific kinetic energy value D. The difference between the peak and the X axis
3	On an energy profile diagram the presence of a catalyst is represented by	A. A highe rpeak representing the activation energy B. A lower peak representing the activation energy C. A change in the energy level of the reactants or products D. A shift in the equilibrium position
4	Catalyst usually act in	A. Initialion step B. Rate determining step C. Fast step D. Intermeiate formation
5	Increasing the temperature of a chemical reaction increases the rate of a reaction because.	A. Both the collision frequency and collision energies of reactant molecules increase B. Collision frequency of reactatn molecule increases C. Activation energy increase D. Activiation energy decrease
6	In color change reactions, which method is best	A. Colorimetry B. Titration C. Conductometry D. Manometry
7	Half life is time is which	A. Rate becomes zero B. Concentration reduces to 1/4 C. Concentratio reduces to half D. Product is half formed
8	Reaction rate by conductometry depends on	A. Ionic conductivity B. Pressure C. colour D. Temperature
9	Which one reduces activation energy.	A. Catlyst B. Inhibitor C. Product D. Reactant
10	In energy profile diagram a catalysts.	A. Increse activation energy B. Decreases activation energy C. Provide new pathway D. Both b and c
11	Which unit is common for all orders	A. mol dm-3 B. No common unit C. J mol -1 D. s-1
12	At equilibrium Delta G is	A. Zero B. Positive C. Negative D. Maximum
13	On Boltzmann distribution curve, the area under the curve represents.	A. Activation energy of the rreaction B. Total numebr f moleclues in the sample C. Average kinetic energy of the

		molecules D. Rate constant of the reaction
14	Half life of first order reaction	A. Become zero B. Increases with [A] C. Is independent of [A] D. Decreases with [A]
15	Units of rate constant depend on	A. Order of reaction B. Temperature C. Surface area D. Pressure
16	A zero order reaction has rate	A. Proportional to [A] B. Independent of [A] C. Proportional to time D. Zero
17	Delta G tells about	A. Speed B. Mechanism C. Feasibility D. All of these
18	Half life formula for 1st order reaction is	A. 0.693 /k B. k x t C. 1 /k D. 2 k
19	Unit of rate for gaseous reaction are.	A. mol dm-3 s-1 B. atm C. mol D. kg s-1
20	Collision frequency depends on	A. Pressure B. Concentration C. Temperature D. All of these