Reaction Kinetics

Sr	Questions	Answers Choice
1	In energy profile diagram a catalysts.	A. Increse activation energy B. Decreases activation energy C. Provide new pathway D. Both b and c
2	Delta G tells about	A. Speed B. Mechanism C. Feasibility D. All of these
3	Catalyst usually act in	A. Initialion step B. Rate determining step C. Fast step D. Intermeiate formation
4	Unit of rate for gaseous reaction are.	A. mol dm-3 s-1 B. atm C. mol D. kg s-1
5	Units of rate constant depend on	A. Order of reaction B. Temperature C. Surface area D. Pressure
6	Rate = k [A] ; defines	A. Stoichiometry B. Reaction order C. Reaction mechanism D. Titration
7	Rate of fastest	A. At start B. In middle C. At end D. After completion
8	Rate constant can be determined by	A. Titration B. spectroscopy C. conductometry D. all of these
9	In a chemical reaction, a catalyst.	A. Alters Delta H B. Lowers activation energy C. Is consumed D. Forms product
10	Which factor affects rate of reaction.	A. Concentration B. Temperature C. Surface Area D. All of these
11	Which increases with rise in temperature?	A. Activation energy B. Enthalpy C. Collision frequency D. Molecular weight
12	Overall order is sum of	A. Coefficients B. Exponents in rate law C. Moles D. Products
13	Effect of temperature on rate is	A. Negligible B. Linear C. Exponential D. Logarithmic
14	At equilibrium Delta G is	A. Zero B. Positive C. Negative D. Maximum
15	Half life is time is which	A. Rate becomes zero B. Concentration reduces to 1/4 C. Concentratio reduces to half D. Product is half formed

16	Rate constant dependent of	A. Time B. Temperature C. Concentration D. Catalyst
17	Increase in surface area	A. Increases rate B. Decreases rate C. No effect D. Stops reaction
18	Half life formula for 1st order reaction is	A. 0.693 /k B. k x t C. 1 /k D. 2 k
19	If we double the conentration of a reactant, the rate increases by four times, the reactionis.	A. First Order B. Second Order C. Third Order D. Zero order
20	Which order has constant half life	A. First B. Zero C. Second D. Third