Reaction Kinetics

Sr	Questions	Answers Choice
1	Which of the following methods measures reaction rate.	A. Conductivity B. Titration C. Volume of gas D. All of these
2	k has unit s-1 for	A. First order B. Second order C. Third order D. zero order
3	Rate of fastest	A. At start B. In middle C. At end D. After completion
4	In energy profile diagram a catalysts.	A. Increse activation energy B. Decreases activation energy C. Provide new pathway D. Both b and c
5	Rate of reaction is measured as.	A. Increase Temperature B. Change is concentration per unit time C. Total eergy of reactants D. Change in volume per mole
6	Units of rate constant depend on	A. Order of reaction B. Temperature C. Surface area D. Pressure
7	In a chemical reaction, a catalyst.	A. Alters Delta H B. Lowers activation energy C. Is consumed D. Forms product
8	Effective collisions are those that.	A. Have enough energy and proper orientation B. Are elastic C. Are in vacuum D. Produce no change
9	Considr two reactions with different activation energies at the same temperature, The reaction with the lower activation energy will have	A. A smaller rate constatn B. A larger rate constant C. The same rate constant D. A rate constant that depends on the enthalpy change
10	Unit of half life is	A. mol B. s C. mol dm-3 D. J
11	The lowest stepin a mechanism is	A. Fast step B. Activation step C. Rate determining step D. Initiation step
12	A reaction mechanism describes	A. Experimental conditions B. Overall stoichiometry C. Stepwise molecular events D. Heat of reaction
13	Half life formula for 1st order reaction is	A. 0.693 /k B. k x t C. 1 /k D. 2 k
14	Delta G tells about	A. Speed B. Mechanism C. Feasibility D. All of these
		A. Mass change

15	For gaseous reactions, rate is ofter monitored by	B. pH change C. Pressure change D. Temperature change
16	Homogenous catalysts are in	A. Same phase as reactionts B. Different phase C. Solid only D. Gaseous only
17	If we double the conentration of a reactant, the rate increases by four times, the reactionis.	A. First Order B. Second Order C. Third Order D. Zero order
18	Rate = k [A] ; defines	A. Stoichiometry B. Reaction order C. Reaction mechanism D. Titration
19	Unit of rate for gaseous reaction are.	A. mol dm-3 s-1 B. atm C. mol D. kg s-1
20	On Boltzmann distribution curve, the area under the curve represents.	A. Activation energy of the rreaction B. Total numebr f moleclues in the sample C. Average kinetic energy of the molecules D. Rate constant of the reaction