

## Chemistry Fsc Part 1 Chapter 7 Online Test

Sr	Questions	Answers Choice
1	On an energy profile diagram the presence of a catalyst is represented by	A. A highe rpeak representing the activation energy B. A lower peak representing the activation energy C. A change in the energy level of the reactants or products D. A shift in the equilibrium position
2	Rate of fastest	A. At start B. In middle C. At end D. After completion
3	For gaseous reactions, rate is ofter monitored by	A. Mass change B. pH change C. Pressure change D. Temperature change
4	In color change reactions, which method is best	A. Colorimetry B. Titration C. Conductometry D. Manometry
5	The rate of reaction .	A. Increase as the reaction proceeds     B. Decreases as the reaction     proceeds     C. Remains the same as the reaction     proceeds     D. May decrease or increse as the     reaction proceeds
6	Considr two reactions with different activation energies at the same temperature, The reaction with the lower activation energy will have	A. A smaller rate constatn B. A larger rate constant C. The same rate constant D. A rate constant that depends on the enthalpy change
7	In proceure of catalyst Dolta H	A. Increases B. Decreases
	In pressure of catalyst Delta H	C. Double D. Unchanged
8	Rate law is an expression involving.	
8		D. Unchanged  A. Concentration B. Pressure C. Temperature
	Rate law is an expression involving.	D. Unchanged  A. Concentration B. Pressure C. Temperature D. Catalyst  A. Catlyst B. Inhibitor C. Product
9	Rate law is an expression involving.  Which one reduces activation energy.	D. Unchanged  A. Concentration B. Pressure C. Temperature D. Catalyst  A. Catlyst B. Inhibitor C. Product D. Reactant  A. Constant B. Proportional to [A] C. Inversely proportional with [A]
9	Rate law is an expression involving.  Which one reduces activation energy.  Half life of zero order reaction is	D. Unchanged  A. Concentration B. Pressure C. Temperature D. Catalyst A. Catlyst B. Inhibitor C. Product D. Reactant  A. Constant B. Proportional to [A] C. Inversely proportional with [A] D. Independent to rate constant  A. Initialion step B. Rate determining step C. Fast step
9 10 11	Rate law is an expression involving.  Which one reduces activation energy.  Half life of zero order reaction is  Catalyst usually act in	D. Unchanged  A. Concentration B. Pressure C. Temperature D. Catalyst A. Catlyst B. Inhibitor C. Product D. Reactant  A. Constant B. Proportional to [A] C. Inversely proportional with [A] D. Independent to rate constant  A. Initialion step B. Rate determining step C. Fast step D. Intermeiate formation  A. Activation energy of the reaction B. Temperature of the reaction C. Overall order of the reaction D. Stoichiometry of the balanced

		C. Biological catalyst D. Promoter
15	Half life is time is which	A. Rate becomes zero B. Concentration reduces to 1/4 C. Concentratio reduces to half D. Product is half formed
16	Rate = k [A] ; defines	A. Stoichiometry B. Reaction order C. Reaction mechanism D. Titration
17	Unit of rate for gaseous reaction are.	A. mol dm-3 s-1 B. atm C. mol D. kg s-1
18	Rate determining step controls	A. Overall rate B. Temperature C. Volume D. Final yield
19	Reaction rate by conductometry depends on	A. lonic conductivity B. Pressure C. colour D. Temperature
20	At equilibrium Delta G is	A. Zero B. Positive C. Negative D. Maximum