

Chemistry Fsc Part 1 Chapter 6 Online Test

Cr.	Ougations	Anguara Chaire
Sr	Questions	Answers Choice
1	Delta H is negative and Delta S is position then reaction is.	A. Equilibrium B. Always spontaneous C. Temperature depends D. Non sppontaneous
2	specific heat capacity is amount of heat needed to raise temp of	A. 1 mole B. 1 gram C. 1 kg D. 100 grams
3	Which of the following is always negative in an exothermic reaction.	A. Delta H B. Activation energy C. Entropy D. Delta S
4	A process with increase in entropy and enthalpy is spontaneous at.	A. High temperature B. Low temperature C. All temperature D. never spontaneous
5	Born Haber cycle is used to calculate	A. Bond energy B. Heat of hydration C. Lattice energy D. lonization energy
6	The SI unit of enthalpy	A. Calorie B. eV C. kj mol-1 D. J mol-1
7	Which of the following has maximum heat capacity	A. Water B. Mercury C. Ethanol D. Iron
8	Enthalpy of fusion is the heat required to.	A. Melt a solid B. Boil a liquid C. Freez a liquid D. Vaporize a solid
9	Enthalpy change of a process is measured under	A. Constant volume B. Constant pressure C. Constant temperature D. Constant Energy
10	The sign of Delta H during melting is.	A. Negative B. Positive C. Zero D. Can't be predicted
11	Which of the following affects bond energy.	A. Bond length B. Bond length C. Atomic size D. All of these
12	What type of reation has a negative Delta H and Delta S?	A. Always spntaneous B. Never spontaneous C. Spontaneous at low temp D. Spontaneous at high temp
13	Which of the following is an extensive property.	A. Enthalpy B. Temperature C. Density D. Pressure
14	Heat of combustion of H2 is used to determine	A. Calorific value B. Enthalpy C. lonization energy D. Lattice energy
15	Which of the followna factors would lead to a greater enthalov chage of hydration .	A. A larger ionic reduis and a smaller charge B. A smaller ionic radius and a smaller charge

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16	Which is not a path function	A. Work B. Heat C. Entropy D. Temperature
17	Delta H = q at constant	A. Pressure B. Energy C. Volume D. Temperature
18	The enthalpy change for a reaction depends on.	A. Pathway taken from reactants to products B. Presence of a catalyst C. Intitial and final states of the reactants and products D. Rate of the reaction
19	lon an exothermic reaction , the nergy of products is	A. Greater than reactants B. Less than reactants C. Equal to reactants D. Zero
20	If the pH of solution is 11, what is the [OH-]concentratio in the solution.	A. 1 x 10 ⁻³ M B. 1 x 10 ⁻¹¹ M C. 1 x 10 ⁻² M D. 1 x 10 ^{-14 M}