

Chemistry Fsc Part 1 Chapter 6 Online Test

Sr	Questions	Answers Choice
1	BF ₃ has zero while NH ₃ has 1.49 D dipole moment because.	A. B is less electronegative than N B. F is more electronegative than N C. BF ₃ is pyramidal while NH ₃ is planar D. NH ₃ is pyramidal while BF ₃ is trigonal planar
2	Dipole moment is defined as.	A. Charge x distance B. Charge x Debye C. Charge x displacement D. Charge x bond energy
3	Bond energy depends upon	A. Electronegativity B. Size of atom C. Bond length D. All of these
4	Which bond has more ionic characters in it.	A. C - F B. N - F C. O - F D. F - F
5	Which of the following molecules have highest bond energy	A. F ₂ B. Cl ₂ C. Br ₂ D. I ₂
6	What is bond order.	A. Number of unpaired electrons B. Number of paired electrons C. Number of electrons present in antibonding molecular orbital D. Number of bond formed between two atoms after overlap
7	A molecular orbital can accommodate maximum electron	A. 2 B. 6 C. 8 D. 10
8	According to MOT, which molecular orbital has highest energy.	A. sigma 1s B. pi+ 2s C. pi 2py D. Pi+ 2px
9	Which one of the following molecules is paramagnetic.	A. H ₂ B. He C. N ₂ D. O ₂
10	Which of the following molecules have unpaired electrons in the bonding molecular orbitals.	A. N ₂ B. O ₂ C. B ₂ D. F ₂
11	In BeCl ₂ , the covalent bond is formed due to overlap of	A. sp - s B. sp - p C. sp ² - p D. sp ³ - p
12	The percentage of s characters in sp ³ hybrid orbital is.	A. 25% B. 33.3% C. 50% D. 75%
13	As compared to pure atomic orbitals the hybrid orbitals have.	A. Low energy B. High energy C. Same energy D. None of these
14	How many sigma and pi bonds are present in C ₂ H ₂ .	A. one sigma and two pi B. two pi and one sigma C. Two pi and three sigma D. Three pi and two sigma

15	In which molecule all atoms are coplanar.	A. CH ₄ B. BF ₃ C. NH ₃ D. PH ₃
16	In which one of the following pairs do the molecule have similar shape.	A. BF ₃ and AlCl ₃ B. CO ₂ and H ₂ O C. CH ₄ and PH ₃ D. NH ₃ and BCl ₃
17	Both CH ₄ and NH ₃ are four electron pair system the angles of CH ₄ and NH ₃ are 109.5° and 107.5 ° respectively. This deviation is due to.	A. Hydrogen bonding in ammonia B. Lone pair attraction C. Lone pair occupy more space and repel to other bond pairs D. Lone pair lone pair repulsion