

Chemistry Fsc Part 1 Chapter 4 Online Test

Sr	Questions	Answers Choice
1	The process in which liquid can be made to boil at low temperature is known asdistillation	A. Simple B. Thermal C. Steam D. Vacuum
2	In order to mention the B.P of water at 110 °C the external pressure should be.	A. Between 760 torr and 1200 torr B. Between 200 torr and 760 torr C. 765 torr D. Any value of pressure
3	When water freezes, its volume increase.	A. 12% B. 9% C. 15% D. 18%
4	Which of the given has hydrogen bonding.	A. CH ₄ B. CCl ₄ C. NH ₃ D. NaCl
5	Acetone and chloroform are soluble in each other due to.	A. Intermolecular hydrogen bonding B. Dipole dipole interaction C. Instantaneous dipoles D. All of the above
6	London dispersion force are the only forces present among the.	A. Molecules of water in liquid state B. Atoms of helium in gaseous state at high temperature C. Molecule of solid iodine D. Molecules of hydrogen chloride gas
7	Dipole-dipole forces are present among.	A. Molecules of Iodine B. Atoms of Neon in gaseous state C. Chloroforms' molecules D. CCl ₄ molecules
8	The number of Na ⁺ ions which surround each Cl ⁻ ion in the NaCl crystal lattice is	A. 8 B. 12 C. 6 D. 4
9	NaCl is face centered cubic structure. The Na ion at the face of the unit cell is shared by	A. 2-unit cells B. 4-unit cells C. Only one unit cell D. 8-unit cells
10	How many allotropic forms are present in carbon	A. Two B. Three C. Four D. Five
11	Which one of the following substances is not amorphous	A. Polymer B. Rubber C. Glass D. AgNO ₃
12	Diamond is a bad conductor of electricity because	A. It has a tight structure B. It has a high density C. There are no free electrons present in the crystal of diamond to conduct electronics D. None of these
13	The molecules of CO ₂ in dry ice form the	A. Ionic crystal B. Covalent crystals C. Molecular crystals D. Any type of crystals
14	One of the following liquids has lowest vapour pressure at 32°C. Indicate that liquid	A. Ether B. Chloroform C. Ethanol D. Water
		A. Fractional distillation

15	The distillation of a solution under reduced pressure is called	B. Destructive distillation C. Distillation D. Vacuum distillation
16	Which of the following can form H-bonds	A. NH_3 B. C_2H_6 C. NaCl D. CHCl_3
17	The long chains of amino acids are coiled about one another onto a spiral by	A. Ionic bond B. Van der Waals forces C. Hydrogen bonding D. Overlapping of orbitals
18	The polarizabilities of elements mostly increase down the group due to the reason that	A. The atomic numbers increase B. Number of protons increase C. Number of shells increase along, with increase of shielding effect D. The behavior of the elements remain the same
19	The repulsion of electronic clouds of the molecules are responsible for the attractive forces among the molecules. These forces are	A. Dipole-induced dipole forces B. Ion-dipole forces C. Instantaneous dipole-induced dipole forces D. Dipole-dipole forces
20	In order to mention the B.P. of water at 110°C , the external pressure should be	A. Between 760 torr and 1200 torr B. Between 200 torr and 760 torr C. 760 torr D. Any value of pressure
21	When water freezes at 0°C , its density decreases due to	A. Cubic structure of ice B. Empty spaces present in the structure of ice C. Change of bond lengths D. Change of bond angles
22	NH_3 shows a maximum boiling point among the hydrides of V-A group elements due to	A. Very small size of nitrogen B. Lone pair electrons present on Nitrogen C. Enhanced electronegative character of Nitrogen D. Pyramidal structure of NH_3
23	Acetone and chloroform are soluble in each other due to	A. Intermolecular hydrogen bonding B. Dipole-dipole interaction C. Instantaneous dipoles D. All of the above
24	London dispersion forces are the only forces present among the	A. Molecules of water in liquid state B. Atoms of helium in gaseous state at high temperature C. Molecules of solid iodine D. Molecules of hydrogen chloride gas