Electrochemistry

Sr	Questions	Answers Choice
1	Which element has Eo = 0.00 V?	A. H+ B. H2 C. SHE D. All of the above
2	In the Hall Heroult process, aluminum is obtained by	A. Chemical reaction B. Thermal decomposition C. Elelctrolysis of alumina D. Roasting and leaching
3	In an electrolysis exprement if a charge 96,500 Coluombs is passed trough a solution, the amount of substance liberated or deposited at the electrode is directly relatd to.	A. Mass numebr of the ion B. One mole of electrons being transferred C. Avogadro's nmber of ions being discharged D. Standrd electrode potentil of the metal iion
4	Which of the following changes wold typically lead to an incresse in he rate of electrolysis.	A. Decreasing ghe concentration of the electrolyte B. Incressing the distance between the electrodes C. Decressing the surface area of the electrolytic cell D. Increasing the current passed though the electroytic cell
5	If 1 Faraday of elecricity is passed the mass deposited equals.	A. 1 gram equivalent B. 1 gram C. 1 mole D. 1 atm
6	The principle of measuring DO by Wrinkler's Method is based on.	A. Iodimetry B. Iodometry C. Acid Bse titration D. Complexometry
7	Electrolysis of CuSO4 using copper eletrodes results in	A. Increase in eletolyte concentration B. Decrease in Cu2+ concentration C. No change in electrolyte composition D. Formation of new compound
8	If Zn -Cu galvanic cell works ideally after complete discharge, both comparatments will have	A. CuSO4 Solution B. Zn SO4 Solution C. Cu Ions D. Zn Solid
9	What is the unit of cell potanetial.	A. ampere B. Volt C. Ohm D. Farad
10	Which metal is the best reducing agent.	A. Cu B. K C. Zn D. Fe
11	Electroyte in electroplating of silver	A. AgNo3 B. NaCl C. CuSO4 D. HNO3
12	A Daniell cell produces electricity through	A. Electrolysis B. Radioactivity C. Spontaneous redox reaction D. Endothermic reaction
13	Electrochemical cells convert chemical energy into	A. Electrila energy B. Heat C. Light D. Nuclear energy
14	Electrochemical sereis helps in predicting.	A. Rate of reaction B. Type of bond

		C. Dirction of redox reactions D. Heat released
15	Cell potential is the differecne between	A. Temperature and pressure B. Concentration of ions C. Electrode potentials of chathode and anode D. Mass of electrodes
16	Oxidation numebrof oxygen in H_2O	A. 1 B. -2 C. 0 D. +1
17	In a galvanic cell, the anode is	A. Negaive and site of oxidation B. Postive and site of oxidation C. Negative and site of reduction D. Postive and site of reduction
18	The more positive the standard reduction potential	A. The stronger the reducing agent B. The weaker the oxidizing agent C. The stronger the oxidizing agent D. No effect
19	A positive value for the standared electrode potential of a metal ion metal half cell indicate that	A. The metal is stong reducing agent B. The metal ion is redily oxidized C. The metal ion is redily reduced D. The metal will readily displace hydrogen from dilute acids.
20	A more genative Eo value means the electrode	A. Is a stong oxidizing agent B. Is a strong reducing agent C. Is neurtal D. Has high electronegagtivity