

Electrochemistry

Sr	Questions	Answers Choice
1	If salt bridge is not used, between two half cells in a Galvanic cell, then the voltage.	A. Decrease slowly B. Decrease rapidly C. Does not change D. Drops to zero
2	The reducing agent in a redox reaction	A. Gains electrons B. Loss electrons C. Accepts protons D. Donates protons
3	Standard hydrogen electrode is assigned potential of.	A. 1.0 V B. 0.5 V C. 0.0 V D. -1.0 V
4	The more positive the standard reduction potential	A. The stronger the reducing agent B. The weaker the oxidizing agent C. The stronger the oxidizing agent D. No effect
5	Electrochemical series helps in predicting.	A. Rate of reaction B. Type of bond C. Direction of redox reactions D. Heat released
6	Electrolysis of CuSO_4 using copper electrodes results in	A. Increase in electrolyte concentration B. Decrease in Cu^{2+} concentration C. No change in electrolyte composition D. Formation of new compound
7	The activity series of metals arranges metals in order of their.	A. Atomic mass B. Density C. Ease of oxidation D. Ease of reduction
8	Cell potential is the difference between	A. Temperature and pressure B. Concentration of ions C. Electrode potentials of cathode and anode D. Mass of electrodes
9	Which of the following will react with HCl to liberate H_2 gas	A. Ag B. Mg C. Au D. Cu
10	In a galvanic cell, the anode is	A. Negative and site of oxidation B. Positive and site of oxidation C. Negative and site of reduction D. Positive and site of reduction
11	Electrochemical equivalent is.	A. Mass per mole B. Mass per coulomb C. Charge per second D. Current per mass
12	A salt bridge is usually made from	A. Metallic wires B. Porous glass C. Inert salt solution in agar agar D. Hydrochloric acid
13	In an electrolysis experiment if a charge 96,500 Coulombs is passed through a solution, the amount of substance liberated or deposited at the electrode is directly related to.	A. Mass number of the ion B. One mole of electrons being transferred C. Avogadro's number of ions being discharged D. Standard electrode potential of the metal ion
14	Which element has $E^\circ = 0.00 \text{ V}$?	A. H^+ B. H_2 C. SHE D. H^+/H_2

		D. All of the above
15	Electrolysis is used in the extraction of	A. Silver B. Gold C. Aluminum D. Mercury
16	Electrolysis of brine produces.	A. Hydrogen, oxygen, NaOH B. Sodium, Chlorine, H ₂ O C. Hydrogen, Chlorine, NaOH D. NaOH, Cl ₂ , CO ₂
17	A Daniell cell produces electricity through	A. Electrolysis B. Radioactivity C. Spontaneous redox reaction D. Endothermic reaction
18	If E _o cell is +1.10 V, the reaction is	A. Spontaneous B. Non spontaneous C. At equilibrium D. Endothermic only
19	If Zn-Cu galvanic cell works ideally after complete discharge, both compartments will have	A. CuSO ₄ Solution B. Zn SO ₄ Solution C. Cu ions D. Zn Solid
20	In redox reactions, oxidizing agents are	A. reduced B. Oxidized C. Always metals D. Always gases