

## Chemistry Fsc Part 1 Chapter 10 Online Test

Sr	Questions	Answers Choice
1	The voltage Nickel Cadmium cell is	A. 1 V B. 1.2 V C. 1.4 V D. 1.6 V
2	A single lead cell provides volts	A. 2 B. 4 C. 6 D. 8
3	The reduction potential of Zn is.	A. +0.76 V B0.34 B C. +0.34 V D0.76 V
4	Stronger the oxidizing agent greater is the	A. Oxidation potential B. Reduction potential C. Redox potential D. E.M.F of cell
5	If the salt bridge is not used between two half cells, then the voltage.	A. Decrease rapidly B. Decrease slowly C. Drops to zero D. Does not change
6	The cathodic reaction in the electrolysis of dil H2SO4, with pt electrode sis.	A. Reduction     B. Oxidation     C. Both oxidation and reduction     D. Neither oxidation nor reduction
7	If a strips of Cu metal is placed in a solution of FeSO4	<ul><li>A. Cu will be precipitated down</li><li>B. Fe is precipitated out</li><li>C. Cu and Fe both dissolve</li><li>D. No reaction takes palce</li></ul>
8	Oxidation number of Cr in a C <sub>2</sub> CrO <sub>4</sub> is	A. +2 B. +4 C. +6 D. +8
9	In H2SO4the oxidation number of 'S' is	A. +2 B. +6 C. +8 D. +4
10	Oxidation number of phosphorus in the compound is.	A. +3 B. +4 C. +5 D. +6
11	In silver oxide battery, the cathode is mad up of.	A. AgO B. Ag2O C. Ag2O3 D. Ag
12	The oxidation of O -atom in OF3 is.	A2 B. +2 C1 D. +1
13	The oxidation number of C in C12H22O11 is	A. Zero B 6 C. + 6 D. 12
14	Fuel cells are the means by which chemical energy may be converted into	A. Heat energy B. Magnetic energy C. Sound energy D. Electric energy
15	Electrochemical series is the arrangement of the electrodes in	A. Increasing order of reduction potentials     B. Decreasing order of reduction potentials

		C. Increasing order of oxidation reduction potential  D. There is no fixed arrangement
16	Electromotive force of the cell is the	A. Difference of two electrode potentials B. May be sum or the difference of two electrode potentials C. Sum of two electrode potential D. Depends upon the nature of the cell
17	The electrode reaction of a voltaic cell can be reversed when	A. Concentrations of solutions are changed B. Temperature is increased C. Electrodes are interchanged D. Electric circuit is employed to supply the source of electricity