

Electrochemistry

Sr	Questions	Answers Choice
1	The principle of measuring DO by Winkler's Method is based on.	A. Iodimetry B. Iodometry C. Acid Base titration D. Complexometry
2	The oxidation number of sulfur in H ₂ SO ₄ is	A. +2 B. +4 C. +6 D. +8
3	In electrolysis, reduction always occurs at	A. Anode B. Cathode C. Salt bridge D. Electrolyte
4	Farada's First Law states mass deposited is directly proportional to that m	A. Time B. Voltage C. Resistance D. Quantity of electricity passed
5	Electrolysis of CuSO ₄ using copper electrodes results in	A. Increase in electrolyte concentration B. Decrease in Cu ²⁺ concentration C. No change in electrolyte composition D. Formation of new compound
6	Electrolyte in electroplating of silver	A. AgNO ₃ B. NaCl C. CuSO ₄ D. HNO ₃
7	Standard electrode potential is measured under	A. 100 °C and 1 atm B. 0 °C and 1 M concentration C. 25 °C and 1 M concentration D. 100 °C and 1 M concentration
8	If E _o cell is +1.10 V, the reaction is	A. Spontaneous B. Non spontaneous C. At equilibrium D. Endothermic only
9	Electrolysis is the process of	A. Chemical energy into electrical B. Electrical energy into chemical C. Heat energy into mechanical D. Mechanical energy into chemical
10	In a galvanic cell, the anode is	A. Negative and site of oxidation B. Positive and site of oxidation C. Negative and site of reduction D. Positive and site of reduction
11	Which component maintains electrical neutrality in an electrochemical cell	A. Salt bridge B. Voltmeter C. Electrolyte D. Electrodes
12	The number of coulombs required to deposit 1 mole of Ag	A. 96500 C B. 193000 C C. 1 C D. 241250 C
13	Standard hydrogen electrode is assigned potential of.	A. 1.0 V B. 0.5 V C. 0.0 V D. -1.0 V
14	Which of the following changes would typically lead to an increase in the rate of electrolysis.	A. Decreasing the concentration of the electrolyte B. Increasing the distance between the electrodes C. Decreasing the surface area of the electrolytic cell D. Increasing the current passed through the electrolytic cell

15	If salt bridge is not used, between two half cells in a Galvanic cell, then the voltage.	A. Decrease slowly B. Decrease rapidly C. Does not change D. Drops to zero
16	Electrolysis of brine produces.	A. Hydrogen, oxygen, NaOH B. Sodium, Chlorine, H ₂ O C. Hydrogen, Chlorine, NaOH D. NaOH, Cl ₂ , CO ₂
17	Oxidation number of oxygen in H ₂ O	A. 1 B. -2 C. 0 D. +1
18	The reducing agent in a redox reaction	A. Gains electrons B. Loss electrons C. Accepts protons D. Donates protons
19	Which metal is the best reducing agent.	A. Cu B. K C. Zn D. Fe
20	Which metal will displace Cu from CuSO ₄ Solution.	A. Zn B. Ag C. Hg D. Pb