Electrochemistry

Sr	Questions	Answers Choice
1	Redox reactions always involve.	A. Gain of protons B. Transfer of electrons C. Loss ofneutrons D. Nuclear change
2	In a galvanic cell, the anode is	A. Negaive and site of oxidation B. Postive and site of oxidation C. Negative and site of reduction D. Postive and site of reduction
3	Cell potential is the differecne between	A. Temperature and pressure B. Concentration of ions C. Electrode potentials of chathode and anode D. Mass of electrodes
4	Which species is educed in the followig rection ? Zn + Cu --------- Zn	A. Cu2+ B. Zn C. Cu D. Zn3+
5	What is the unit of cell potanetial.	A. ampere B. Volt C. Ohm D. Farad
6	Electroysis is used in the extration of	A. Silver B. Gold C. Aluminum D. Mercury
7	The reducing agent in a redox reaction	A. Gains electrons B. Loss electrons C. Acceptw protons D. Donates protons
8	Which metal is the best reducing agent.	A. Cu B. K C. Zn D. Fe
9	The principle of measuring DO by Wrinkler's Method is based on.	A. Iodimetry B. Iodometry C. Acid Bse titration D. Complexometry
10	The reaction at the cathode is alwyas	A. Oxidation B. Neutral C. Reduction D. Ionization
11	Electrolysis of CuSO4 using copper eletrodes results in	A. Increase in eletolyte concentration B. Decrease in Cu2+ concentration C. No change in electrolyte composition D. Formation of new compound
12	Electrolysis is the process of	A. Chemical energy into electrical B. Electrical energy into chemical C. Heat energy into mechanical D. Mechanical energy into chemical
13	Electrochemical cells convert chemical energy into	A. Electrila energy B. Heat C. Light D. Nuclear energy
14	In the Hall Heroult process, aluminum is obtained by	A. Chemical reaction B. Thermal decomposition C. Elelctrolysis of alumina D. Roasting and leaching
		A. 96500 C

15	The numebr of coulombs required to deposit 1 mole of Ag	B. 193000 C241250 C C. 1 C D. 241250 C
16	Electrochemical sereis helps in predicting.	A. Rate of reaction B. Type of bond C. Dirction of redox reactions D. Heat released
17	Oxidation numebrof oxygen in H_2O	A. 1 B. -2 C. 0 D. +1
18	If salt bridge in not used, between twohalf cells in a Galvanic cell, than the voltage.	A. Decrease slowly B. Decrease repidly C. Does not change D. Drope to zero
19	If 1 Faraday of elecricity is passed the mass deposited equals.	A. 1 gram equivalent B. 1 gram C. 1 mole D. 1 atm
20	Which of the following is not a use of electrochemistry	A. Manufacture of explosives B. Electroplating C. Extractio of metals D. Electrorefining