

Electrochemistry

Sr	Questions	Answers Choice
1	Electrochemical cells convert chemical energy into	A. Electric energy B. Heat C. Light D. Nuclear energy
2	Electrochemical series helps in predicting.	A. Rate of reaction B. Type of bond C. Direction of redox reactions D. Heat released
3	Which of the following is not a use of electrochemistry	A. Manufacture of explosives B. Electroplating C. Extraction of metals D. Electrorefining
4	In an electrochemical cell, which ion migrates to cathode through the salt bridge.	A. Anions B. Cations C. Electrons D. Protons
5	Electrolysis of brine produces.	A. Hydrogen, oxygen, NaOH B. Sodium, Chlorine, H ₂ O C. Hydrogen, Chlorine, NaOH D. NaOH, Cl ₂ , CO ₂
6	Electrochemical equivalent is.	A. Mass per mole B. Mass per coulomb C. Charge per second D. Current per mass
7	The principle of measuring DO by Winkler's Method is based on.	A. Iodimetry B. Iodometry C. Acid Base titration D. Complexometry
8	For a redox reaction to be spontaneous, ΔG should be	A. Positive B. Negative C. Zero D. Constant
9	One Faraday is equal to	A. 1 C B. 96500 C C. 1 J D. 96500 J
10	The oxidation number of sulfur in H ₂ SO ₄ is	A. +2 B. +4 C. +6 D. +8
11	In redox reactions, oxidizing agents are	A. reduced B. Oxidized C. Always metals D. Always gases
12	The reaction at the cathode is always	A. Oxidation B. Neutral C. Reduction D. Ionization
13	Electroplating involves.	A. Electrolysis for coating B. Spontaneous oxidation C. Redox without electrodes D. Thermolysis
14	Cell potential is the difference between	A. Temperature and pressure B. Concentration of ions C. Electrode potentials of cathode and anode D. Mass of electrodes
15	The activity series of metals arranges metals in order of their.	A. Atomic mass B. Density C. Ease of oxidation D. Electronegativity

D. Ease of reduction

16 In electrolysis, reduction always occurs at

- A. Anode
- B. Cathode
- C. Salt bridge
- D. Electrolyte

17 The metal deposited at cathode during electrolysis of NaCl is.

- A. Cl₂
- B. Na
- C. H₂
- D. Cu

18 The cathode in an electrolytic cell is.

- A. Positive electrode
- B. Negative electrode
- C. Neutral electrode
- D. None

19 Which metal will displace Cu from CuSO₄ Solution.

- A. Zn
- B. Ag
- C. Hg
- D. Pb

20 Electrons in a galvanic cell flow from

- A. Anode to cathode
- B. Cathode to anode
- C. Salt bridge to cathode
- D. Electrolyte to anode