

## Periodic Table and Periodic Properties

Sr	Questions	Answers Choice
1	Peroxides are those in which oxidation state of oxygen is	<p>A. <math>-1</math></p> <p>B. <math>+2</math></p> <p>D. <math>-2</math></p> <p>E. <math>-1/2</math></p>
2	Which element is used in S.O.S flares and in firework.	<p>A. Mg Powder</p> <p>B. Na powder</p> <p>C. <math>\text{CaCO}_3</math></p> <p>D. <math>\text{KO}_2</math></p>
3	The term "Halogen" mean	<p>A. Ash former</p> <p>B. Salt Former</p> <p>C. Copper giver</p> <p>D. Sulphur giver</p>
4	Select which is not alkaline earth metal	<p>A. Ba</p> <p>B. Be</p> <p>C. Sr</p> <p>D. CO</p>
5	Group 16 elements are also called.	<p>A. Chalcogens</p> <p>B. Pnictogens</p> <p>C. Alkali metals</p> <p>D. Halogens</p>
6	L- Pauling developed a scale for	<p>A. Ionization energy</p> <p>B. Electron affinity</p> <p>C. Electronegativity</p> <p>D. None</p>
7	Tendency to lose electrons called	<p>A. metallic character</p> <p>B. Electron affinity</p> <p>C. Ionization energy</p> <p>D. Electronegativity</p>
8	Which of the following atoms show more than one oxidation states.	<p>A. Phosphorous</p> <p>B. Magnesium</p> <p>C. Sodium</p> <p>D. Aluminium</p>
9	Na burns with oxygen to give.	<p>A. Golden yellow flame</p> <p>B. Golden brown flame</p> <p>C. Crimson red flame</p> <p>D. Intense white flame</p>
10	Which oxidation state will not shown by S	<p>A. <math>-2</math></p> <p>B. <math>+4</math></p> <p>C. <math>+8</math></p> <p>D. <math>+2</math></p>
11	Least electronegative element is.	<p>A. Cs</p> <p>B. H</p> <p>C. Li</p> <p>D. He</p>
12	Modern periodic table provides information about.....elements.	<p>A. 100</p> <p>B. 118</p> <p>C. 110</p> <p>D. 130</p>
13	By 1700 A.D only----- elements were recognized.	<p>A. 3</p> <p>B. 6</p> <p>C. 12</p> <p>D. 10</p>
14	Which is the correct trend in variation of electronegativity along a period of the periodic table.	<p>A. It decreases from left to right across a period</p> <p>B. It increases from left to right across a period</p> <p>C. It remains constant</p> <p>D. It has no definite trend</p>
15	2nd Value of electron affinity	<p>A. Will be negative always</p> <p>B. Will always be positive</p>

		<p>C. <math>\leq</math> Equal to 1st value</p> <p>D. <math>&gt;</math> Double than 1st Value</p>
16	Metallic character increases	<p>A. Downward and forward from left to right</p> <p>B. Upward and forward from left to right</p> <p>C. Downward and backward from right to left</p> <p>D. Upward and backward from right to left</p>
17	Which property increases as you go down a group in the periodic table?	<p>A. Atomic radius</p> <p>B. Electron affinity</p> <p>C. Electronegativity</p> <p>D. Ionization Energy</p>
18	Acidic oxides are	<p>A. Metallic oxides</p> <p>B. Non Metallic oxides</p> <p>C. Metalloid oxides</p> <p>D. All</p>
19	What is the oxidation state of sulfur in the sulfate ion $\text{SO}_4^{2-}$	<p>A. +4</p> <p>B. +6</p> <p>C. +2</p> <p>D. 0</p>
20	Which set of the following conditions results in higher ionization energy.	<p>A. Smaller atom and greater nuclear charge</p> <p>B. Smaller atom and smaller nuclear charge</p> <p>C. Larger atom and greater nuclear charge</p> <p>D. Large atom and the smaller nuclear charge</p>