

Periodic Table and Periodicity

Sr	Questions	Answers Choice
1	Second and third periods are called	A. 1st transition series B. Normal periods C. 2nd transition series D. 3rd transilition series
2	Which of the following groups contain alkaline earth metals.	A. I A B. II A C. VII A D. VIII A
3	Period number tells about the	A. No. of valence electrons B. No. of electronic shells C. Both a and b D. None of the above
4	The elements of group 1 and 2 are placed in which block	A. s B. p C. d D. f
5	How many electrons are present in the valence shell of group 1 elements.	A. 1 B. 2 C. 3 D. 4
6	Which are good conductor of heat and electricity	A. Metals B. Non metals C. Metalloids D. All of these
7	Electron affinity of fluorine in kJ mol^{-1} is	A. -328 B. 328 C. -330 D. -340
8	Which halogen element exists in a liquid state at room temperature	A. Bromine B. Chlorine C. Fluorine D. Iodine
9	Which one of the following halogens has lowest electronegativity	A. Iodine B. Chlorine C. Fluorine D. Bromine
10	Actinide series starts after the element	A. Actinium B. Lanthanum C. Osmium D. Silver
11	Which of the following element belongs to VIII A.	A. Xe B. Mg C. Br D. Na
12	How many periods are present in the modern periodic table.	A. 7 B. 8 C. 10 D. 12
13	How many electrons can nitrogen accept in its outermost shell.	A. 2 B. 3 C. 4 D. 5
14	The minimum amount of energy which is required to remove an electron from valence shell of the gaseous state of an atom is called.	A. Potential energy B. Ionization energy C. Electron affinity D. Electronegativity
15	Main group elements are arranged ingroups.	A. 7 B. 6 C. 8 D. 10

16	Atomic number of actinium is	A. 57 B. 60 C. 89 D. 80
17	With the increase of atomic numebr , the number of electron in an atom also.	A. Decreases B. First increases then decreases C. Increases D. None of the above
18	Atomic number of lanthanum is	A. 57 B. 58 C. 59 D. 60
19	The unit of ionization energy is	A. nm and pm B. kJ mol^{-1} C. Pascal D. Newton
20	Point out the incorrect statement about electron affinity	A. It decreases in a period B. It decreases in a group C. It is measure din kJmol^{-1} D. None of these