

## Periodic Table and Periodicity

Sr	Questions	Answers Choice
1	The electronegativity of carbon is	A. 2.5 B. 2.0 C. 1.0 D. 4.0
2	Electronegativity of oxygen is.	A. 3,1 B. 3,3 C. 3,4 D. 3,2
3	Which one of the followign halogens has highest electronegativity	A. Iodine B. Fluorine C. Chlorine D. Bromine
4	The ability of an atom to attract the shared pair of electrons towards itself in a molecule is called	A. Ionization energy B. Electronegativity C. Shielding effect D. Electron affinity
5	Electron affinity of fluorine in $\text{kJ mol}^{-1}$ is	A. -328 B. 328 C. -330 D. -340
6	Unit of electron affinity is.	A. $\text{kJ mol}^{-1}$ B. $\text{kJ mol}$ C. pm D. Newton
7	When we move from left to right in a period, ionization energy.	A. No effect B. Decreases C. Increases D. None of these
8	When we move top to bottom in group, ionization energy.	A. Increases B. No effect C. Decreases D. None of these
9	The unit of ionization energy is	A. nm and pm B. $\text{kJ mol}^{-1}$ C. Pascal D. Newton
10	The minimum amount of energy which is required to remove an electron from valence shell of the gaseous state of an atom is called.	A. Potential energy B. Ionization energy C. Electron affinity D. Electronegativity
11	When we move from top to bottom in a group atomic size.	A. Decrease B. Increases C. First increases then decreases D. None of the above
12	When we move from left to right in a period, atomic number	A. Decreases B. Increases C. First increases then decreases D. None of the above