

Periodic Table and Periodicity

Sr	Questions	Answers Choice
1	The elements of group 3 to 12 are called.	A. Normal elements B. Halogens C. Noble gases D. Transition elements
2	The element having less value of ionization energy and less value of electron affinity is likely to belong to.	A. Group 1 B. Group 13 C. Group 16 D. Group 17
3	All the elements of Group II A are less reactive than alkali metals. This is because these elements have.	A. Decreased nuclear charge B. Similar electronic configuration C. High ionization energies D. Relatively greater atomic size.
4	Point out the incorrect statement about electron affinity	A. It decreases in a period B. It decreases in a group C. It is measured in kJ mol^{-1} D. None of these
5	17th Group of elements contain electrons in their outer most shell	A. 4 B. 5 C. 7 D. 6
6	How many periods are present in the modern periodic table	A. 7 B. 8 C. 10 D. 12
7	Electron affinity of fluorine in kJ mol^{-1} is	A. -328 B. 328 C. -330 D. -340
8	Which of the following is a metal	A. Magnesium B. Carbon C. Hydrogen D. Nitrogen
9	Elements of Group 1 are called.	A. Alkali Metals B. Alkali earth metals C. Transition metals D. Halogen
10	Atomic number of lanthanum is	A. 57 B. 58 C. 59 D. 60
11	How many elements are present in 6th period.	A. 2 B. 8 C. 18 D. 32
12	Main group elements are arranged ingroups.	A. 7 B. 6 C. 8 D. 10
13	Which of the following groups contain alkaline earth metals.	A. I A B. II A C. VII A D. VIII A
14	Lanthanide series starts after the element	A. Osmium B. Actinium C. Lanthanum D. None of these
15	Metals lose their electrons easily because.	A. They are electronegative B. They have electron affinity C. They are electropositive D. Good conductors of heat

16	Which element is the most reactive element?	A. Florine B. Oxygen C. Chlorine D. Nitrogen
17	Mark the icorrect statement about ionization energy.	A. It is measured in kJmol-1 B. It is absorption of energy C. It decreased in a period D. It decrease in a group
18	How many elements are present in 1st period.	A. 1 B. 2 C. 8 D. 18
19	Elements re classified into four blocks depending upon	A. Shell B. Atomic mass C. Sub -Shell D. Atomic Number
20	When we mvove from left to right in a period, ionization energy.	A. No effect B. Decreases C. Increases D. None of these
