Energetics

Sr	Questions	Answers Choice
1	---------- acts a catalyst promoting the breakdown of ozone.	A. I2 B. Br2 C. Cl2 D. None
2	Bond formation energy of one O-H bond is.....	A. 488 kJ/mol B. 484 kJ/mol C. 486 kJ/mol D. 489 kJ/mol
3	The part of the universe that we want to focus our attention called.	A. Surrounding B. Energy C. System D. Both a and b
4	The enthalpy of reaction C+O2 ------ CO2	A. -571.6 kJ B. -393.5 kJ C. +53.8 kJ D. -110.5 kJ
5	The enthalpy of reaction 2H2 +O2-----2H2O	A. -571.6 kJ B. -110.5 kJ C. -393.5 kJ D. +53.8 kJ
6	Which is not produced in an aerobic respiration.	A. Carbon dioxide B. Lactic acid C. Water D. Energy
7	---------- of the energy used by traditional electric bulb is wasted in producing heat.	A. 60% B. 50% C. 70% D. 90%
8	No reaction occurs if the energy of reacting particles.....activation energy.	A. Lower than B. Greather than C. Nearest to D. Equal to
9	When old bonds are broken, the energy is.	A. Release B. Remain same C. Consume D. None of these
10	Activation energy of a chemical reaction must be..... the everage kinetic energy of reacting molecules.	A. Equal to B. Grether than C. Lower than D. None of these
11	Formation of NO is	A. Exothrmic B. Endothermic C. No Heat Change D. None of these
12	Aerobic respiration releases.....energy than anaerobic respiration.	A. Equal B. Less C. More D. None of these
13	----------acts are reserve energy sources.	A. Enzymes B. Vitamins C. Proteins D. Lipids
14	All chemical reaction involves.	A. Enzymes B. Catalyst C. Energy changes D. All of these
15	Bond dissociation for O2 is	A. 505 kJ/mol B. 705 kJ/mol C. 605 kJ/mol D. 498 kJ/mol

16	The word energy is used in physics ofr the firt time.	A. 1902 B. 1858 C. 1805 D. 1802
17	If the Delta H value is negative then reaction witll be	A. Endotermic B. Exothermic C. May or may not be exothermic or endothermic D. None of these
18	When new bonds ae formed, the energy is	A. Consume B. Remain same C. Release D. None of these
19	The enthalpy of reaction H2+I2--- 2HI	A. -571.6 k J B. +53.8 kJ C. 11 kJ D. -393.5 kJ
20	When NaOH and HCl are mixed the temperature increases. The reaction	A. Exothermic with a negative enthalpy chagne. B. Endothermic with a positive enthaly change. C. Endothermic with a negatie enthalpy change D. Exothermic with a positive enthealpy change