

Stoichiometry

Sr	Questions	Answers Choice
1	Mass of 3 moles of oxygen atoms is.	A. 64 g B. 16 g C. 32 g D. 48 g
2	A necklace has 6 g of diamonds in it. What are the number of carbon atoms in it?	A. 3.01×10^{23} B. 1.003×10^{23} C. 12.04×10^{23} D. 6.02×10^{23}
3	Which term is the same for one mole of oxygen and one mole of water.	A. atoms B. mass C. Molecules D. Volume
4	Limes is another name of	A. Sodium hydroxide B. Sodium Carbonate C. Calcium Carbonate D. Silicon dioxide
5	Avogadro was a scientist	A. Italian B. Greek C. German D. African
6	Without stoichiometry which industry cannot exist.	A. Metal B. Petroleum C. Leather D. Chemical
7	Which of the following is insoluble salt	A. KCl B. AgCl C. NaCl D. CaCl ₂
8	Formula of common salt is	A. NaCl B. AgCl C. LiCl D. KCl
9	Which compound has same molecular and empirical formula.	A. C ₆ H ₁₂ O ₆ B. H ₂ O ₂ C. H ₂ O D. C ₆ H ₆
10	Empirical formula of acetic acid (CH ₃ COOH) is	A. CHO B. CH C. CH ₂ O D. None of these
11	Value of Avogadro's number is.	A. 6.6×10^{20} B. 6.02×10^{23} C. 6.00×10^{24} D. 1.32×10^{23}
12	Empirical formula of hydrogen peroxide.	A. HO B. CO C. CHO D. CH
13	Stoichiometric calculations are used to prepare.	A. Soaps B. Shampoo C. Perfumes D. All of these
14	Empirical formula of sand is.	A. SiO ₂ B. SiO ₃ C. SiO ₄ D. SiO
15	How many moles of molecules are there in 16 g oxygen.	A. 0.05 B. 0.1 C. 0.5 D. 1

16	Formula mass of K_2SO_4 is.	<p>A. 174 amu B. 110 amu C. 180 amu D. 145 amu</p>
17	How many atoms are present in one gram of H_2O ?	<p>A. 1002×10^{23} atom B. 6.022×10^{23} atom C. 0.334×10^{23} atom D. 2.004×10^{23} atom</p>
18	How many atomic mass units (amu) are there in one gram.	<p>A. 1 amu B. 6.022×10^{23} C. 10 amu D. 6.022×10^{22}</p>
19	Which one of the following compounds will have the highest percentage of the mass of nitrogen?	<p>A. N_2H_4 B. $CO(NH_2)_2$ C. NH_3 D. NH_2OH</p>
20	1- gram atom of carbon contain how many moles.	<p>A. 1 mole B. 2 mole C. 6 moles D. 12 moles</p>