

Stoichiometry

Sr	Questions	Answers Choice
1	Empirical formula of hydrogen peroxide.	A. HO B. CO C. CHO D. CH
2	Which compound has same molecular and empirical formula.	A. C ₆ H ₁₂ O ₆ B. H ₂ O C. H ₂ O D. C ₆ H ₆
3	The mass number of sodium is.	A. 19 B. 31 C. 27 D. 23
4	Value of Avogadro's number is.	A. 6.6×10^{-20} B. 6.02×10^{23} C. 6.00×10^{24} D. 1.32×10^{23}
5	Limes is another name of	A. Sodium hydroxide B. Sodium Carbonate C. Calcium Carbonate D. Silicon dioxide
6	A necklace has 6 g of diamonds in it. What are the number of carbon atoms in it?	A. 3.01×10^{23} B. 1.003×10^{23} C. 12.04×10^{23} D. 6.02×10^{23}
7	The molecular mass of H ₂ SO ₄ is.	A. 9.8 g B. 98 amu C. 98 g D. 9.8 amu
8	Which of the following is insoluble salt	A. KCl B. AgCl C. NaCl D. CaCl ₂
9	Without stoichiometry which industry cannot exist.	A. Meta B. Petroleum C. Leather D. Chemical
10	A compound with chemical formula Na ₂ CX ₃ has formula mass 106 amu. Atomic mass of the element X is.	A. 16 B. 23 C. 12 D. 106
11	Mass of 3 moles of oxygen atoms is.	A. 64 g B. 16 g C. 32 g D. 48 g
12	How many number of moles are equivalent to 8 gram of CO ₂ .	A. 0.18 B. 0.15 C. 0.24 D. 0.21
13	How many moles are there in 25 g of H ₂ SO ₄ ?	A. 0.765 Moles B. 0.255 moles C. 0.4 moles D. 0.51 moles
14	Stoichiometric calculators are used to prepare.	A. Soaps B. Shampoo C. Perfumes D. All of these
15	How many atoms of carbon are present in one molecule of glucose.	A. 11 B. 22 C. 12 D. 6

16	What mass of 95% CaCO ₃ will be required to neutralize 50 cm ³ of 0.5 M HCl solution.	A. 9.5 g B. 1.45 g C. 1.32 g D. 1.25 g
17	Empirical formula of Benzene is.	A. CH ₂ O B. CH C. C ₂ H D. CH ₂
18	Which one of the following compounds will have the highest percentage of the mass of nitrogen?	A. N ₂ H ₄ B. CO (NH ₂) ₂ C. NH ₃ D. NH ₂ OH
19	How many atoms are present in one gram of H ₂ O?	A. 1002×10^{23} atom B. 6.022×10^{23} atom C. 0.334×10^{23} atom D. 2.004×10^{23} atom
20	How many moles of molecules are there in 16 g oxygen.	A. 0.05 B. 0.1 C. 0.5 D. 1