Stoichiometry

Sr	Questions	Answers Choice
1	Empirical formula of acetic acid (CH3COOH) is	A. CHO B. CH C. CH2O D. None of these
2	A compound with chemicla formula Na_2CX_3 has formula mass 106 ami. Atomicmass of the elemetn X is.	A. 16 B. 23 C. 12 D. 106
3	Empirical formula of hydrogen peroxide.	A. HO B. CO C. CHO D. CH
4	What is the mass of Al in 204 g of aluminium oxide Al_2O_3	A. 26 g B. 54 g C. 108 g D. 27 g
5	Without stoichiometry which industry cannot exist.	A. Meta B. Petroleium C. Leather D. Chemical
6	How many moles are there in 25 g of H_2SO_4?	A. 0.765 Moles B. 0.255 moles C. 0.4 moles D. 0.51 moles
7	Stochiometric calculators are used to prepare.	A. Soaps B. Shampo C. Perfumes D. All of these
8	What mass of 95% CaCO3 will be requred to neutralize 50 cm 3 of 0.5 M HCl solution.	A. 9.5 g B. 1.45 g C. 1.32 g D. 1.25 g
9	A nceklaee has 6 g of diamonds in it . What are the numberof carbon atoms in it?	A. 3.01×10^{23} B. 1.003×10^{23} C. 12.04×10^{23} D. 6.02×10^{23}
10	How many atom of carbon are present in one molecule of glucose.	A. 11 B. 22 C. 12 D. 6
11	If one mole of carbon contains x atoms what is the number of atoms contained in 12 g of Mg.	A. 1.5 x B. 0.5 x C. x D. 2x
12	Which of the following is insoluble salt	A. KCl B. AgCl C. NaCl D. CaCl2
13	1 gram formula of NaCl contan is grams.	A. 100 g B. 58.5 g C. 32 g D. 49 g
14	The molecular mass of H2SO4 is.	A. 9.8 g B. 98 amu C. 98 g D. 9.8 amu
15	Which compund has same molecular and emiprical formula.	A. C6H12 O6 B. H2O2 C. H2O D. C6H6

16	How many atoic mass units (amu) are there in one gram.	A. 1 amu B. 6.022×10^{23} C. 10 amu D. 6.022×10^{22}
17	Whof the following represent sand?	A. NaCl B. CaCO3 C. H2O D. CH2O
18	Emperical Formula of Glucose is	A. CH2O B. CHO C. CHO2 D. C2HO
19	How many atoms are present in one gram of H2O?	A. 1002×10^{23} atom B. 6.022×10^{23} atom C. 0.334×10^{23} atom D. 2.004×10^{23} atom
20	The mass number of sodium is.	A. 19 B. 31 C. 27 D. 23