

Stoichiometry

Sr	Questions	Answers Choice
1	The mass of one molecule of water is.	A. 18 g B. 18 mg C. 18 kg D. 18 amu
2	If one mole of carbon contains x atoms what is the number of atoms contained in 12 g of Mg.	A. 1.5 x B. 0.5 x C. x D. 2x
3	What is the mass of 4 moles of hydrogen gas.	A. 1 g B. 1.008 g C. 8.064 g D. 4.032 g
4	Which one of the following compounds will have the highest percentage of the mass of nitrogen?	A. N ₂ H ₄ B. CO (NH ₂) ₂ C. NH ₃ D. NH ₂ OH
5	Empirical formula of Benzene is.	A. CH ₂ O B. CH C. C ₂ H D. CH ₂
6	Which term is the same for one mole of oxygen and one mole of water.	A. atoms B. mass C. Molecules D. Volume
7	Empirical formula of acetic acid (CH ₃ COOH) is	A. CHO B. CH C. CH ₂ O D. None of these
8	Formula mass of K ₂ SO ₄ is.	A. 174 amu B. 110 amu C. 180 amu D. 145 amu
9	Mass of 3 moles of oxygen atoms is.	A. 64 g B. 16 g C. 32 g D. 48 g
10	Formula of common salt is	A. NaCl B. AgCl C. LiCl D. KCl
11	Stoichiometric calculators are used to prepare.	A. Soaps B. Shampoo C. Perfumes D. All of these
12	How many atoms of carbon are present in one molecule of glucose.	A. 11 B. 22 C. 12 D. 6
13	What is the mass of Al in 204 g of aluminium oxide Al ₂ O ₃	A. 26 g B. 54 g C. 108 g D. 27 g
14	What mass of 95% CaCO ₃ will be required to neutralize 50 cm ³ of 0.5 M HCl solution.	A. 9.5 g B. 1.45 g C. 1.32 g D. 1.25 g
15	How many moles of molecules are there in 16 g oxygen.	A. 0.05 B. 0.1 C. 0.5 D. 1

16	How many atomic mass units (amu) are there in one gram.	A. 1 amu B. 6.022×10^{23} C. 10 amu D. 6.022×10^{22}
17	A compound with chemical formula Na_2CX_3 has formula mass 106 amu. Atomic mass of the element X is.	A. 16 B. 23 C. 12 D. 106
18	Which is the correct formula of calcium phosphide.	A. CaP B. Ca_3P_2 C. CaP_2 D. Ca_2P_3
19	1 gram formula of NaCl contains is grams.	A. 100 g B. 58.5 g C. 32 g D. 49 g
20	40 g of H_3PO_4 contains number of moles	A. 0.58 g B. 4.8 g C. 5.8 g D. 0.408 g