

Stoichiometry

Sr	Questions	Answers Choice
1	Formula mass of K_2SO_4 is.	A. 174 amu B. 110 amu C. 180 amu D. 145 amu
2	Whof the following represent sand?	A. NaCl B. $CaCO_3$ C. H_2O D. CH_2O
3	How many moels of molecules are there in 16 g oxygen.	A. 0.05 B. 0.1 C. 0.5 D. 1
4	Whcih law is obeyed in chemicla calculations?	A. Law of mass action B. Law of conservation of mass C. Law of definite proportion D. Both a and b
5	Empirical formula of hydrogen peroxide.	A. HO B. CO C. CHO D. CH
6	The molecular mass of H_2SO_4 is.	A. 9.8 g B. 98 amu C. 98 g D. 9.8 amu
7	Value of Avagadro's number is.	A. 6.6×10^{-20} B. 6.02×10^{23} C. 6.00×10^{24} D. 1.32×10^{23}
8	Empricial formula of sand is.	A. SiO_2 B. SiO_3 C. SiO_4 D. SiO
9	Avogadro was a scientist	A. Italian B. Greek C. Garman D. African
10	Mass of 3 moles of oxygen atoms is.	A. 64 g B. 16 g C. 32 g D. 48 g
11	Empirical formula of acetic acid (CH_3COOH) is	A. CHO B. CH C. CH_2O D. None of these
12	Empirical formula of Benzene is.	A. CH_2O B. CH C. C_2H D. CH_2
13	Numbr of hydrogen atoms present in 18 g of water.	A. $2 \times N_A$ B. $N_{sub>A</sub>}$ C. $3 \times N_{sub>A</sub>}$ D. $1/2 N_{sub>A</sub>}$
14	What mass of 95% $CaCO_3$ will be required to neutralize 50 cm ³ of 0.5 M HCl solution.	A. 9.5 g B. 1.45 g C. 1.32 g D. 1.25 g
15	A nceklaee has 6 g of diamonds in it . What are the numberof carbon atoms in it?	A. 3.01×10^{23} B. 1.003×10^{23} C. 12.04×10^{23} D. 6.02×10^{23}

16	The mass of one molecule of water is.	A. 18 g B. 18 mg C. 18 kg D. 18 amu
17	If one mole of carbon contains x atoms what is the number of atoms contained in 12 g of Mg.	A. 1.5 x B. 0.5 x C. x D. 2x
18	What is the mass of Al in 204 g of aluminium oxide Al_2O_3	A. 26 g B. 54 g C. 108 g D. 27 g
19	40 g of H_3PO_4 contains numebr of moles	A. 0.58 g B. 4.8 g C. 5.8 g D. 0.408 g
20	How many moles are there in 25 g of H_2SO_4 ?	A. 0.765 Moles B. 0.255 moles C. 0.4 moles D. 0.51 moles
