

## Chemical Bonding

Sr	Questions	Answers Choice
1	The force of attraction between water molecule is.	A. Ionic bonding B. Covalent bonding <b>C. Hydrogen Bonding</b> D. Coordinate Covalent bonding
2	Triple covalent bond involves how many electrons.	<b>A. Six</b> B. Four C. Eight D. Three
3	Covalent Bond involves the	A. Sharing of electrons B. Repulsion of electrons C. Acceptance of electrons D. Donation of electrons
4	Which ionic compound has the highest melting point.	A. RbCl B. KCl C. LiCl <b>D. NaCl</b>
5	Keeping in view the intermolecular forces of attraction, indicate which compound has the highest boiling point	A. H <sub>2</sub> S B. HF C. NH <sub>3</sub> <b>D. H<sub>2</sub>O</b>
6	Identify the covalent compound	<b>A. NaCl</b> B. H <sub>2</sub> O C. KF D. MgO
7	Which molecule contains a single covalent bond.	<b>A. CH<sub>4</sub></b> B. C <sub>2</sub> H <sub>4</sub> C. C <sub>2</sub> H <sub>2</sub> D. O <sub>2</sub>
8	Which types of attractive forces are present in ionic compounds.	A. Covalent bonds <b>B. Electrostatic forces of attraction</b> C. Metallic bonds D. Coordinate covalent bonds
9	Which element is capable of forming all the three types of bonds, covalent, coordinate covalent or ionic.	<b>A. Carbon</b> B. Silicon C. Magnesium D. Oxygen
10	The electropositive elements have the tendency to	<b>A. Lose electrons</b> B. Gain electrons C. Share electrons D. All of these
11	How many lone pairs are present on nitrogen in ammonia molecule.	<b>A. One</b> B. Two C. Three D. Four
12	Silicon belongs to Group IVA. It has ....electrons in the valence shell	<b>A. 2</b> B. 6 C. 3 D. 4
13	Why is H <sub>2</sub> O a liquid while H <sub>2</sub> S is a gas?	A. Because in water, the atomic size of oxygen is smaller than that of Sulphur B. Because water can easily freeze into ice <b>C. Because water is a polar compound</b> and there exists strong forces of attraction between its molecules D. Because H <sub>2</sub> O molecule is lighter than H <sub>2</sub> S

14	Ionic compound have	<b>B. Low melting and boiling point</b> <b>C. High melting and boiling points</b> <b>D. High melting and low boiling points</b>
15	Which of the following atoms obey duplet rule.	A. O <sub>2</sub> B. Cl <sub>2</sub> <b>C. H<sub>2</sub></b> D. Li <sub>2</sub>
16	Which of the following atoms will not form cation or anion.	A. Atomic no. 16 <b>B. Atomic no. 18</b> C. Atomic no. 17 D. Atomic No. 19
17	Which types of bond is present between NH <sub>3</sub> and BF <sub>3</sub>	A. Covalent Bond B. Ionic Bond <b>C. Co-ordinate covalent bond</b> D. Metallic Bond
18	A bond formed by the mutual sharing an electron pair is called.	A. Ionic bond B. Metallic bond <b>C. Covalent bond</b> D. Coordinate covalent bond
19	The compound formed by opposite charges are known as.	A. Metallic solids <b>B. Ionic compounds</b> C. Non-polar Covalent compound D. None of the above
20	Noble gases are non-reactive, because they do not.	A. Gain electrons B. Lose electrons C. Share electrons <b>D. All of these</b>