

## Chemical Bonding

Sr	Questions	Answers Choice
1	The fore of attraction betwene water molecule is.	A. Ionic bonding B. Covalent bonding C. Hydrigen Bonding D. Co ordinate Covalent bonding
2	Triple covalent bond involves how many electrons.	A. Six B. Four C. Eight D. Three
3	Covalent Bond involves the	A. Sharing of electrons B. Repulsion of electrons C. Acceptance of electrons D. Donation of electrons
4	Which ionic compound has the highest melting point.	A. RbCl B. KCl C. LiCl D. NaCl
5	Keeping in view the intermolecular forces of attraction, indicate which compound has the highest boiling point	A. $H_2S$ B. HF C. $NH_3$ D. $H_2O$
6	Identify the covalent compound	A. NaCl B. $H_2O$ C. KF D. MgO
7	Which molecule contains a single vovalent bond.	A. $CH_4$ B. $C_2H_4$ C. $C_2H_2$ D. $O_2$
8	Which types of attractive forces are present in ionic compounds.	A. Covalent bonds B. Electrostatic forces of attraction C. Mtallic bonds D. Coordinate covalent bonds
9	Which element is capable of forming all the three types of bonds, covalent coordinate covalent or ionic.	A. Carbon B. Silicon C. Magnesium D. Oxygen
10	The electropositive elements have the tendency to	A. Lose electrons B. Gain electrons C. Share electrons D. All of these
11	How many lone pairs are present on nitrogen in ammonia molecule.	A. One B. Two C. Three D. Four
12	Silicon belongs to Group IVA . It has ....electrons in the valence shell	A. 2 B. 6 C. 3 D. 4
13	Why is $H_2O$ liquid while $H_2S$ is a gas?	A. Because in water, the atomic size of oxygen is smaller than that of Sulphur B. Because water can easily freeze into ice C. Because water is a polar compound and there exists strong forces of attraction between its molecules D. Because $H_2O$ molecule is lighter than $H_2S$
		A. Low melting and high boiling points

14	Ionic compound have	B. Low melting and boiling point C. High melting and boiling points D. High melting and low boiling points
15	Which of the following atoms obey duplet rule.	A. $O_{2/2}$ B. $Cl_{2/2}$ C. $H_{2/2}$ D. $Li_{2/2}$
16	Which of the following atoms will not form cation or anion.	A. Atomic no. 16 B. Atomic no. 18 C. Atomic no. 17 D. Atomic No. 19
17	Which types of bond is present between $NH_3$ and $BF_3$	A. Covalent Bond B. Ionic Bond C. Coordinate covalent bond D. Metallic Bond
18	A bond formed by the mutual sharing an electron pair is called.	A. Ionic bond B. Metallic bond C. Covalent bond D. Coordinate covalent bond
19	The compound formed by opposite charges are known as.	A. Metallic solids B. Ionic compounds C. Non-polar Covalent compound D. None of the above
20	Noble gases are non-reactive, because they do not.	A. Gain electrons B. Lose electrons C. Share electrons D. All of these