

Chemical Bonding

Sr	Questions	Answers Choice
1	How many electrons are involved in the formation of single covalent bond	A. One B. Two C. Three D. Four
2	Keeping in view the intermolecular forces of attraction, indicate which compound has the highest boiling point	A. H_2S B. HF C. NH_3 D. H_2O
3	Triple covalent bond involves how many electrons.	A. Six B. Four C. Eight D. Three
4	Octet rule is	A. Attainment of eight electrons in its valence shell B. Description of eight electrons C. Picture of electronic configuration D. Pattern of electronic configuration
5	Transfer of electron between elements result in.	A. Coordinate covalent bonding B. Ionic bonding C. Metallic bonding D. Covalent bonding
6	Metal have the tendency to lose electrons due to.	A. High ionization energies B. Low ionization energies C. Low electron affinity D. None of the above
7	A bond formed by the mutual sharing an electron pair is called.	A. Ionic bond B. Metallic bond C. Covalent bond D. Coordinate covalent bond
8	How many valence shell electrons are there in Na^+ ion.	A. 8 B. 9 C. 1 D. 10
9	Silicon belongs to Group IVA. It haselectrons in the valence shell	A. 2 B. 6 C. 3 D. 4
10	Which one of the following is the weakest force among the atoms.	A. Intermolecular force B. Ionic force C. Metallic force D. Covalent forces
11	An atom having six electrons in its valence shell will achieve noble gas electronic configuration by	A. Gaining one electron B. Gaining two electrons C. Losing all electrons D. Losing two electrons
12	During the formation of ionic bond heat is.	A. Remains same B. Absorbed C. Released D. Both a and b
13	Which metal has the lowest melting point?	A. Li B. Na C. Rb D. K
14	Noble gases are non-reactive, because they do not.	A. Gain electrons B. Lose electrons C. Share electrons D. All of these
15	Atoms react with each other because.	A. they are attracted towards each other B. They are short of electrons

C. They want to disperse
D. They want attain stability

16 A bond pair is covalent molecules usually has.

A. One electron
B. Two electron
C. Three electron
D. Four electron

17 When molten copper and often zinc are mixed together , they give rise to a new substnce called brass. Predict what type of bond is formed between copper and zinc.

A. Ionic bond
B. Coordinate Covalent bond
C. Metallic bond
D. Covalent Bond

18 The compund formed by oppoite charges are known as.

A. Metallic solids
B. Ionic compounds
C. Non-polar Covalent compound
D. None of the above

19 Covalent bond is most commonly found between the elements of group

A. 1 to 13
B. 16 to 18
C. 13 to 17
D. 15 to 18

20 A bond formed between two non metals is expected to be

A. Ionic
B. Coordinate covalent
C. Metallic
D. Covalent
