

Chemical Bonding

Sr	Questions	Answers Choice
1	How many covalent bonds does C_2H_2 molecule have.	A. Two B. Three C. Four D. Five
2	Which of the following atoms obey duplet rule.	A. $O_{2/2}$ B. $Cl_{2/2}$ C. $H_{2/2}$ D. $Li_{2/2}$
3	The force of attraction between water molecules is.	A. Ionic bonding B. Covalent bonding C. Hydrogen Bonding D. Coordinate Covalent bonding
4	Keeping in view the intermolecular forces of attraction, indicate which compound has the highest boiling point	A. H_2S B. HF C. NH_3 D. H_2O
5	How many valence shell electrons are there in Na^+ ion.	A. 8 B. 9 C. 1 D. 10
6	Which of the following atoms will not form cation or anion.	A. Atomic no. 16 B. Atomic no. 18 C. Atomic no. 17 D. Atomic No. 19
7	Which among the following has a double covalent bond.	A. Ethane B. Methane C. Acetylene D. Ethylene
8	Metals have the tendency to lose electrons due to.	A. High ionization energies B. Low ionization energies C. Low electron affinity D. None of the above
9	Silicon belongs to Group IVA. It haselectrons in the valence shell	A. 2 B. 6 C. 3 D. 4
10	A covalent bond formed by the mutual sharing of two pairs of electrons between bonded atoms is called.	A. Single covalent bond B. Double covalent bond C. Triple covalent bond D. Polar covalent bond
11	Which group of the periodic table has the tendency to gain electrons.	A. Group -1 B. Group -17 C. Group-2 D. Group -18
12	Ionic compounds are good conductors of electricity in	A. Solution B. Molten state C. Solid state D. both a and b
13	How many electrons are involved in the formation of a single covalent bond	A. One B. Two C. Three D. Four
14	Which element is capable of forming all the three types of bonds, covalent coordinate covalent or ionic.	A. Carbon B. Silicon C. Magnesium D. Oxygen
15	Covalent Bond involves the	A. Sharing of electrons B. Repulsion of electrons C. Acceptance of electrons D. Donation of electrons

16	Hydrogen and Helium follow.	A. Octet rule B. Triple rule C. Duplet rule D. None of these
17	Which metal has the lowest melting point?	A. Li B. Na C. Rb D. K
18	Why is H ₂ O a liquid while H ₂ S is a gas?	A. Because in waer, the atomic size of oxygen is smaller than that of Sulphur B. Because water can easily freeze into ice C. Because water is a polar compound ad tere exists strong forces of attrction betwene its molecules D. Because H ₂ O molecule is lighter than H ₂ S
19	Water has high boiling point as compared a alcohol due to	A. Low density B. High surface tension C. Hydrogen bonding D. High vapour pressure
20	Octet rule is	A. Attaingn of eight electrons in its valence shell B. Discription of eight electrons C. Pictur eof electronic configuration D. Pattern of electronic configuration