

Chemical Bonding

Sr	Questions	Answers Choice
1	During the formation of ionic bond heat is.	A. Remains same B. Absorbed C. Released D. Both a and b
2	Every atom has a natural tendency to accommodate electrons in its valence shell	A. 2 or 6 B. 2 or 4 C. 2 or 8 D. 2 or 10
3	In the formation of AlF_3 , aluminum atom loses.....electrons.	A. 1 B. 4 C. 3 D. 2
4	When molten copper and olden zinc are mixed together , they give rise to a new substance called brass. Predict what type of bond is formed between copper and zinc.	A. Ionic bond B. Coordinate Covalent bond C. Metallic bond D. Covalent Bond
5	Number of electrons in nitrogen molecule is.	A. 2 B. 4 C. 6 D. 8
6	How many valence shell electrons are there in Na^+ ion.	A. 8 B. 9 C. 1 D. 10
7	Transfer of electron between elements result in.	A. Coordinate covalent bonding B. Ionic bonding C. Metallic bonding D. Covalent bonding
8	Identify the compound which is not soluble in water	A. KBr B. $MgCl_2$ C. C_6H_6 D. NaCl
9	A bond formed between two non metals is expected to be	A. Ionic B. Coordinate covalent C. Metallic D. Covalent
10	The electropositive elements have the tendency to	A. Lose electrons B. Gain electrons C. Share electrons D. All of these
11	Which of the following bonds is expected to be the weakest.	A. $Cl-Cl$ B. $C-C$ C. $F-F$ D. $O-O$
12	Which molecule contains a single covalent bond.	A. CH_4 B. C_2H_4 C. C_2H_2 D. O_2
13	When an electronegative element combines with electropositive element, the type of bonding is.	A. Covalent B. Polar Covalent C. Ionic D. Coordinate Covalent
14	Metals have the tendency to lose electrons due to.	A. High ionization energies B. Low ionization energies C. Low electron affinity D. None of the above
15	Hydrogen bonding is always found in	A. Non-polar molecules B. Homonuclear molecules C. Polar Molecules D. All of the above

16	Nitrogen molecule contains.	A. Polar covalent bond B. Triple Covalent bond C. Double covalent bond D. Single covalent bond
17	Which of the following atoms will not form cation or anion.	A. Atomic no. 16 B. Atomic no. 18 C. Atomic no. 17 D. Atomic No. 19
18	Which of the following atoms obey duplet rule.	A. $O_{2^{2-}}$ B. $Cl_{2^{2-}}$ C. $H_{2^{2-}}$ D. $Li_{2^{2-}}$
19	In metals, the hold of nucleus over the valence shell electrons is weak due to.	A. High electron affinity B. Large sized atoms C. High ionization energies D. All of the above
20	Which types of attractive forces are present in ionic compounds.	A. Covalent bonds B. Electrostatic forces of attraction C. Metallic bonds D. Coordinate covalent bonds