

## Chemical Bonding

Sr	Questions	Answers Choice
1	During the formation of ionic bond heat is.	A. Remains same B. Absorbed <b>C. Released</b> D. Both a and b
2	Every atom has a natural tendency to accommodate electrons in its valence shell	A. 2 or 6 B. 2 or 4 <b>C. 2 or 8</b> D. 2 or 10
3	In the formation of $\text{AlF}_3$ , aluminum atom loses.....electrons.	A. 1 B. 4 <b>C. 3</b> D. 2
4	When molten copper and often zinc are mixed together , they give rise to a new substnce called brass. Predict what type of bond is formed between copper and zinc.	A. Ionic bond B. Coordinate Covalent bond <b>C. Metallic bond</b> D. Covalent Bond
5	Number of electronsin nitrogen molecule is.	A. 2 B. 4 <b>C. 6</b> D. 8
6	How many valance shell electrons are there in $\text{Na}^+$ ion.	<b>A. 8</b> B. 9 C. 1 D. 10
7	Transfer of electron between elements result in.	A. Coordinate covalent bonding <b>B. Ionic bonding</b> C. Metallic bonding D. Covalent bonding
8	Identify the compound whcihis not soluble in water	A. $\text{KBr}$ B. $\text{MgCl}_2$ <b>C. <math>\text{C}_6\text{H}_6</math></b> D. $\text{NaCl}$
9	A bond formed between two non metals is expected to be	A. Ionic B. Coordinate covalent C. Metallic <b>D. Covalent</b>
10	The electropositive elements have the tendency to	<b>A. Lose electrons</b> B. Gain electrons C. Share electrons D. All of these
11	Which of the following bond s is expected to b the weakest.	A. $\text{Cl}-\text{Cl}^{-2}$ B. C-C <b>C. F-F</b> D. O -O
12	Which molecule contains a single voalent bond.	<b>A. <math>\text{CH}_4</math></b> B. $\text{C}_2\text{H}_4$ C. $\text{C}_2\text{H}_2$ D. $\text{O}_2$
13	When an electronegative element combines with electrpositive element, the type of bonding. is.	A. Covalent B. Polar Covalent <b>C. Ionic</b> D. Coorinae Covalent
14	Metal have the tendency to lose elecrons due to.	A. High ionization energies <b>B. Low ionization energies</b> C. Low electron affinity D. None of the above
15	Hydrogen bound ing is always found in	A. Non-polar molecules B. Homoatomic molecules <b>C. Polar Molecules</b> D. All of the above

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16	Nitrogen molecule contains.	A. Polar covalent bond B. <b>Triple Covalent bond</b> C. Double covalent bond D. Single covalent bond
17	Whcih of the following atoms willnot form cation or anion.	A. Atomic no. 16 B. <b>Atomic no. 18</b> C. Atomic no. 17 D. Atomic No. 19
18	Whcih of the following atoms obey duplet rule.	A. O <sub>2</sub> B. Cl <sub>2</sub> <b>C. H<sub>2</sub></b> D. Li <sub>2</sub>
19	In metals, the hold of nucieus over the valence shell elecrons is veak due to.	A. Highelectron affinity <b>B. Large sized atoms</b> C. High ionization energies D. All of the above
20	Which types of attractive forces are presentin ionic compounds.	A. Covalent bonds <b>B. Electrostatic forces of attraction</b> C. Mtallic bonds D. Coordinate covalent bonds

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