

ECAT Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	Volcanoes produce SO ₂ :	A. 47% B. 57% C. 67% D. 77%
2	With the increase in carbon no. the solubility of carboxylic acids	A. Increases B. Decreases C. Remains same D. None of these
3	Hydration energy depends upon:	A. Change of ion B. Size of ion C. Charge to size ratio D. Heat changes
4	In the reaction between chlorine and U.V. the propagation step is.	A. CH ₃ +CL -----CH ₃ +HCl B. CH ₄ +Cl-----CH ₃ Cl C. CH ₄ +CL -----CH ₃ CL +H D. CH+CL ₂ -----CL +HCl
5	The value of K _p is greater than K _c for a gaseous reaction when	A. Number of molecules of products is greater than the reactants B. Number of molecules of reactants is greater than those of products C. Number of molecules of reactants and products equal D. Catalyst is added
6	When a bond breaks	A. Heat is evolved B. Heat is absorbs C. No change in heat contents takes place D. Temperature increases
7	S _N 1 reaction of alkylhalides leads to	A. Retention of configuration B. Racemisation C. Inversion of configuration D. None of these
8	A suitable solvent should dissolve maximum amount of solute at its boiling point and minimum amount at :	A. Freezing point. B. Room temperature. C. Boiling point. D. Sea level temperature.
9	Carbon monoxide is pollutant as it	A. Inactivates nerves B. Inhibits glycolysis C. Combines with oxygen D. Combines with hemoglobin
10	Question Image	A. Rate is independent of concentration of water since it is in excess B. Rate is independent of concentration of ester since it is in excess C. Rate depends upon the concentration of acid catalyst added D. Rate = $k[\text{CH}_3\text{COOC}_2\text{H}_5]^3[\text{H}_2\text{O}]^{1/2}$
11	Question Image	A. $\text{RCH}(\text{CH}_3)_3\text{CO}_2\text{H} + \text{CH}_3\text{OH}$ B. $\text{RCH}(\text{CH}_3)_3\text{CO}_2\text{H} + \text{HCO}_2\text{H}$ C. $\text{RCH}(\text{CH}_3)_3\text{OH} + \text{CO}_2$ D. $\text{RCH}(\text{CH}_3)_3\text{OH} + \text{HCO}_2\text{H}$
12	Which of the following has highest oxidation potential	A. Be B. Li C. Na D. Ca
13	Which electronic configuration corresponds to an element of Group 12-A of the periodic table?	A. 1s, 2s ² , 2p ⁶ , 3s ² , 3p ⁶ , 4s ² B. 1s ² , 2s ² , 2p ⁶ , 3s ² , 3p ¹ C. 1s ² , 2s ² , 2p ⁶ D. 1s ² , 2s ² , 2p ⁶

14	Which among the following species has the highest ionization energy?	A. Ne B. F C. Li D. B
15	Periodic law was given by:	A. Al-Razi B. Dobriener C. Newland D. Mendeleev
16	The number of atoms present in molecule determines its:	A. Molecularity B. Atomicity C. Basicity D. Acidity
17	When 0.1 g of magnesium is treated with an excess of hydrochloric acid, what volume of gas at room temperature and pressure will be produced	A. 10 cm ³ B. 25 cm ³ C. 48 cm ³ D. 100 cm ³
18	Which is the strongest acid?	A. HClO B. HClO ₂ C. HClO ₃ D. HClO ₄
19	Butyric acid was named from butyrum means:	A. Red out B. Vinegar C. Butter D. Milk
20	1 mole of N ₂ and 2 moles of H ₂ are allowed to react in a 1 dm ³ vessel. At equilibrium 0.8 mole of NH ₃ is formed. The concentration of H ₂ in the vessel is	A. 0.6 mole B. 0.8 mole C. 0.2 mole D. 0.4 mole