

## ECAT Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	In an exothermic reaction, a 10° rise in temperature will	<p>A. Decrease the value of equilibrium constant</p> <p>B. Double the value of <math>K_c</math></p> <p>C. Not produce any change in <math>K_c</math></p> <p>D. Produce some increase in <math>K_c</math></p>
2	With the increase in carbon no. the solubility of carboxylic acids	<p>A. Increases</p> <p>B. Decreases</p> <p>C. Remains same</p> <p>D. None of these</p>
3	Alkenes normally have _____ geometry	<p>A. Tetrachedral</p> <p>B. Linear</p> <p>C. Planer</p> <p>D. None</p>
4	What is going to replace the petroleum?	<p>A. Silica</p> <p>B. Silicates</p> <p>C. Silicones</p> <p>D. Silicon</p>
5	The actual number of atoms or molecules taking part in rate determining step is	<p>A. Rate of reaction</p> <p>B. Velocity of reaction</p> <p>C. Order of reaction</p> <p>D. Molecularly</p>
6	When copper is allowed to react with $\text{HNO}_3$ , the reaction is slow in the beginning, finally becomes very fast. It is due to the formation of an auto catalyst which is	<p>A. <math>\text{Cu}(\text{NO}_3)_2</math></p> <p>B. <math>\text{CuO}</math></p> <p>C. <math>\text{O}_2</math></p> <p>D. <math>\text{HNO}_2</math></p>
7	A solution of sodium sulphate was electrolysed using some inert electrodes. The products at the electrodes are	<p>A. <math>\text{O}_2</math>, <math>\text{H}_2</math></p> <p>B. <math>\text{O}_2</math>, Na</p> <p>C. <math>\text{O}_2</math>, <math>\text{SO}_2</math></p> <p>D. <math>\text{O}_2</math>, <math>\text{S}_2\text{O}_8^{2-}</math></p>
8	Atmospheric pollutant is	<p>A. <math>\text{CO}_2</math></p> <p>B. CO</p> <p>C. <math>\text{O}_2</math></p> <p>D. <math>\text{N}_2</math></p>
9	In Friedal-Craft's alkylation besides $\text{AlCl}_3$ the other reactants are	<p>A. <math>\text{C}_6\text{H}_6 + \text{NH}_3</math></p> <p>B. <math>\text{C}_6\text{H}_6 + \text{NH}_4</math></p> <p>C. <math>\text{C}_6\text{H}_6 + \text{CH}_3\text{Cl}</math></p> <p>D. <math>\text{C}_6\text{H}_6 + \text{CH}_3\text{COCl}</math></p>
10	The highest temperature at which a substance can exist as a liquid is called its	<p>A. Critical temperature</p> <p>B. Zero temperature</p> <p>C. Absolute temperature</p> <p>D. None of above</p>
11	The function of salt bridge in galvanic cell is	<p>A. To prevent accumulation of ions in two half</p> <p>B. To add salt ions in two half</p> <p>C. To block flow of ions between two half</p> <p>D. None of these</p>
12	Which of the following species is paramagnetic?	<p>A. <math>\text{CO}_2</math></p> <p>B. NO</p> <p>C. <math>\text{O}^{2-}</math></p> <p>D. CN</p>
13	All covalent bonds formed between the two atoms are non-polar when	<p>A. Covalent bond between two non-metal atoms</p> <p>B. Covalent bond between metal and non-metal</p> <p>C. Covalent bond between two atoms of same element</p> <p>D. Covalent bond between metal</p>

atoms

14	Saponification of ethyl benzoate with caustic soda	A. Benzyl alcohol, ethanoic acid B. Sodium benzoate, ethanol C. Benzoic acid, sodium ethoxide D. Phenol, ethanoic acid
15	The force which holds together two or more atoms or ions to form a large variety of compounds is called:	A. A chemical bond. B. An ionic bond. C. A covalent bond. D. A coordinate bond.
16	Which is not a property of ether:	A. Very weak hydrogen bonding B. High b.p C. Slightly soluble D. Inflammable
17	The reaction rate is expressed in the units of	A. $\text{mol dm}^{-3}\text{s}^{-1}$ B. $\text{mol dm}^{-3}$ C. $\text{mol dm}^{-3}\text{N}^{-1}\text{s}^{-1}$ D. $\text{dm}^{-3}\text{s}^{-1}$
18	The boiling point of $\text{NH}_3$ is maximum among the hydrides of group V elements due to:	A. Enhanced electronegative character of Nitrogen B. Pyramidal structure of $\text{NH}_3$ C. Very small size of Nitrogen. D. Enhanced electropositive character of Nitrogen
19	During the process of crystallization, the hot saturated solution:	A. is cooled very slowly to get large size crystals B. is cooled at a moderate rate to get medium sized crystals of the product C. is evaporated to get the crystals of the products D. is mixed with an immiscible liquid to get the pure crystals of the product.
20	Aldol condensation is actually	A. Electrophilic addition of carbonation B. Electrophilic addition of carbonium ion C. Nucleophilic addition of carbonation D. Nucleophilic addition of carbonium ion