

## ECAT Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	Baukelite is a polymer obtained from two monomers	A. Phenol and ethanol B. Phenol and methanol <b>C. Phenol and methanal</b> D. Phenol and acetone
2	A solution is a homogeneous mixture of two or more kinds different :	A. Molecular. B. Covalent substance <b>C. Ionic Substances</b> D. Both (a) and (c)
3	A beaker contains 9 grams of water. The number of H-atoms is	<b>A. <math>6.02 \times 10^{23}</math></b> B. $3.01 \times 10^{23}$ C. $6.02 \times 10^{24}$ D. $3.01 \times 10^{24}$
4	The equilibrium constant in a reversible chemical reaction at a given temperature	A. Depends on the initial concentration of the reactants B. Depends on the concentration of one of the products at equilibrium <b>C. Does not depend on the initial concentration of reactants</b> D. It is characteristic of the reaction
5	For a given process, the heat changes at constant volume ( $q_v$ ) are related to other as:	A. $q_v = q_p - \Delta n R T$ B. $q_v = q_p + \Delta n R T$ C. $q_v = q_p$ <b>D. <math>q_v = q_p - \Delta n R T</math></b>
6	Factor which slows down the rate of reaction is	A. Small size of the particles of the reactant B. High temperature of reaction C. More concentration of reactant <b>D. Lowering the temperature</b>
7	Toluene is also called	A. Hydroxyl benzene <b>B. Methyl benzene</b> C. ethyl benzene D. None
8	The reaction between fat and NaOH is called	A. Esterification B. Hydrogenolysis C. Fermentation <b>D. Saponification</b>
9	Ground and surface waters are contaminated and become polluted due to the human activity. Which human activity will not cause water pollution	A. Live stock waste B. Agricultural pesticides C. Oil breaks and spills <b>D. All of the above</b>
10	The amount of electricity that can deposit 108 g of silver from silver nitrate solution is	A. 1 ampere B. 1 coulomb <b>C. 1 faraday</b> D. 2 ampere
11	In the reaction between chlorine and U.V. the propagation step is.	<b>A. <math>CH_3 + Cl \rightarrow CH_3 + HCl</math></b> B. $CH_4 + Cl \rightarrow CH_3Cl + H$ C. $CH_4 + Cl \rightarrow CH_3Cl + H$ D. $CH + Cl_2 \rightarrow Cl + HCl$
12	Question Image	A. Heat of reaction B. Heat of formation <b>C. Heat of neutralization</b> D. Heat of combustion
13	The reaction of sodium with ethanol gives	A. $CH_3CH_2O^- Na^+$ B. NaH C. $H_2$ <b>D. <math>C_2H_5O^- Na^+</math></b>
		<b>A. Internal energy of system</b>

14	Heat absorbed by a system when its volume does not change is equal to	B. Enthalpy of system C. Increase in internal energy of system D. Increase in enthalpy of system
15	Aromatic hydrocarbons are the derivatives of:	A. Normal series of paraffins B. Alkene C. Benzene D. Cyclohexane
16	What will be the molarity of solution if 103 g $(\text{NH}_4)_2\text{SO}_4$ is dissolved per 600 $\text{cm}^3$ of water	A. 2.32 M B. 3.32 M C. 4.32 M D. 1.30 M
17	In endothermic reactions, the heat content of the:	A. Products is more than that of reactants. B. Reactants is more than than to products. C. Both (a) and (b). D. Reactants and products are equal.
18	The noble gas have	A. Very low ionization energies B. High boiling points C. No electron pair interaction D. No van Der Waal's forces
19	A cell constant is generally found by measuring the conductivity of aqueous solution of	A. $\text{BaCl}_2$ B. KCl C. NaCl D. $\text{MgCl}_2$
20	Which of the following discoveries resulted in a version of the Mendeleef's periodic law	A. The nucleus of atom by Rutherford B. The elements polonium and radium by the Curies C. Atomic numbers by Moseley D. x-rays by Roentgen