ECAT Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	Platinum is poisoned by	A. Arsenic B. Silver C. Argon D. Zinc
2	Atoms and molecules can either gain or lose electrons, forming charge particles called:	A. Positrons<o:p></o:p> B. Photons<o:p></o:p> C. Ions<o:p></o:p> D. Electrons<o:p></o:p>
3	If an endothermic reaction is allowed to take place very rapidly in air, the temperature of the surrounding air :	A. Remains constant. B. Decreases. C. Increases. D. Fluctuates rapidly.
4	If the ionic product of a solution is less than the solubility product, the solution is	A. Supersaturated B. Unsaturated C. Ideal D. Saturated
5	The strongest forces are	A. Debye forces B. London dispersion forces C. Dipole-dipole attraction D. Hydrogen
6	Choose the correct answer of transition elements?	A. Transition elements have low melting points B. Transition elements do not have catalytic activity C. Transition elements exhibit variable oxidation states D. Transition elements exhibit inert pair effect
7	Schrodinger wave equation describes electron completely because	A. It describes a set of four quantum number B. It describes the particle nature of electron C. It measures wavelength of electron D. It describes electron moving in specific orbit
8	Scientist(s) made contributions in the field designing a periodic table	A. Newland B. Dobriener C. Mendeleev & Mosley D. All of these
9	The number of atoms present in molecule determines its:	A. Molecularity B. Atomicity C. Basicity D. Acidity
10	The stability of acromatic compounds decreases with ------- in the no. of its resonance structure	A. Decrease B. Increase C. Remain constant D. Partially decreases
11	75% plant nutrients are present in the fertilizer:	A. Urea B. NH₄NO₃ C. NH₃ D. (NH₄)₂HPO₄
12	3.6% WN solution of HCl has the molairity	A. 1.0 B. 1.15 C. 0.98 D. 1.98

13	The concentration units independent of temperature would be	A. Normality B. Mass-volume precent C. Molality D. Molarity
14	The specific conductance of a 0.1 N KCl solution at 23°C is 0.012 $ohm^{-1}cm^{-1}$. The resistance of cell containing the solution at the same temperature was found to be 55 ohm. The cell constant will be	A. 0.142 cm⁻¹ B. 0.66 cm⁻¹ C. 0.916 cm⁻¹ D. 1.12 cm⁻¹
15	The ph of 10-3 mole dm-3 of an aqueous solution of H_2SO_4 is :	A. 3.0<o:p></o:p> B. 2.7<o:p></o:p> C. <span style="font-size:11.0pt;line-height:107%; font-family:"Calibri",sans-serif;mso-ascii-theme-font:minor-latin;mso-fareast-font-family: Calibri;mso-fareast-theme-font:minor-latin;mso-hansi-theme-font:minor-latin; mso-bidi-font-family:Arial;mso-bidi-theme-font:minor-bidi;mso-ansi-language: EN-US;mso-fareast-language:EN-US;mso-bidi-language:AR-SA">2.0 D. 1.5<o:p></o:p>
16	Amongst the following the compound that can be most readily sulphonated is	A. Toluene B. Benzene C. Nitrobenzene D. Chlorobenzene
17	Nylon 6, 6 is prepared by the reaction of nexamethylene with	A. Formic acid B. Acetic acid C. Adipic acid D. None of these
18	Cotton is _____% cellulose	A. 90 B. 100 C. 99 D. 30
19	Question Image	A. 32 B. 64 C. 16 D. 4
20	In the rate equation when the concentration of reactants are unity, then rate is equal to	A. Instantaneous rate B. Average rate C. Active mass of products D. Specific rate constant