

ECAT Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	A carbon atom carrying a positive charge and attached to three other atoms of groups is called	A. Caronium ion B. Carbanion C. Oconium ion D. Carba ion
2	Which of the following half molar solutions will have lowest freezing point	A. Solution of non-volatile, none electrolyte B. Solution of non volatile, weak electrolyte C. Solution of non volatile strong electrolyte D. Solution of volatile, non electrolyte
3	Each molecule of haemoglobin is 68000 times heavier than one atom of	A. C B. H C. N D. O
4	The solutions of NaCl and KCl are prepared separately by dissolving same amount of solute in water, which of the following statements is true fro these solutions ?	A. KCl solution will have higher boiling point than NaCl solution. B. Both the solutions have same boiling points. C. KCl and NaCl solutions possess same vapour pressure. D. KCl solution possesses lower freezing point than NaCl solution.
5	Preparation of vegetable ghee involves	A. Halogenations B. Hydrogenation C. Hydroxylation D. Dehydrogenation
6	Question Image	A. Neutrons are attracted by nucleus B. Neutrons carry out nuclear reactions C. Neutrons carry no charge D. Neutrons are electromagnetic radiations
7	Which of the following is not an element?	A. Graphite B. Diamond C. 22-Carat gold D. Rhombic sulphur
8	The four quantum numbers of the valency electron of potassium are	A. 4, 1, 1, 1/2 B. 4, 0, 0, 1/2 C. 4, 1, 0, 1/2 D. 4, 4, 0, 1/2
9	As compared to molar solution, in the molal solution the quantity of solvent is :	A. Comparatively lesser B. More or less equal C. Comparatively greater D. Very large
10	Stainless steel contains ion and carbon along with	A. Ni and Cr B. Cr and Co C. Co and Mn D. Mn and Ni
11	The reaction of Zinc with copper sulphate solution is an example of	A. Oxidation reduction reaction B. Spontaneous reaction C. Spontaneous redox reaction D. Non-spontaneous reaction
12	An organic compound, on treatment with Br ₂ in CCl ₄ gives bromoderivative of an alkene. The compound will be	A. CH ₃ - CH = CH ₂ B. CH ₃ CH = CHCH ₃ C. HC = CH D. H ₂ C = CH ₂
13	A molecule in which sp ² hybrid orbitals are used by the central atom in forming covalent bonds in	A. He ₂ B. SO ₂ C. PCl ₅ D. N ₂
		A. Heat is evolved

14	When a bond breaks	<p>B. Heat is absorbed</p> <p>C. No change in heat contents takes place</p> <p>D. Temperature increases</p>
15	Which of the following oxides is peroxide?	<p>A. Na_2O_2</p> <p>B. MnO_2</p> <p>C. BaO</p> <p>D. SO_2</p>
16	The correct order of reactivity of halogens with alkanes is	<p>A. I_2 > Br_2 > Cl_2 > F_2</p> <p>B. I_2 > Cl_2 > F_2 > Br_2</p> <p>C. F_2 > Cl_2 > I_2 > Br_2</p> <p>D. F_2 > Cl_2 > Br_2 > I_2</p>
17	Which of the following is a typical transition metal	<p>A. Sc</p> <p>B. Y</p> <p>C. Ra</p> <p>D. Co</p>
18	The order of frequency of the following radiations ultraviolet, visible, infrared and microwave is	<p>A. Microwave > infrared > visible > ultraviolet</p> <p>B. Visible > ultraviolet > microwave > infrared</p> <p>C. Ultraviolet > visible > infrared > microwave</p> <p>D. Infrared > microwave > ultraviolet > visible</p>
19	Out of 110 known elements, transition elements are	<p>A. 40</p> <p>B. 60</p> <p>C. 50</p> <p>D. 80</p>
20	Value of rate constant k is specific for a reaction, and varies from reaction to reaction. The value of k of a reaction changes with	<p>A. Time</p> <p>B. Temperature</p> <p>C. Concentration of reactants</p> <p>D. Order of reaction</p>