

## ECAT Chemistry Online Test

Sr	Questions	Answers Choice
1	Which is not the condition for the formation of smog	A. There must be sufficient NO gas B. There must be sunlight to help photo chemical reaction to take place C. Air must be blowing swiftly D. There must be SO <sub>2</sub> in the air
2	When acetylene is passed through a copper tube at 300°C, it polymerizes to	A. Polyacetylene B. polyethylene C. Benzene D. None of these
3	ph of the buffer CH <sub>3</sub> COOh + CH <sub>3</sub> COONa is 3.76. If the mixture contains 1 molar acetic acid and 0.1 molar sodium acetate, then pKa of this buffer is	A. 3.76 B. 4.76 C. 5.76 D. 6.76
4	Acetamides are formed by the reaction of carboxylic acids with	A. Acids B. Bases C. Salts D. NH <sub>3</sub>
5	Ozone in most of the tropical regions acts as a pollutant and causes	A. Damages to eyes B. Aggravates asthma C. Chest discomfort D. All of these
6	Question Image	A. Ethanol in the presence of concentrated sulphuric acid     B. Potassium hydroxide     C. Sodium     D. Sodium carbonate
7	Estimation of Na in sea water is an example of :	A. Numerical analysis.     B. Qualitative analysis.     C. Quantitative analysis.     D. None of above.
8	Two solutions of NaCl and KCl are prepared separately by dissolving 0.1 M of the solute in water. Which of the following statements is not true for these solution	A. KCl solution will have higher boiling point than NaCl solution B. Both the solutions have same boiling C. KCl and NaCl solution possess same vapour pressure D. KCl solution possess same freezing point at NaCl solution
9	The formation of compounds like PF <sub>5</sub> , BCl <sub>3</sub> , SF <sub>6</sub> indicates that	A. These halides are ionic     B. These halides are covalent     C. They are Lewis acids     D. Octet rule not obeyed so the rule is not universal
10	The detection of functional group is called:	A. Numerical analysis B. Qualitative analysis C. Quantitative analysis D. Combustion analysis
11	Question Image	A. Measuring pH B. Measuring density C. Titration against standard NaOH D. Titration against standard KMnO <sub>4</sub> solution
12	The covalent bonds are	A. Unidirectional B. Bi-directional C. Non-directional D. Multi-directional
13	Precipitation will occur until the ionic product becomes	A. Equal to K <sub>sp</sub> B. Lesser than K <sub>sp</sub> C. Greater than K <sub>sp</sub> D. None of these
		A. Na <sub>2</sub> CO <sub>3</sub>

High properties of the following metal exhibits more than one oxidation?  A Na B. Mg C. Fe D. Al  A 7 B. zero C. 14 D. 10  A Weak acid and its salt with weak base C. Weak base and its salt with weak base C. Weak base and its salt with strong acid D. Strong base and its salt with strong acid D. Strong base and its salt with weak acid  A Order of reaction can determined by an experiment C. Order of reaction can determined from a balance equation only. C. Order of reaction can determined increases by increasing temperate. D. Order of reaction can determined increases by increasing temperate. D. Order of reaction can determined increases by increasing temperate. D. Order of reaction can determined increases by increasing temperate. D. Order of reaction can determined increases by increasing temperate. D. Order of reaction can determined increases by increasing temperate. D. Order of reaction can determined increases by increasing temperate. D. Order of reaction can determined increases by increasing temperate. D. Order of reaction can determined increases by increasing temperate. D. Order of reaction can determined increases by increasing temperate. D. Order of reaction can determined increases by increasing temperate. D. Order of reaction can determined from a balance equation only. C. Order of reaction can determined increases by increasing temperate. D. Order of reaction can determined from a balance equation only. C. Order of reaction can determined from a balance equation only. C. Order of reaction can determined from a balance equation only. C. Order of reaction can determined from a balance equation only. C. Order of reaction can determined from a balance equation only. C. Order of reaction can determined from a balance equation only. C. Order of reaction can determined from a balance equation only. C. Order of reaction can determined from a balance equation only. C. Order of reaction can determined from a balance equation only. C. Order of reaction can determined from a balance equation only. C. Order of re	14	Solvay process is used in the manufacture of	B. NaHCO <sub>3</sub> C. CaCl <sub>2</sub> D. All
pH of 1 molar NaOH is    PH of 1 molar NaOH is   PH of 1 molar NaOH is	15	Which of the following metal exhibits more than one oxidation?	B. Mg C. Fe
B. Strong acid and its salt with weak base C. Weak base and its salt with strong acid D. Strong base and its salt with weak acid  A. Order of reaction can determined by an experiment B. Order of reaction can determined from a balance equation only. C. Order of reaction can determined increases by increasing temperate. D. Order of reaction must be in whole number and not in fraction.  A. Liquid at room temperature B. Good electrical conductors C. Good reducing agents D. Easily reduced  A. Water B. Na <sub>2</sub> CO <sub>3</sub> CO <sub>CO<sub>3</sub>CO<sub>CO<sub>3</sub>CO<sub>CO<sub>3</sub>CO<sub>CO<sub>3</sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<sub>CO<s< td=""><td>16</td><td>pH of 1 molar NaOH is</td><td>B. zero C. 14</td></s<></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub></sub>	16	pH of 1 molar NaOH is	B. zero C. 14
The unit of rate constant K ismole <sup>-1</sup> dm <sup>3</sup> for a chemical reaction, the order of reaction is:  The unit of rate constant K ismole <sup>-1</sup> dm <sup>3</sup> for a chemical reaction, the order of reaction is:  The unit of rate constant K ismole <sup>-1</sup> dm <sup>3</sup> for a chemical reaction, the order of reaction is:  C. Order of reaction can determined increases by increasing temperate. D. Order of reaction must be in whole number and not in fraction.  A. Liquid at room temperature B. Good electrical conductors C. Good reducing agents D. Easily reduced  A. Water B. Na <sub>2</sub> CO <sub>3</sub> CO <sub>3</sub> CO<800H	17	Base buffer solution can be prepared by mixing	B. Strong acid and its salt with weak base C. Weak base and its salt with strong acid D. Strong base and its salt with weak
19 Ionic hydrides are generally  B. Good electrical conductors C. Good reducing agents D. Easily reduced  A. Water B. Na <sub>2</sub> CO <sub>3</sub> C. NaOH	18	The unit of rate constant K ismole $^{-1}$ dm $^3$ for a chemical reaction, the order of reaction is :	by an experiment B. Order of reaction can determined from a balance equation only. C. Order of reaction can determined increases by increasing temperate. D. Order of reaction must be in whole
20 Absolute alcohol is obtained by adding rectified apirit in alcohol:  B. Na <sub>2</sub> CO <sub>3</sub> C. NaOH	19	lonic hydrides are generally	B. Good electrical conductors     C. Good reducing agents
	20	Absolute alcohol is obtained by adding rectified apirit in alcohol:	B. Na <sub>2</sub> CO <sub>3</sub> C. NaOH