

## ECAT Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	Question Image	<p>A. A strong reducing agent</p> <p>B. A strong oxidising agent</p> <p>C. Better oxidising agent than hydrogen</p> <p>D. Less reducing agent than hydrogen</p>
2	$H_2 + I_2 \rightleftharpoons 2HI$ In the above equilibrium system, if the concentration of reactants at 25°C is increased, the value $K_C$ will :	<p>A. Remains Constant</p> <p>B. Increases</p> <p>C. Decreases</p> <p>D. Depends upon nature of reactants</p>
3	Quantum number values for 2p orbitals are	<p>A. <math>n = 2, l = 1</math></p> <p>B. <math>n = 1, l = 2</math></p> <p>C. <math>n = 1, l = 0</math></p> <p>D. <math>n = 2, l = 0</math></p>
4	Best way to prevent rusting of iron is by	<p>A. Making iron cathode</p> <p>B. Putting it in saline water</p> <p>C. Both of these</p> <p>D. None of these</p>
5	When we perform the same reaction by taking two different initial concentrations of a reactant for a second order reaction then	<p>A. Reaction becomes exothermic</p> <p>B. Energy of activation is different</p> <p>C. Mechanism of reaction is changed</p> <p>D. Half life period is changed</p>
6	Hybridization in alkanes is:	<p>A. sp</p> <p>B. <math>sp^2</math></p> <p>C. <math>sp^3</math></p> <p>D. <math>dsp^2</math></p>
7	All are ortho & Para directing except	<p>A. X</p> <p>B. OH</p> <p>C. <math>NR_2</math></p> <p>D. <math>NH_2</math></p>
8	Which of the following enzymes brings about the hydrolysis of fats?	<p>A. Urease</p> <p>B. Maltase</p> <p>C. Zymase</p> <p>D. Lipase</p>
9	The reaction of alcohol with $SOCl_2$ in the presence of pyridine as catalyst gives	<p>A. Acids</p> <p>B. Acid chloride</p> <p>C. Alkyl halide</p> <p>D. Benzene</p>
10	The hydroxyl derivatives of aromatic hydrocarbons which have the -OH group directly bonded to the ring C-atom are called	<p>A. Alcohols</p> <p>B. Ketones</p> <p>C. Esters</p> <p>D. Phenols</p>
11	Ionization potential increases in moving from left to right in a period	<p>A. Because nuclear charge increase</p> <p>B. Because atomic size decrease</p> <p>C. Both (a) and (b)</p> <p>D. Because atomic size increases</p>
12	Which one is used as a fuel for internal combustion engines in many European countries and Brazil	<p>A. <math>C_2H_5OH</math></p> <p>B. <math>CH_3OH</math></p> <p>C. <math>CH_3COOH</math></p> <p>D. <math>C_2H_5OH</math></p>
13	Zeotropic mixture	<p>A. Obey Henry's law</p> <p>B. Obey Raoult's law</p> <p>C. Does not obey Raoult's law</p> <p>D. Obey Dalton's law</p>
14	A Solution containing 6.8 g of non-ionic solute in 100g of water was found to freeze at -0.93°C. If $K_f$ for water is 1.86 and molecular mass of solute is	<p>A. 13.6</p> <p>B. 34</p> <p>C. 68</p> <p>D. 136</p>

A.  $Na^+ F^-$   
 B.  $Na^+ F^-$

15	The positive ion is always smaller than the neutral atom while the negative ion is always bigger than the neutral atom. The atomic and ionic radii of Na, F, Na <sup>+</sup> , F <sup>-</sup> are in ppm	<div>15 / 95 1 / 2 136</div> <div>C. Na F Na<sup>+</sup> F<sup>-</sup></div> <div>72 95 136 157</div> <div>D. Na F Na<sup>+</sup> F<sup>-</sup></div> <div>157 136 95 72</div>
16	The energy absorbed when an electron is added to a gaseous atom to form a gaseous ion is called	<div>A. Electron affinity</div> <div>B. Ionization energy</div> <div>C. Both of these</div> <div>D. None of these</div>
17	Given data (i) heat of formation of CO <sub>2</sub> is -393.7 KJ mole <sup>-1</sup> (ii) heat of formation of H <sub>2</sub> O is -285.8 KJ mole <sup>-1</sup> (iii) heat of combustion of CH <sub>4</sub> is -890.00 KJ mole <sup>-1</sup> Enthalpy of formation of methane from C and H <sub>2</sub> is calculated by Hess's law by	<div>A. Adding i + ii + iii</div> <div>B. Adding 2(i) and ii and subtracting iii</div> <div>C. adding i + iii and subtracting ii</div> <div>D. Adding i + 2(ii) and subtracting iii</div>
18	Acetone reacts with HCN to form a cyanohydrin. It is an example of	<div>A. Electrophilic addition</div> <div>B. Electrophilic substitution</div> <div>C. Nucleophilic addition</div> <div>D. Nucleophilic substitution</div>
19	Aliphatic carboxylic acids have carboxyl group attached to:	<div>A. Alkyl group</div> <div>B. Aryl group</div> <div>C. Phenyl group</div> <div>D. Benzyl group</div>
20	The order of reactivity of halogen acids towards alkenes	<div>A. HCl &gt; HBr &gt; HI</div> <div>B. HBr &gt; HCl &gt; HI</div> <div>C. HCl &gt; HBr</div> <div>D. HI &gt; HBr &gt; HCl</div>