

ECAT Chemistry Online Test

Sr	Questions	Answers Choice
1	Bohr model of atom is contradicted by:	A. Planck's quantum theory B. Pauli's exclusion theory C. Heisenberg's uncertainty principle. D. All of above.
2	Which one of the following step is not involved in determination of empirical formula	A. Determination % of each element B. Determination of gram atom of each element C. Determination of isotopes of each element D. Determination of atomic ratio of element
3	Which is a good nucleophile?	A. F ⁻¹ B. Cl ⁻¹ C. Be ⁻¹ D. I ⁻¹
4	Which one of the following is the most abundant organic substance found in nature	A. Fructose B. Starch C. Glucose D. Cellulose
5	Lattice energy of NaCl	A. +5000 KJ B344 KJ C776 KJ D411 KJ
6	Chlorofluorocarbon, CF_2Cl_2 , plays an effective role in removing $O_3(ozone)$ in the stratosphere. Which reaction does not cause the depletion of ozone	
7	Which of the following oxides is amopheric in character?	A. CaO B. CO ₂ C. SiO ₂ D. SnO ₂
8	Which woody raw material is used of the manufacture of paper pulp	A. Cotton B. Bagasse C. Poplar D. Rice straw
9	Which is the formula of tetra-ammine chloronitro platinum (VI) sulphate?	A. [Pt(NH ₃) ₄ (NO ₂)]SO ₄ B. [PtNO ₂ Cl(NH ₃) ₄ C. [PtCl(NO ₂) (NH ₃) C. [PtCl(NO ₂) (NH ₃) [Pt(NH ₃) Sub>4 D. [Pt(NH ₃) ₄ (NO ₂)
10	If the radius of first Bohr orbit be a_0 , then the radius of third Bohr orbit would be	A. 3 x a _o B. 6 x a _o C. 9 x a _o D. 1/2 x a _o
11	Which of the following species has unpaired electrons in ant-bonding molecular orbitals?	A. O2 ⁺² <o:p></o:p> B. N2 ⁻² <o:p></o:p> C. B ₂ <o:p></o:p> D. F ₂ <o:p></o:p> D. F ₂ <o:p></o:p> D. F ₂ <o:p></o:p>
12	Diversity of organic compounds in millions is:	A. Four B. Five C. Six D. Seven

13	Which one of the following has continuous solubility curve	A. NH ₄ NO ₃ B. CaCl C. CaCl ₂ . 6H ₂ O D. Na ₂ SO ₄ . 10H ₂ O
14	Carbon atom holding halogen in aryl halides is	A. sp ² -hybridesed B. sp ³ -hybridesed C. sp-hybridesed D. sp ³ d-hybridesed
15	The value of charge on electron is	A. 1.602 x 10 ⁻¹⁹ coulombs B. 1.602 x 10 ⁻¹⁸ coulombs C. 1.602 x 10 ⁻¹⁷ coulombs D. 1.602 x 10 ⁻¹⁶ coulombs
16	Energy can neither be created nor destroyed, although it can be transformed from one form to another. This is a statement of	A. Law of conservation of matter B. Law of definite proportions C. Law of conservation of energy D. None of these
17	The bond order of individual C - C bond is benzene is	A. One B. Two C. Between one and two D. One and two alternately
18	The elements of group IA are called	A. Chalocogens B. Halogens C. Alkali metals D. Alkaline earth metals
19	The weight of pure NaOH required to prepare 250 cm ³ of 0.1 N solution is	A. 4 g B. 1 g C. 2 g D. 5 g
20	On hydrolysis of starch, we finally get	A. Glucose B. Fructose C. Both D. Sucrose