ECAT Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	Alkaline earth metals are usually	A. Reducing agent B. Oxidizing agent C. Amphoteric D. Acidic
2	Select from the following which one is alcohol?	A. CH₃-CH₂-OH B. CH₃-O-CH₃ C. CH₃COOH D. CH₂-CH₂-BR
3	Which of the following is not the reactions of lithium	
4	During the preparation of alkanes the hydrogenation of alkenes or alkynes the catalyst may be	A. H₂SO₄ B. Ni C. Fe₂O₃ D. Al₂O₃
5	Among the following the most reactive towards alcoholic KOH is	A. CH₂= CHBr B. CH₃COCH₂CH₂Br C. CH₃CH₂Br D. CH₃CH₂CH₂Br
6	The reaction of aromatic acyl chloride and phenol in the presence of a base NaOH or pyridine is called	A. Kolbe's reaction B. Perkin's reaction C. Sandmeyer's reaction D. Schotten Baumann reaction
7	The orbitals having n + l = 5 are	A. 2p, 3d,3s B. 3p, 3d, 5s C. 3s, 4p, 4d D. 5s, 4p, 3d
8	General formula of aliphatic carboxylic acids:	A. R---OH B. R----COH C. R----CO---R D. RCOOH
9	Sulphuric acid generates nitronium ion by reacting with	A. Nitric acid B. Nitrogen gas C. Nitrous acid D. Potassium nitrate
10	Acetylene had a characteristic ethereal smell resembling that of	A. Ginger B. Vinegar C. Garlic D. Onion
11	How many secondary alcoholic groups are present in the structure of glucose OHC CHOH CHOHCH OH CHOH CH_2OH	A. 1 B. 2 C. 3 D. 4
12	What happens when reaction is at equilibrium and more reactant is added :	A. Forward reaction rate is increased. B. Forward reaction rate is decreased. C. Backward reaction rate is increased. D. Equilibrium remains unchanged.
13	Charge of an electron is:	A. <span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">1.6 x 10⁻¹⁹ C<o:p></o:p> B. 9.1 x <span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">10⁻³⁴ C<o:p></o:p> C. 1.7588 x <span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-

		serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">10¹¹ C<o:p></o:p> D. 6.62 x <span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">10⁻³⁴ C<o:p></o:p>
14	Homologues differ from each other by an integral number of	A. CH₂groups B. CH₃groups C. CH groups D. CH₄groups
15	Ethanal may be converted into a three-carbon acid in a two-step process. Which compound is the intermediate	A. CH₃CO₂H B. CH₃CN C. CH₃CH₂CN D. CH₃CH(OH)CN
16	Which is possible in ethers?	A. Reactivity high B. Oxidation and reduction C. Reactivity towards bases D. Towards acids
17	All C - H bond lengths of benzene ring is	A. 1.07A° B. 1.09A° C. 1.08A° D. None
18	Electrons in degenerate orbitals are placed in separate orbitals with same spin according to	A. Hund's rule B. Pauli exclusion principle C. Aufbau principle D. Mosley's law
19	Which of the following molecules have multiple bonds	A. CH₄ B. C₂H₄ C. C₂H₆ D. CCl₄
20	Half life period of a first order reaction is independent of:	A. Presence of catalyst. B. Conditions of temperature C. Initial concentration of the compound D. All of above