

ECAT Chemistry MCQ's Test For Full Book

Answers Choice The depression of freezing point is directly proportional to Answers Choice B. Molarity of the solution C. Mostrity of the solution D. Molarity D. None A 4th C. Gith D. 7 th D.			
1 The depression of freezing point is directly proportional to Chebric of the solution Chebric of the	Sr	Questions	Answers Choice
The sun of all the energies of atoms, molecule, ion, within system is called Series starting form 3gV to 4gCD is in period: Which of the following gives icdoform on heating with a solution of I ₂ containing Ns ₂ CO ₃ ? All the following decompose easily on heating to give oxygen except All the following decompose easily on heating to give oxygen except A solution can be A solution can be A solution can be A solution can be A colors separated by dark spaces There are no bouldards Spectrum of white light is continuous becuase All covalent bonds formed between the two atoms are non-polar when All covalent bonds formed between the two atoms are non-polar when The experimental relationship between a reaction rate and the concentration of reactants is called The experimental relationship between a reaction rate and the concentration of reactants is a called and complex to the start of the solutions of the same short of the solutions of the same short of same elements The experimental relationship between a reaction rate and the concentration of reactants is a called and subtracted and unsupplications and subtracted region The experimental relationship between a reaction rate and the concentration of reactants is a called and subtracted and unsupplications and subtracted region and the solutions is a called and subtracted region and the solutions is a called and subtracted region and the solutions is a called and subtracted region and the solutions is a called and subtracted region and the solutions is a called and subtracted region and the solutions is a subtracted and subtracted region and the solutions is a subtracted and subtracted region and subtracte	1	The depression of freezing point is directly proportional to	B. Molarity of the solution C. Molality of the solution
Series starting form 3gY to 4gCD is in period: Series starting form 3gY to 4gCD is in period: Which of the following gives iodoform on heating with a solution of becontaining NazCO37 A Ethyl alcohol B. Acetone C. Bith and on the following gives iodoform on heating with a solution of becontaining NazCO37 All the following decompose easily on heating to give oxygen except A Solution can be A Colors separated by dark spaces B. Saturated and unsaturated D. Superstaturated and saturated D. Solution can define C. Shurated and saturated D. Solution can define C. Shurated and saturated D. Solution can define C. Shurated and saturated D. Solution can be A Colors separated by dark spaces B. There are no boundary lines between the color can be contained by the solutions are in infrared region D. C. The radiations are in infrared region D. C. The radiations are in infrared color can be contained by the solution series in the solution of the solution series in the solution	2	The sun of all the energies of atoms, molecule, ion, within system is called	B. K.E. of the system C. Internal energy
4 Which of the following gives iodoform on heating with a solution of I ₂ containing Na ₂ CO ₃ ? B. Acistone C. Ethyl alcohol as well as acetone D. Methyl alcohol D. Methyl	3	Series starting form 39Y to 48CD is in period:	B. 5th C. 6th
All the following decompose easily on heating to give oxygen except A plustured and concentrated and start and and dilute constant and dilute constant and sturrated and unsaturated and sturrated and unsaturated constant and sturrated and unsaturated constant and sturrated and unsaturated and sturrated and unsaturated constant and sturrated and unsaturated constant and sturrated and unsaturated constant and sturrated and unsaturated and unsaturated constant and sturrated and unsaturated unsaturated unsaturated and unsat	4	Which of the following gives iodoform on heating with a solution of I ₂ containing Na ₂ CO ₃ ?	B. Acetone C. Ethyl alcohol as well as acetone
8 Saturated and dilute C. Supersaturated and saturated D. Supersaturated and dilute C. Supersaturated and saturated D. Supersaturated and dilute D. Supersaturated and dilute D. Supersaturated and dilute D. Supersaturated and dilute D. Supersaturated and saturated D. Supersaturated and dilute D. Supersaturated and dilute D. Supersaturated and dilute D. Supersaturated and dilute D. Supersaturated and saturated D. Supersaturate and satu	5	All the following decompose easily on heating to give oxygen except	B. Potassium chlorateC. Mercuric oxide
For the are no boundary lines between the colours C. The radiations are in infrared region D. The radiations and the content of th	6	A solution can be	B. Saturated and dilute C. Saturated and unsaturated
8 Lattice energy of NaCl 8344 KJ C776 KJ D411 KJ A. Covalent bond between two non-metal atoms B. Covalent bond between two non-metal atoms B. Covalent bond between metal and non-metal atoms B. Covalent bond between metal and non-metal atoms C. Covalent bond between two atoms of same element D. Covalent bond between metal and same element D. Covalent bond between metal atoms B. Rate law C. Activated complex D. Molecularity A. Gases B. Liquids C. Waxy solids D. Solids 12 In which type of following solutions we don't know the total volume of the solutions: B. Percentage weight/weight B. Percentage volume/volume C. Percentage volume/volume D. Percentage volume/volume C. Percentage volume/weight D. Covalent bond between metal and non-metal atoms A. Order or reaction B. Rate law C. Activated complex D. Molecularity A. Gases B. Liquids C. Waxy solids D. Solids A. Percentage weight/volume C. Percentage volume/volume D. Percentage volume/weight B. Percentage volume/weight A. Change in internal energy B. Enthaply change C. Temperature change D. Work done by the system A. Electrophilic addition B. Electrophilic substitution	7	Spectrum of white light is continuous becuase	B. There are no boundary lines between the colours C. The radiations are in infrared region D. The radiatins fall in ultraviolet
9 All covalent bonds formed between the two atoms are non-polar when 10 The experimental relationship between a reaction rate and the concentration of reactants is called 11 Alkanes containing carbon C ₁₈ ownwards are 12 In which type of following solutions we don't know the total volume of the solutions: 13 The heat energy change during a chemical reaction at constant pressure and at a given temperature is called 14 Ouestion Image All covalent bond between metal and non-metal C. Covalent bond between two atoms of same element D. Covalent bond between metal and non-metal C. Covalent bond between two atoms of same element D. Covalent bond between two atoms of same element D. Covalent bond between metal and non-metal C. Covalent bond between two atoms of same element D. Covalent bond between metal and non-metal C. Covalent bond between metal and non-metal C. Covalent bond between two atoms of same element D. Covalent bond between metal and non-metal C. Covalent bond between metal and non-metal C. Covalent bond between two atoms of same element D. Covalent bond between metal and non-metal covalent bond between two atoms of same element D. Covalent bond between metal and non-metal covalent bond between metal and non-metal covalent bond between two atoms of same element D. Covalent bond between metal and non-metal covalent bond between metal and non-m	8	Lattice energy of NaCl	B344 KJ C776 KJ
The experimental relationship between a reaction rate and the concentration of reactants is called C. Activated complex D. Molecularity A. Gases B. Liquids C. Waxy solids D. Solids In which type of following solutions we don't know the total volume of the solutions: A. Percentage weight/weight B. Percentage weight/volume C. Percentage volume/volume D. Percentage volume/volume D. Percentage volume/weight A. Change in internal energy B. Enthaply change C. Temperature change D. Work done by the system A. Electrophilic addition B. Electrophilic substitution	9	All covalent bonds formed between the two atoms are non-polar when	metal atoms B. Covalent bond between metal and non-metal C. Covalent bond between two atoms of same element D. Covalent bond between metal
Alkanes containing carbon C ₁₈ ownwards are B. Liquids C. Waxy solids D. Solids A. Percentage weight/weight B. Percentage weight/volume C. Percentage volume/volume D. Percentage volume/weight The heat energy change during a chemical reaction at constant pressure and at a given temperature is called A. Change in internal energy B. Enthaply change C. Temperature change D. Work done by the system A. Electrophilic addition B. Electrophilic substitution	10		B. Rate law C. Activated complex
In which type of following solutions we don't know the total volume of the solutions: B. Percentage weight/volume C. Percentage volume/volume D. Percentage volume/weight A. Change in internal energy B. Enthaply change C. Temperature change D. Work done by the system A. Electrophilic addition B. Electrophilic substitution	11	Alkanes containing carbon C ₁₈ ownwards are	B. Liquids C. Waxy solids
The heat energy change during a chemical reaction at constant pressure and at a given temperature is called B. Enthaply change C. Temperature change D. Work done by the system A. Electrophilic addition B. Electrophilic substitution	12	In which type of following solutions we don't know the total volume of the solutions :	B. Percentage weight/volumeC. Percentage volume/volume
14 Ouestion Image B. Electrophilic substitution	13		B. Enthaply change C. Temperature change
	14	Question Image	B. Electrophilic substitution

		D. Nuclophilic addition
15	Which of the following is a pseudo solid?	A. <pre>A. <pre>class="MsoNormal">CaF₂ <o:p></o:p> B. Glass<o:p> </o:p> C. NaCL<o:p> </o:p> D. All<o:p> </o:p></pre></pre>
16	Which one of the following statements is true about discovery of neutrons?	 A. These particles were formed by the bombardment of Alpha-particles on Beryllium. B. These particles are formed by the spiting of alpha-particles. C. These particles were discovered by natural radioactivity. D. None of above.
17	At present cement plants in Pakistan are:	A. Ten B. Twenty Two C. Four D. Twenty Four
18	Which is a mixture of low boiling hydrocarbon	A. Natural gas B. Petroleum C. Wood D. Graphite
19	In endothermic reactions, the heat content of the:	A. Products is more than that of reactants. B. Reactants is more than than to products. C. Both (a) and (b). D. Reactants and products are equal.
20	$N_2 3H_2 \!\!\!=\! 2NH_3$ Which of the following change will favorthe formation of more $^{NH} \!\!\! 3$ at equilibrium in above reaction :	A. By adding NH _{3.} B. By removing H _{2.} C. By decreasing pressure. D. By increasing pressure.