ECAT Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	Question Image	A. Decrease in temperature favour more dissolution of the salt B. Increase in temperature favour more dissolution of the salt C. Lowering pressure favour more dissolution of the salt D. Increasing pressure favour more dissolution of the salt
2	Which has soapy touch?	A. Na2B4O7 B. H3BO3 C. Ca2B6O11 D. HBO2
3	Question Image	A. 450°C B. 250°C C. 850°C D. 1000°C
4	The order of decreasing ease of reaction with ammonia is	A. Anhydrides, esters, ethers B. Anhydrides, ethers, esters C. Ethers, anhydrides, esters D. Esters, ethers, anhydrides
5	Which of the following discoveries resulted in a version of the Mendeleefs periodic law	A. The nucleus of atom by Rutherford B. The elements polonium and radium by the Curies C. Atomic numbers by Moseley D. x-rays by Roentgen
6	which one of the following is a heteroheneous catalysis	
7	The osmotic pressure of solution increases if	A. Temperature is decreased B. Solution constant is increased C. Number of solute molecules are increased D. Volume is increased
8	Alcohals can be distinguished using test:	A. Lucas B. Tollen's C. Koib's D. William's
9	$CnH_{2n}$is the general formula of	A. Alkanes B. Alkanes C. alkynes D. None of above
10	l-Phenylethanol can be prepared by reaction of benzaldehyde with	A. Methyl iodide an Magnesium B. Methyl bromide C. Methyl bromide and AlBr₃ D. C₂H₅I and Mg
11	Enzymes, in the living systems	A. Provide energy B. Provide immunity C. Transport oxygen D. Catalyze biochemical processes
12	Type of isomerism in BrCH=CHBr is:	A. Structural isomerism B. <blockquote style="margin: 0 0 0 40px; border: none; padding: 0px;">Conformational isomerism</blockquote> C. Geometrical isomerism D. Positional isomerism
13	The addition of unsymmetrical reagent to unsymmetrical alkene is in accordance with the rule:	A. Hund's rule B. Markownikov's rule C. Pauli's Exclusion Principle D. Auf bau Principle
		A. sp hybridized

14	Hybridization of each carbon atom in benzene ring is	B. sp²hybridized C. sp³ D. dsp²
15	Keeping in view the size of atoms, which order is the correct one	A. Mg > Sr B. Ba > Mg C. Lu > Ce D. Cl > I
16	What is the molarity of the solution that contains 20 grams of NaOH in 500 ml of solution [Na = 23, O = 16, H = 1]	A. 0.25 B. 0.5 C. 1 D. 20
17	Hydrochloric acid available in the laboratory is 36% w/w. The density of HCl solution is 1.19 g cm^{-3}. The molarity of HCl solution is	A. 10.23 moles dm⁻³ B. 11.55 moles dm⁻³ C. 11.73 moles dm⁻³ D. 12.67 moles dm⁻³
18	$N_2 3H_2 \rightleftharpoons 2NH_3$ Which of the following change will favorthe formation of more NH_3 at equilibrium in above reaction :	A. By adding NH_{3.} B. By removing<span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">H_{2.} C. By decreasing pressure. D. By increasing pressure.
19	Which of the following reacts with chloroform and base to form phenyl isocynaide?	A. Nitrobenzene B. Phenol C. Chlorobenzene D. Aniline
20	NaH is	A. Ionic hydride B. Complex hydride C. Covalent hydride D. Interstitial hydride